



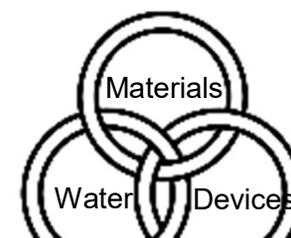
Since 1959

# Chemical reactions and dynamics in nanoparticles

**T. Pradeep**

Institute Professor, IIT Madras

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**May 12-13**

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# Atomically Precise Clusters of Noble Metals: Emerging Link between Atoms and Nanoparticles

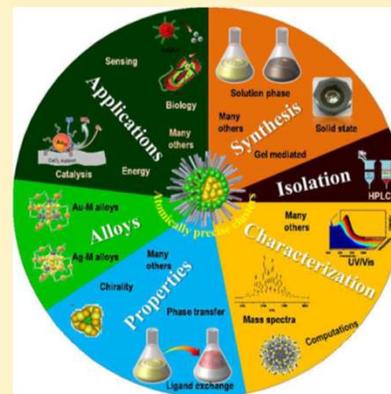
Indranath Chakraborty<sup>†</sup> and Thalappil Pradeep<sup>\*†</sup>

DST Unit of Nanoscience (DST UNS) and Thematic Unit of Excellence, Department of Chemistry, Indian Institute of Technology Madras, Chennai 600036, India

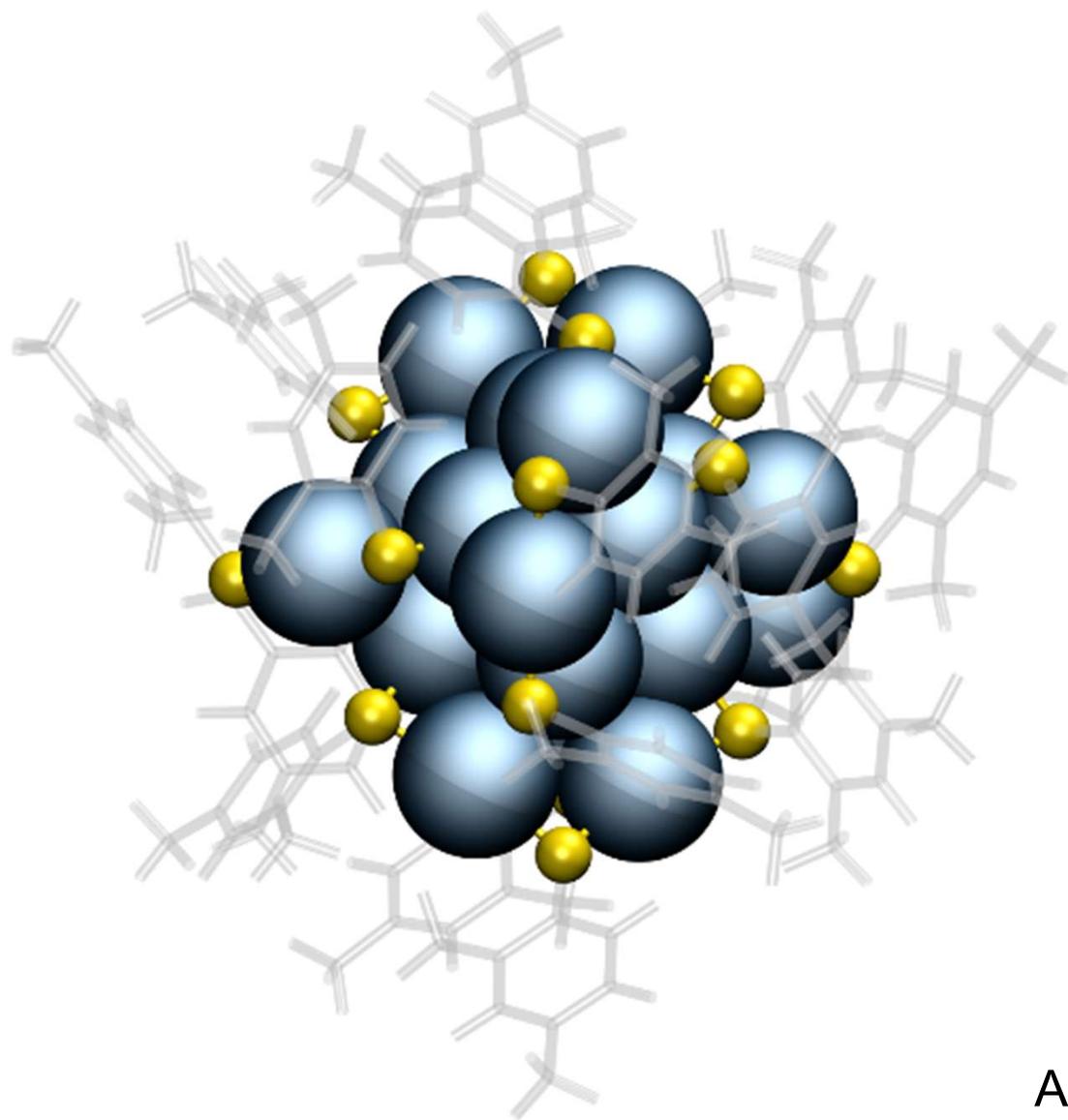
 Supporting Information

Citations: >1100

**ABSTRACT:** Atomically precise pieces of matter of nanometer dimensions composed of noble metals are new categories of materials with many unusual properties. Over 100 molecules of this kind with formulas such as  $\text{Au}_{25}(\text{SR})_{18}$ ,  $\text{Au}_{38}(\text{SR})_{24}$ , and  $\text{Au}_{102}(\text{SR})_{44}$  as well as  $\text{Ag}_{25}(\text{SR})_{18}$ ,  $\text{Ag}_{29}(\text{S}_2\text{R})_{12}$ , and  $\text{Ag}_{44}(\text{SR})_{30}$  (often with a few counterions to compensate charges) are known now. They can be made reproducibly with robust synthetic protocols, resulting in colored solutions, yielding powders or diffractable crystals. They are distinctly different from nanoparticles in their spectroscopic properties such as optical absorption and emission, showing well-defined features, just like molecules. They show isotopically resolved molecular ion peaks in mass spectra and provide diverse information when examined through multiple instrumental methods. Most important of these properties is luminescence, often in the visible–near-infrared window, useful in biological applications. Luminescence in the visible region, especially by clusters protected with proteins, with a large Stokes shift, has been used for various sensing applications, down to a few tens of molecules/ions, in air and water. Catalytic properties of clusters, especially oxidation of organic substrates, have been examined. Materials science of these systems presents numerous possibilities and is fast evolving. Computational insights have given reasons for their stability and unusual properties. The molecular nature of these materials is unequivocally manifested in a few recent studies such as intercluster reactions forming precise clusters. These systems manifest properties of the core, of the ligand shell, as well as that of the integrated system. They are better described as protected molecules or *aspicules*, where *aspis* means shield and *cules* refers to molecules, implying that they are “shielded molecules”. In order to understand their diverse properties, a nomenclature has been introduced with which it is possible to draw their structures with positional labels on paper, with some training. Research in this area is captured here, based on the publications available up to December 2016.



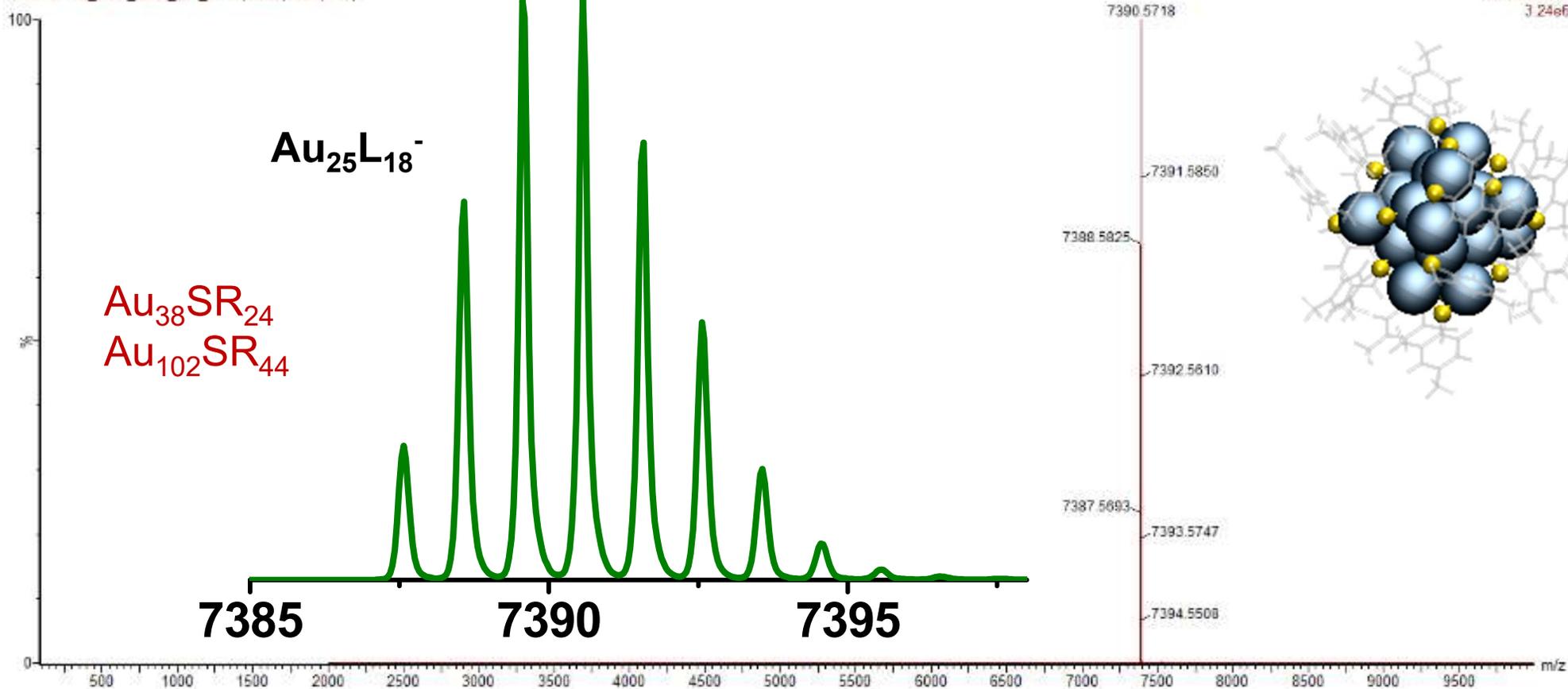
Also the pioneering work of R. W. Murray, Robert L. Whetten, Uzi Landman, Tatuya Tsukuda, Yuichi Negishi, Hannu Hakkinen, Rongchao Jin, Nanfeng Zheng, Terry Bigioni, Osman Bakr, Kornberg, Jianping Xie, C. M. Aikens, Thomas Buergi, Amala Dass, Ackerson, De-en Jiang, .... A. W. Castleman Jr., H. Schmidbauer, .... Robin Ras, Olli Ikkala



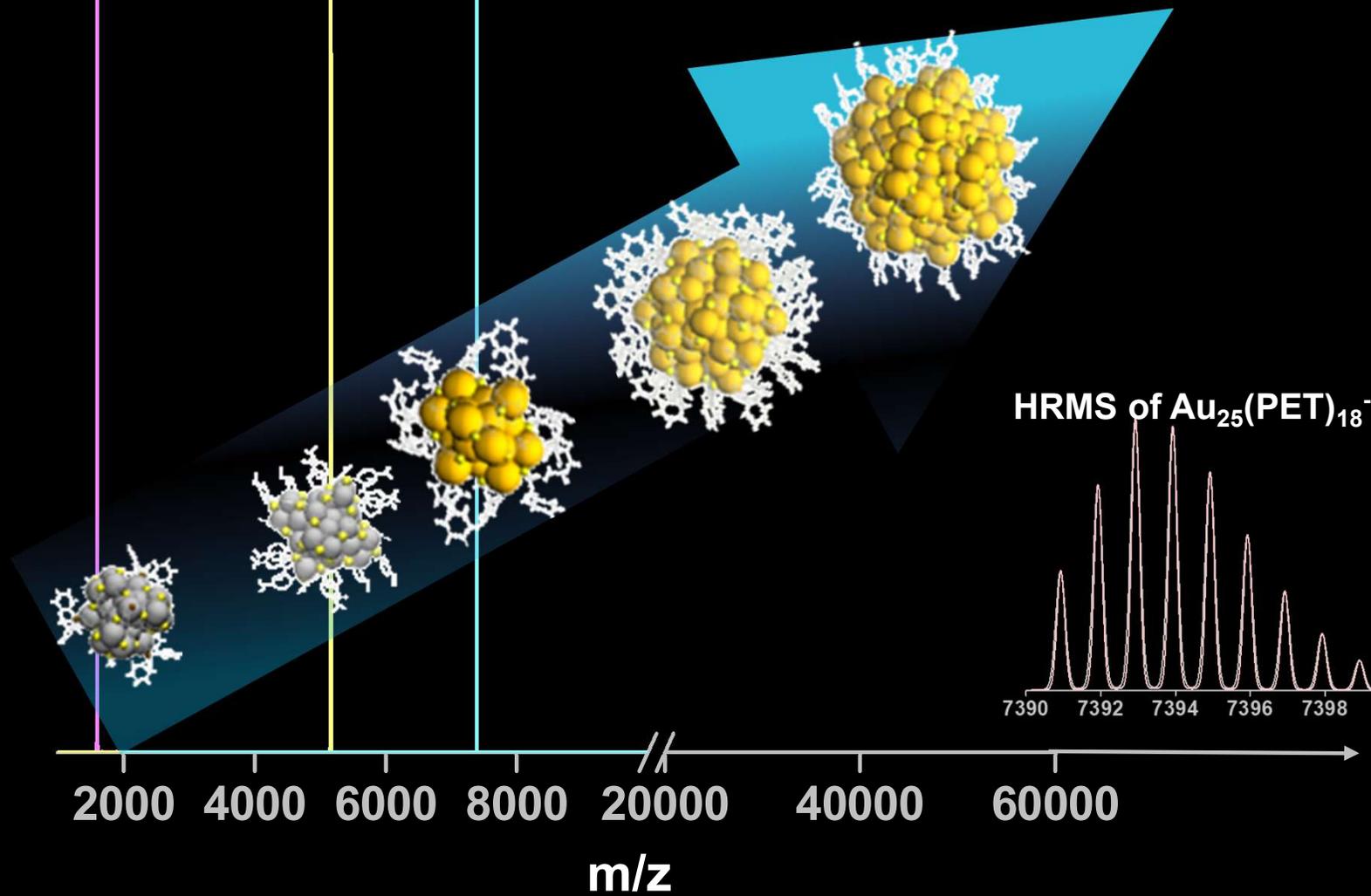
$\text{Au}_{25}, \text{Ag}_{25}, \text{Ag}_{29}$

# Molecular formula, Molecular weight

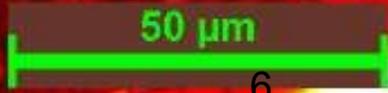
AU25PET18\_RES\_NEG\_MS\_3 32 (0.558) Cm (5.80)



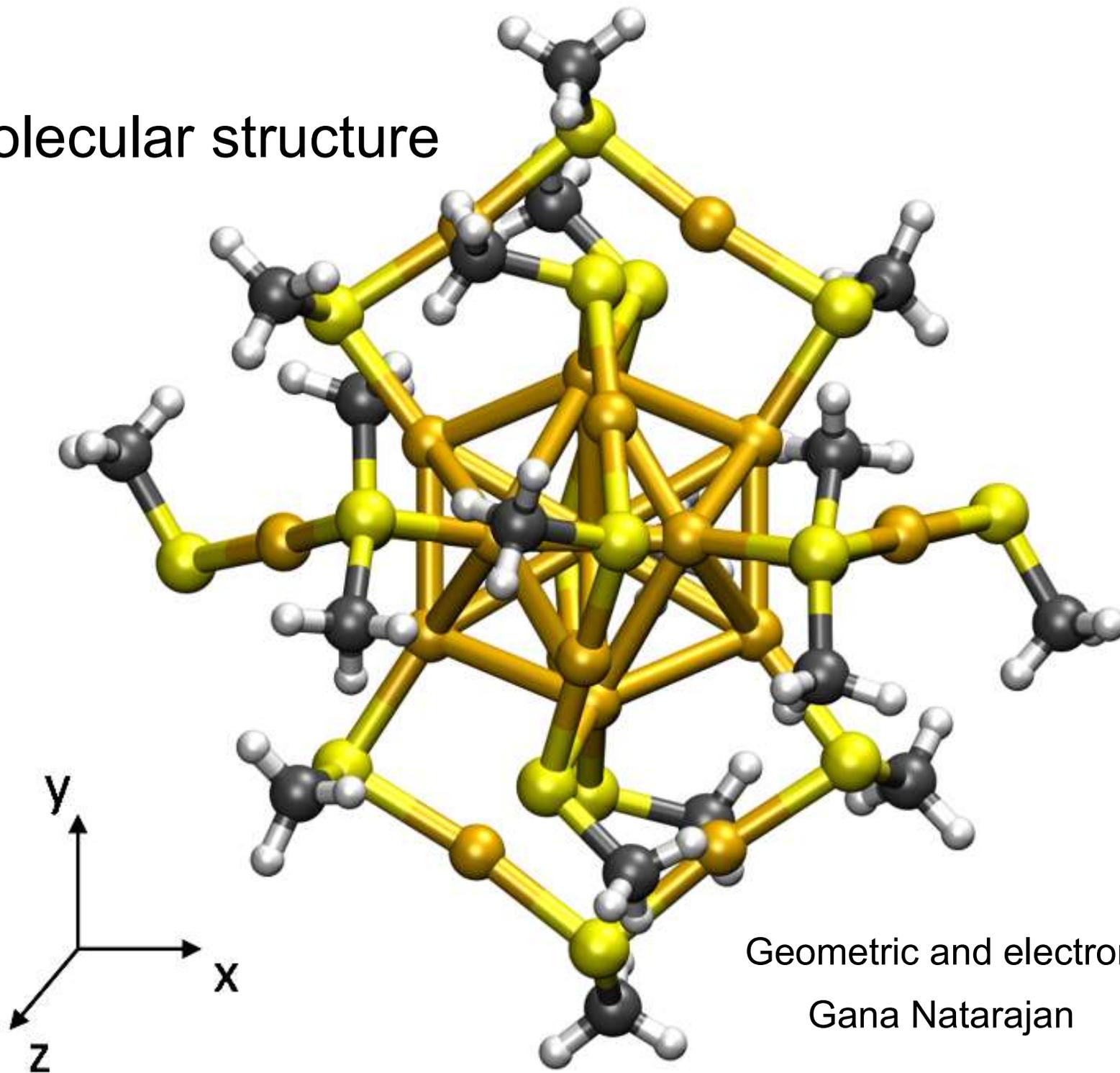
$\text{Ag}_{29}(\text{BDT})_{12}^{3-}$   $\text{Ag}_{25}(\text{DMBT})_{18}^{-}$   $\text{Au}_{25}(\text{PET})_{18}^{-}$



They make high quality crystals



# Molecular structure



Geometric and electronic shells

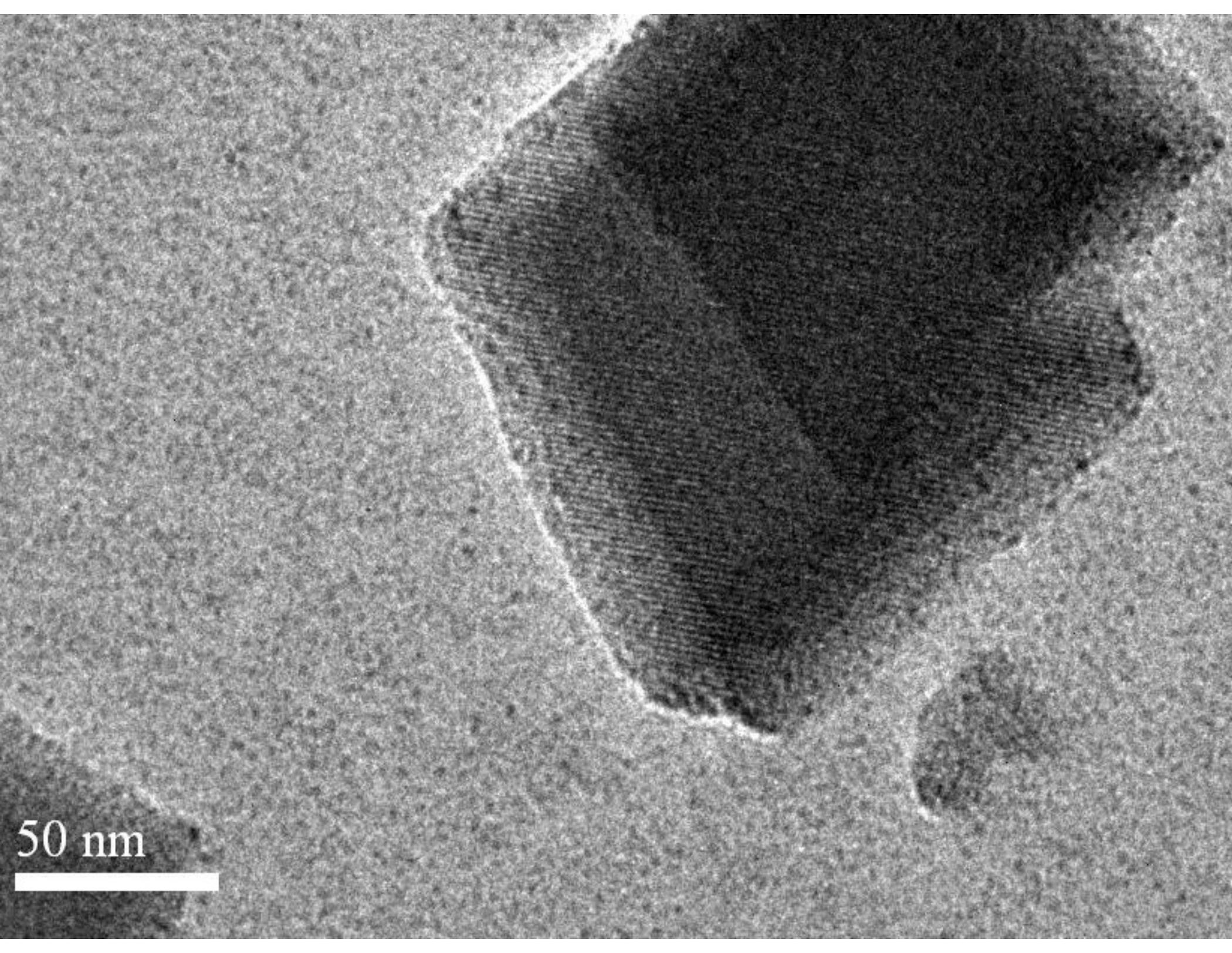
Gana Natarajan

# Molecules and their properties

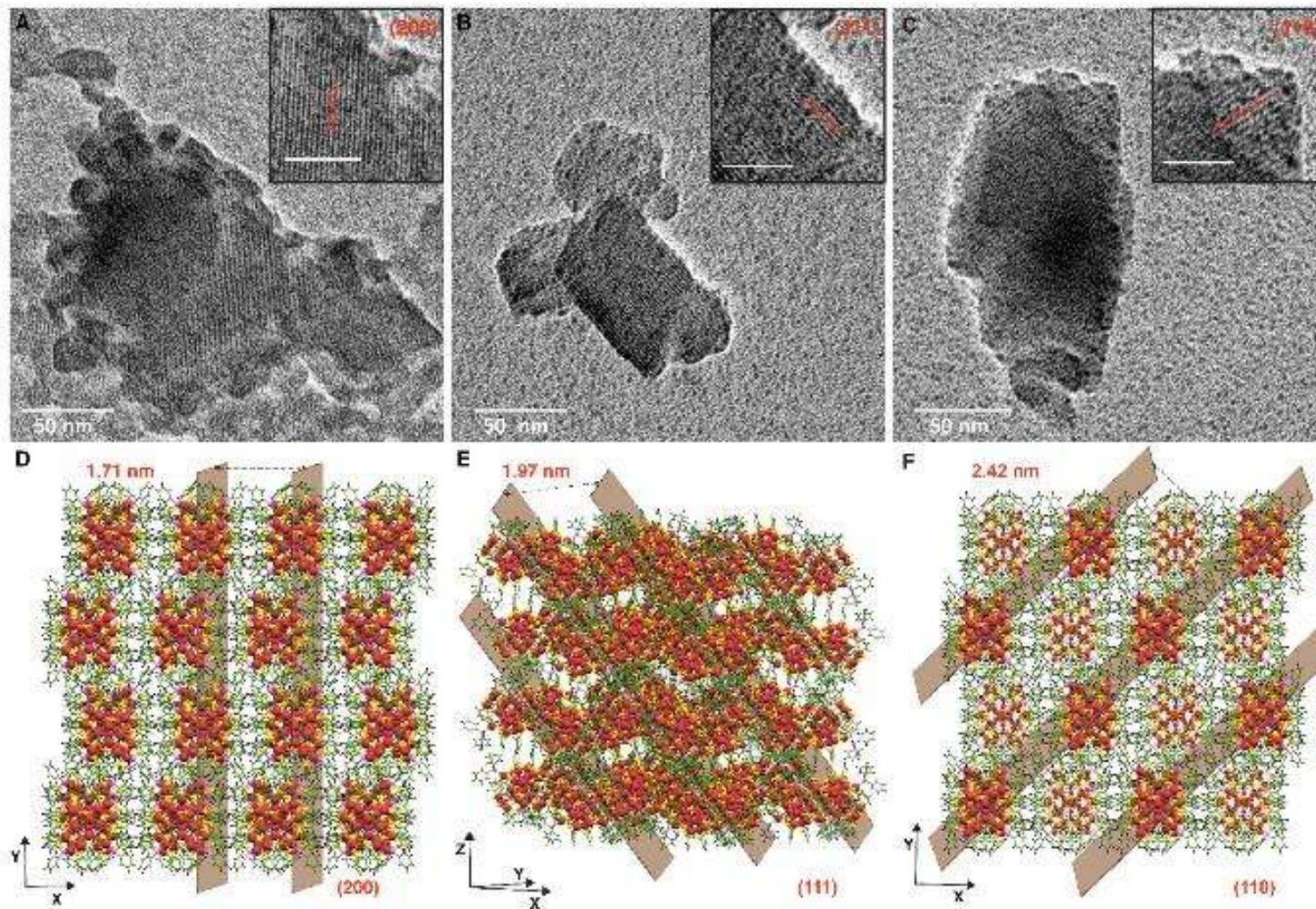
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Chemical formula	H <sub>2</sub> O
Molecular weight	18.0148
Critical temperature	373.91°C
Critical pressure	22.05 MPa
Critical density	315.0 kg/m <sup>3</sup>
Triple point temperature	0.01°C
Triple point pressure	615.066 Pa
Normal boiling point	100.0°C
Normal freezing point	0.0°C
Density of ice at normal melting point	918.0 kg/m <sup>3</sup>
Maximum density, 3.98°C	999.973 kg/m <sup>3</sup>
Viscosity, 25°C	0.889 mN s/m <sup>2</sup>
Surface tension, 25°C	72 mN/m
Heat Capacity, 25°C	4.1796 kJ/kg.K
Enthalpy of vaporisation, 100°C	2,257.7 kJ/kg
Enthalpy of fusion, 0°C	333.8 kJ/kg
Velocity of sound, 0°C	1.403 km/s
Dielectric constant, 25°C	78.40
Electrical conductivity, 25°C	8 μS/m
Refractive index, 25°C	1.333
Liquid compressibility, 10°C	480. × 10 <sup>-12</sup> m <sup>2</sup> /N
Coefficient of thermal expansion, 25°C	256.32 × 10 <sup>-6</sup> K <sup>-1</sup>
Thermal Conductivity, 25°C	0.608 W/m.K

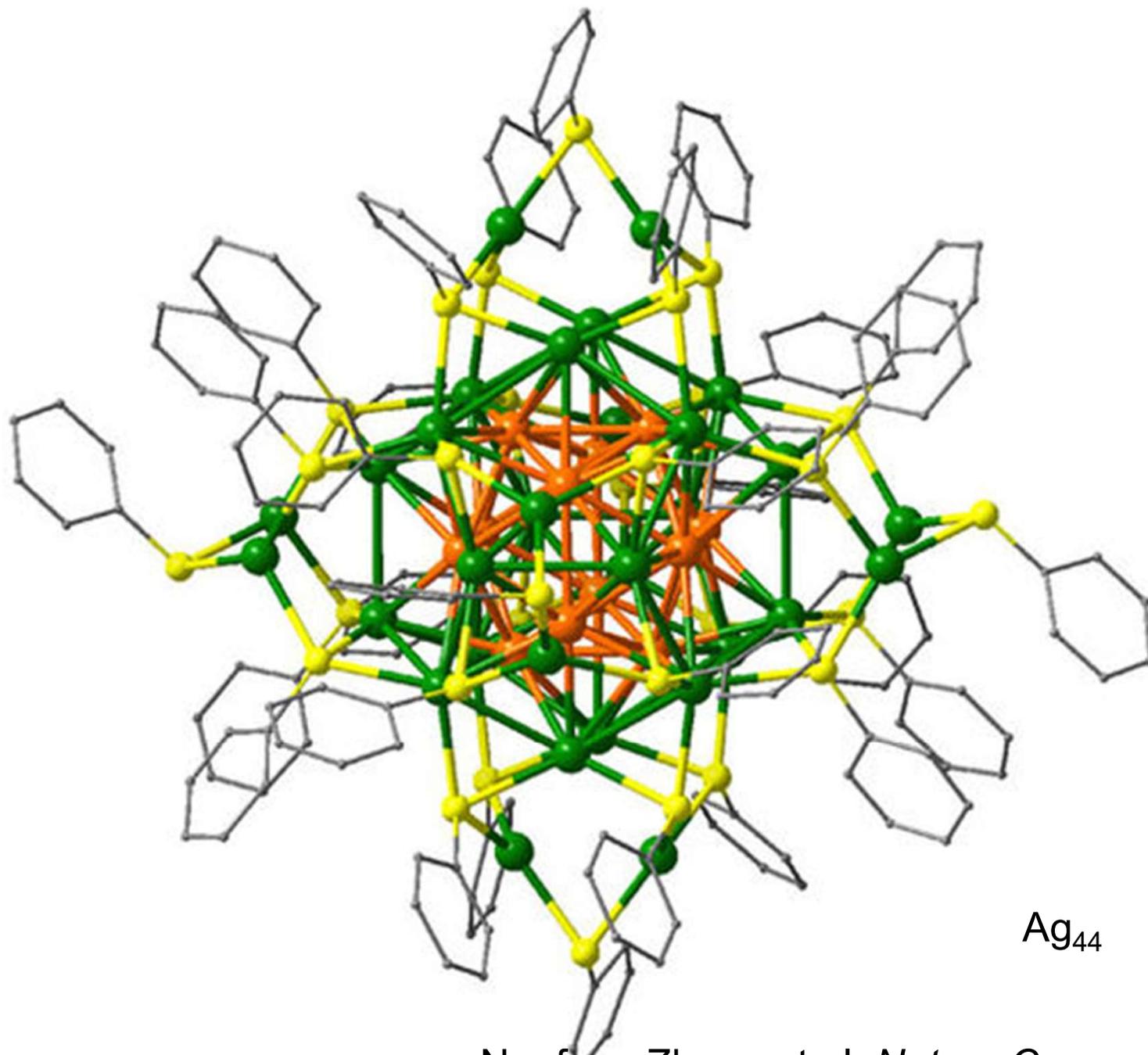
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50 nm



Ananthu Mahendranath et al. Chem.Comm.2021



Ag<sub>44</sub>

Nanfeng Zheng et al. *Nature Communications* 4, 2013  
Terry Bigioni et al. *Nature* 2013

# Molecular reactions



Reactions on clusters

Reactions between clusters

# Inter-cluster reactions

**J | A | C | S**  
JOURNAL OF THE AMERICAN CHEMICAL SOCIETY

Article

[pubs.acs.org/JACS](https://pubs.acs.org/JACS)

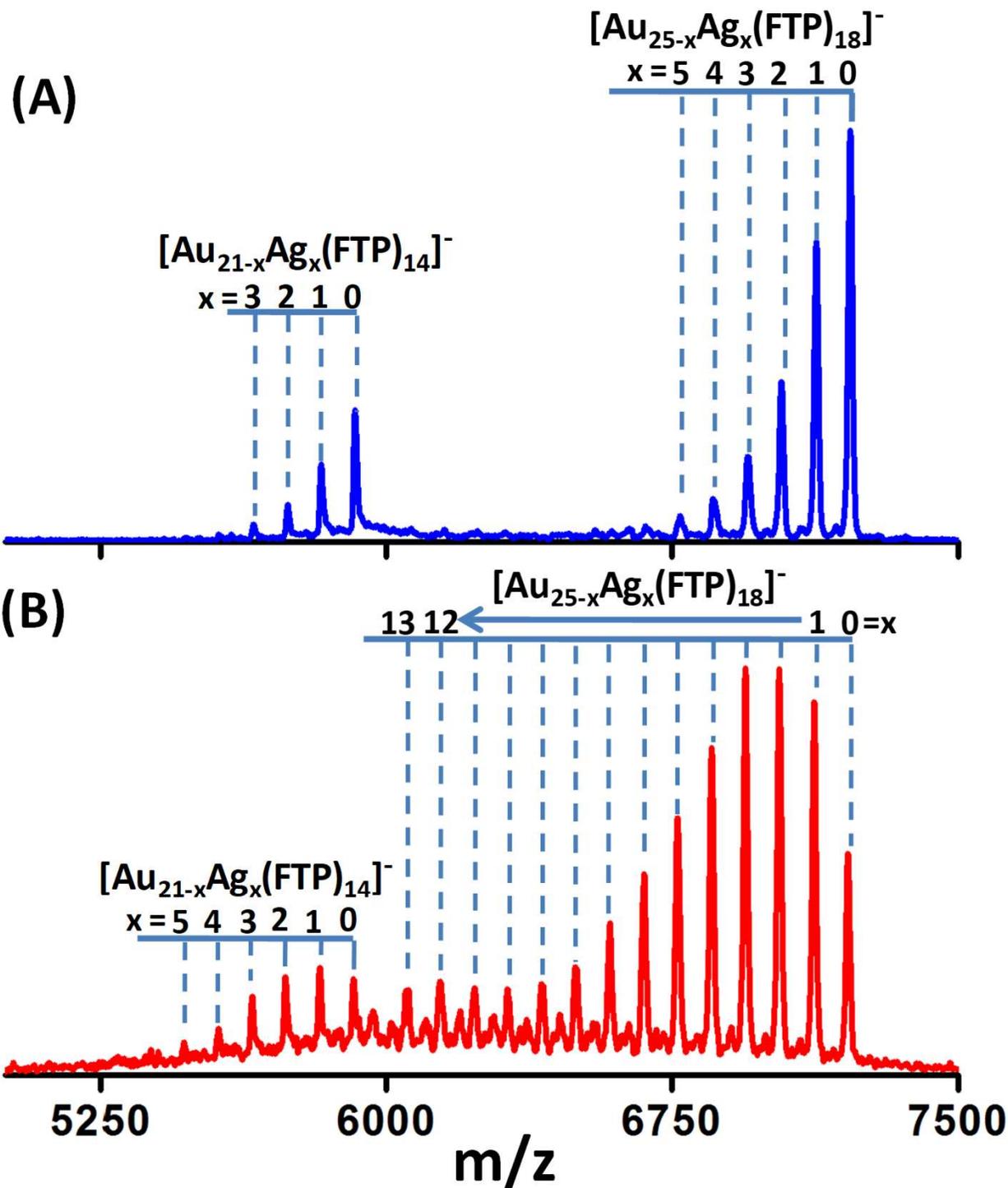
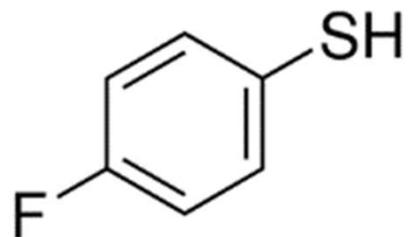
## Intercluster Reactions between $\text{Au}_{25}(\text{SR})_{18}$ and $\text{Ag}_{44}(\text{SR})_{30}$

K. R. Krishnadas, Atanu Ghosh, Ananya Baksi, Indranath Chakraborty,<sup>†</sup> Ganapati Natarajan,  
and Thalappil Pradeep\*

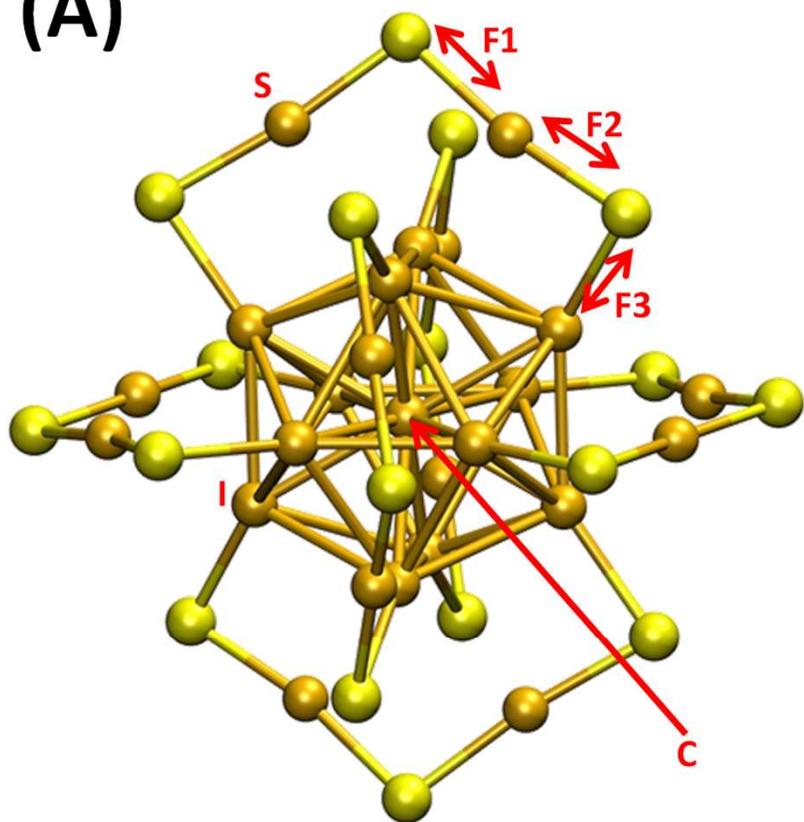
DST Unit of Nanoscience (DST UNS) and Thematic Unit of Excellence, Department of Chemistry, Indian Institute of Technology  
Madras, Chennai, 600 036, India

 Supporting Information

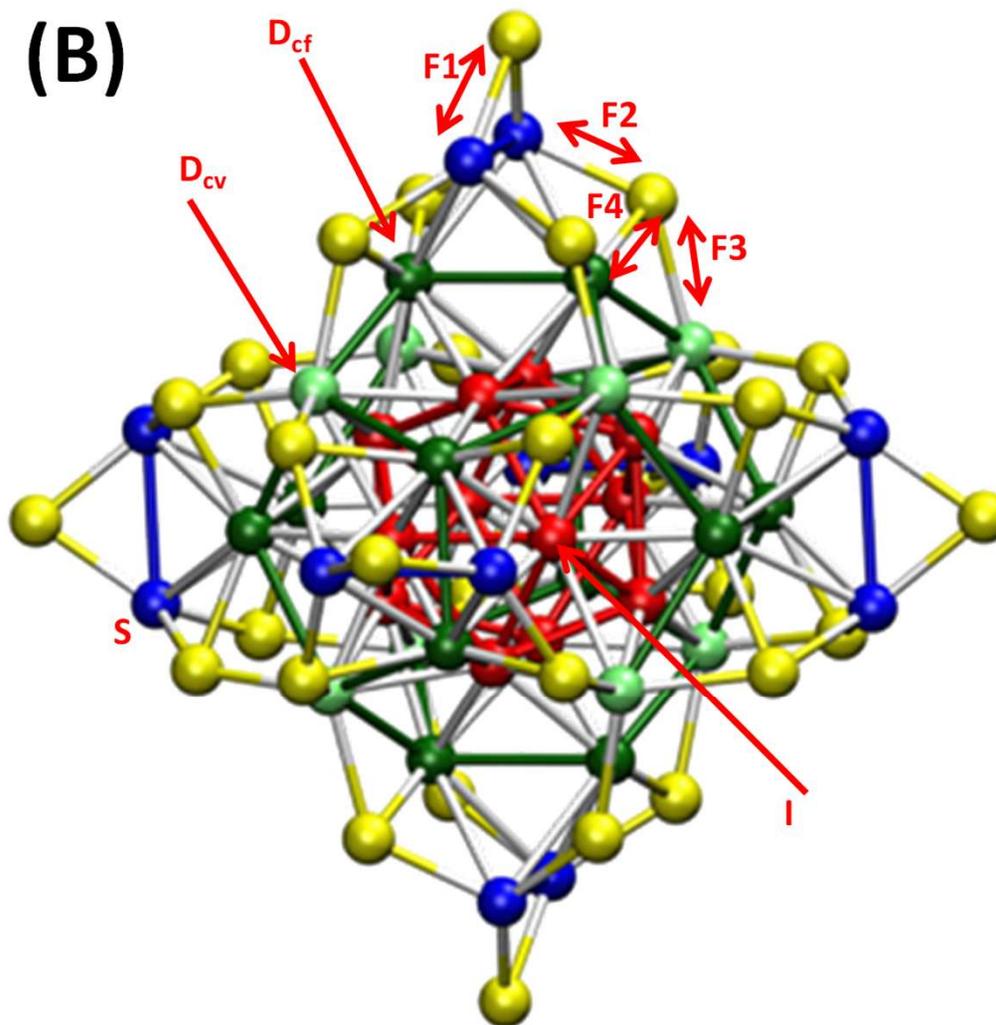




(A)



(B)



**Energies for the substitution reaction of (A) Au in  $\text{Ag}_{44}(\text{SR})_{30}$ , (B) Ag in  $\text{Au}_{25}(\text{SR})_{18}$  and (C) the overall reaction energies (in eV) as a function of their positions in product clusters,  $\text{Au}_x\text{Ag}_{44-x}(\text{SR})_{30}$  and  $\text{Au}_{25-x}\text{Ag}_x(\text{SR})_{18}$  for  $x=1$**

**(A) Location of Au in  $\text{Au}_x\text{Ag}_{44-x}(\text{SR})_{30}$**

	$\Delta E/\text{eV}$
Icosahedron (I)	-0.72
Dodecahedron: cube vertex ( $D_{cv}$ )	-0.14
Dodecahedron: cube face ( $D_{cf}$ )	-0.32
Staples (S)	-0.48

**(B) Location of Ag in  $\text{Au}_{25-x}\text{Ag}_x(\text{SR})_{18}$**

	$\Delta E/\text{eV}$
Central atom (C)	+0.71
Icosahedron (I)	+0.23
Staples (S)	+0.44

**(C) Locations of Au in  $\text{Au}_x\text{Ag}_{44-x}(\text{SR})_{30}$**

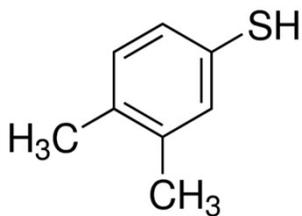
Location of Ag in $\text{Au}_{25-x}\text{Ag}_x(\text{SR})_{18}$	I	$D_{cv}$	$D_{cf}$	S
C	-0.015	+0.564	+0.388	+0.226
I	-0.486	+0.093	-0.083	-0.245
S	-0.276	+0.303	+0.127	-0.035

# **Ag<sub>25</sub>-Au<sub>25</sub> experiments**

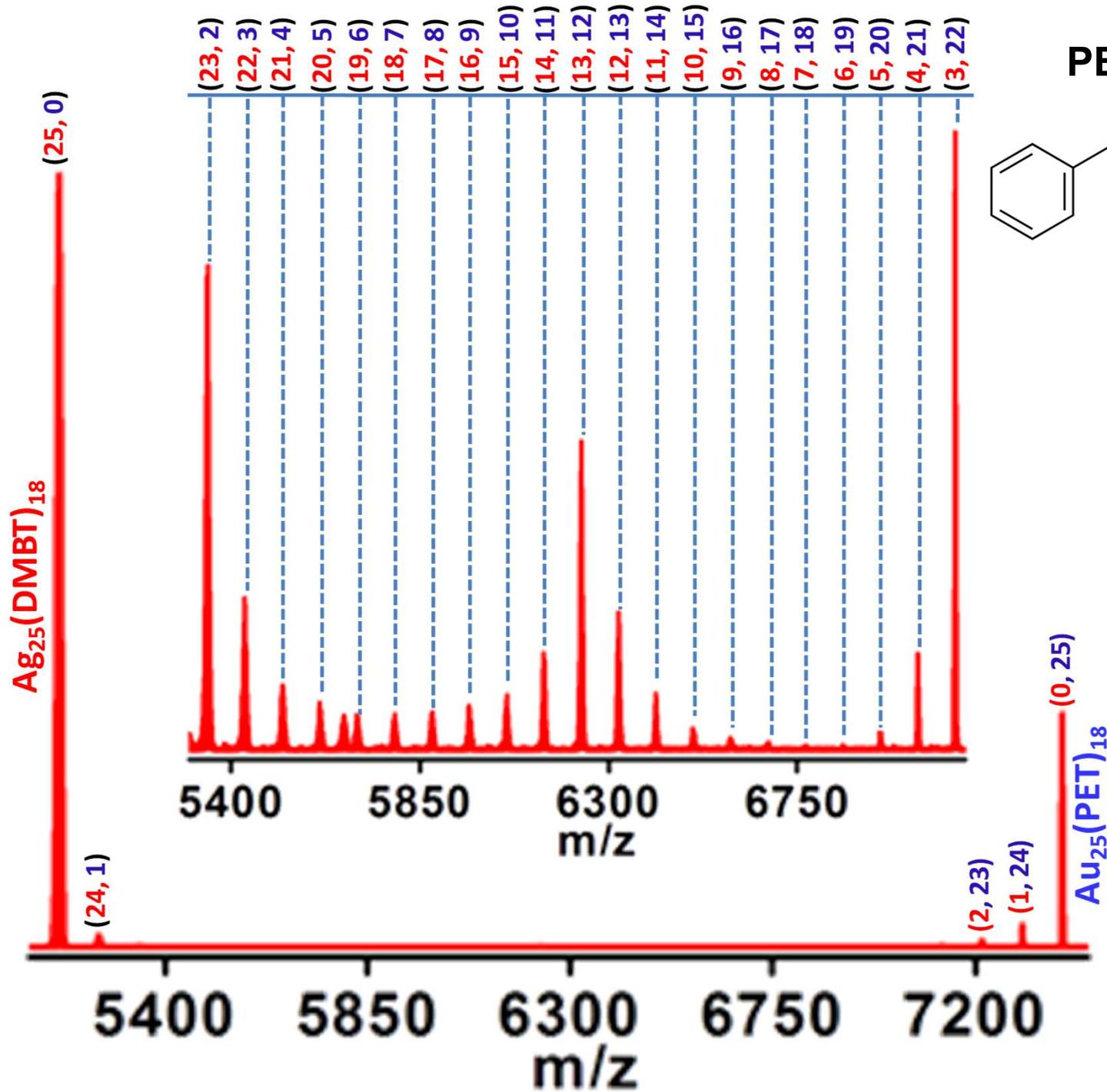
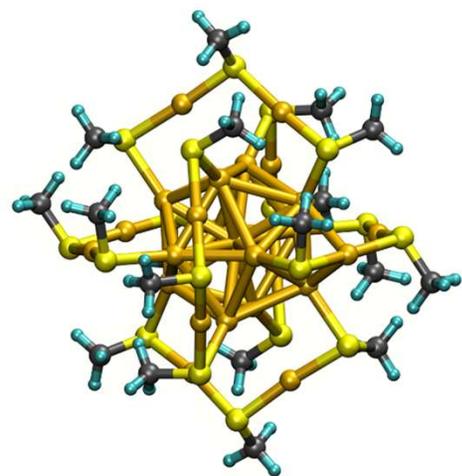
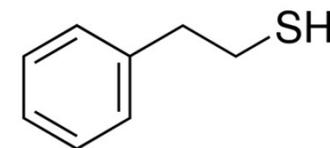
K. R. Krishnadas et al. *Nature Commun.* 2016

# Reaction between $\text{Au}_{25}(\text{PET})_{18}$ and $\text{Ag}_{25}(\text{DMBT})_{18}$

DMBT

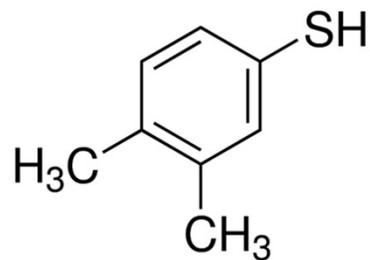


PET

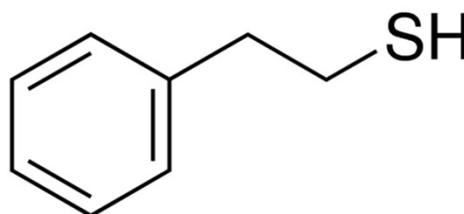


# $[\text{Ag}_{25}(\text{DMBT})_{18} + \text{Au}_{25}(\text{PET})_{18}]^{2-}$

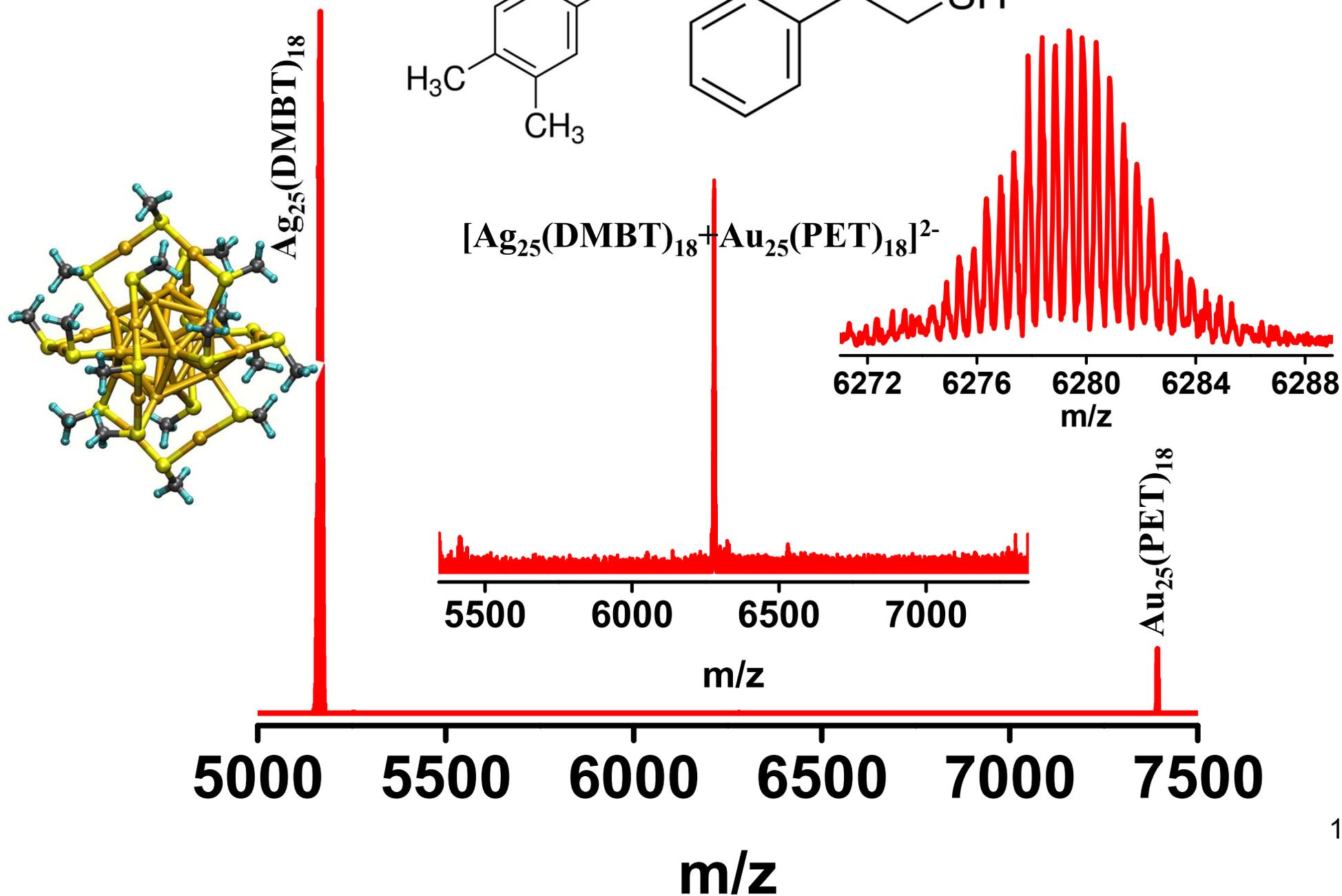
DMBT



PET

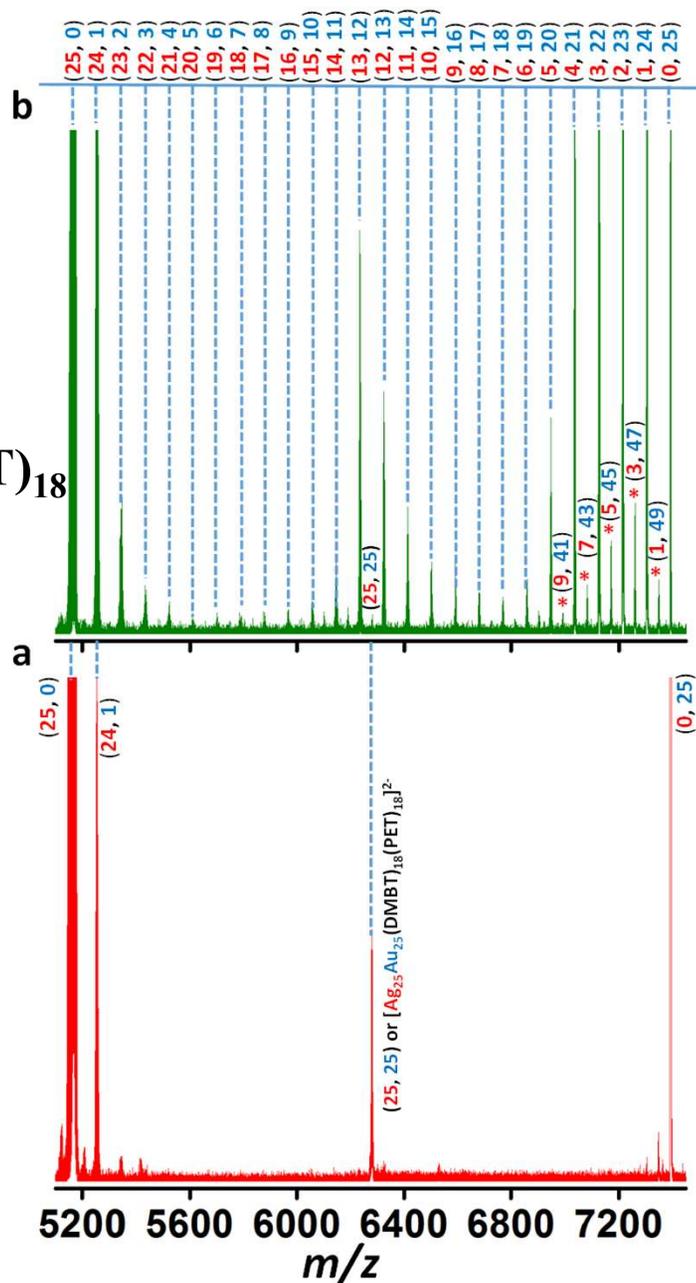


$[\text{Ag}_{25}(\text{DMBT})_{18} + \text{Au}_{25}(\text{PET})_{18}]^{2-}$



# Evolution of alloy clusters from the dianionic adduct, $[\text{Ag}_{25}\text{Au}_{25}(\text{DMBT})_{18}(\text{PET})_{18}]^{2-}$

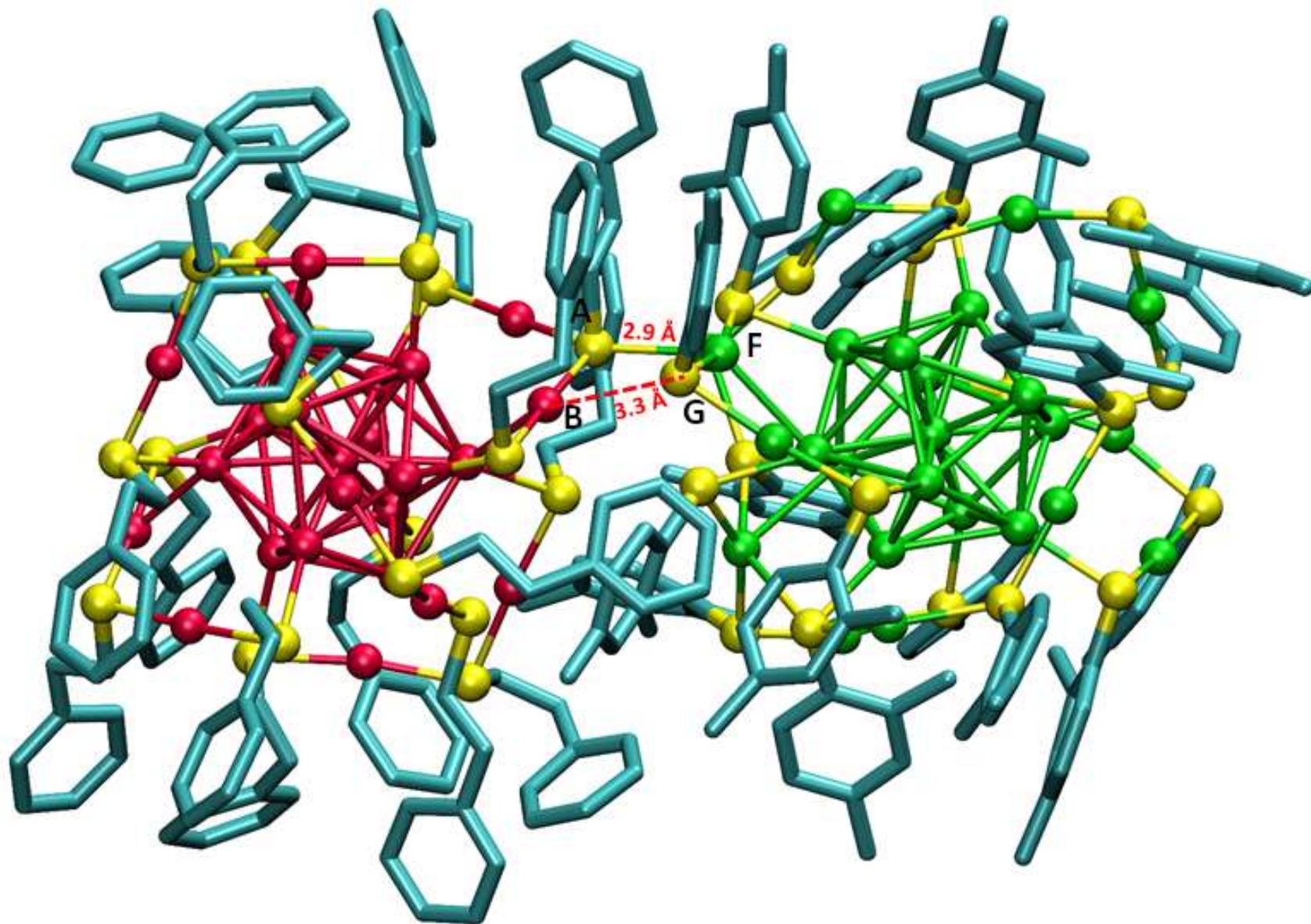
$\text{Ag}_{25}(\text{DMBT})_{18}:\text{Au}_{25}(\text{PET})_{18}$   
0.3:1.0

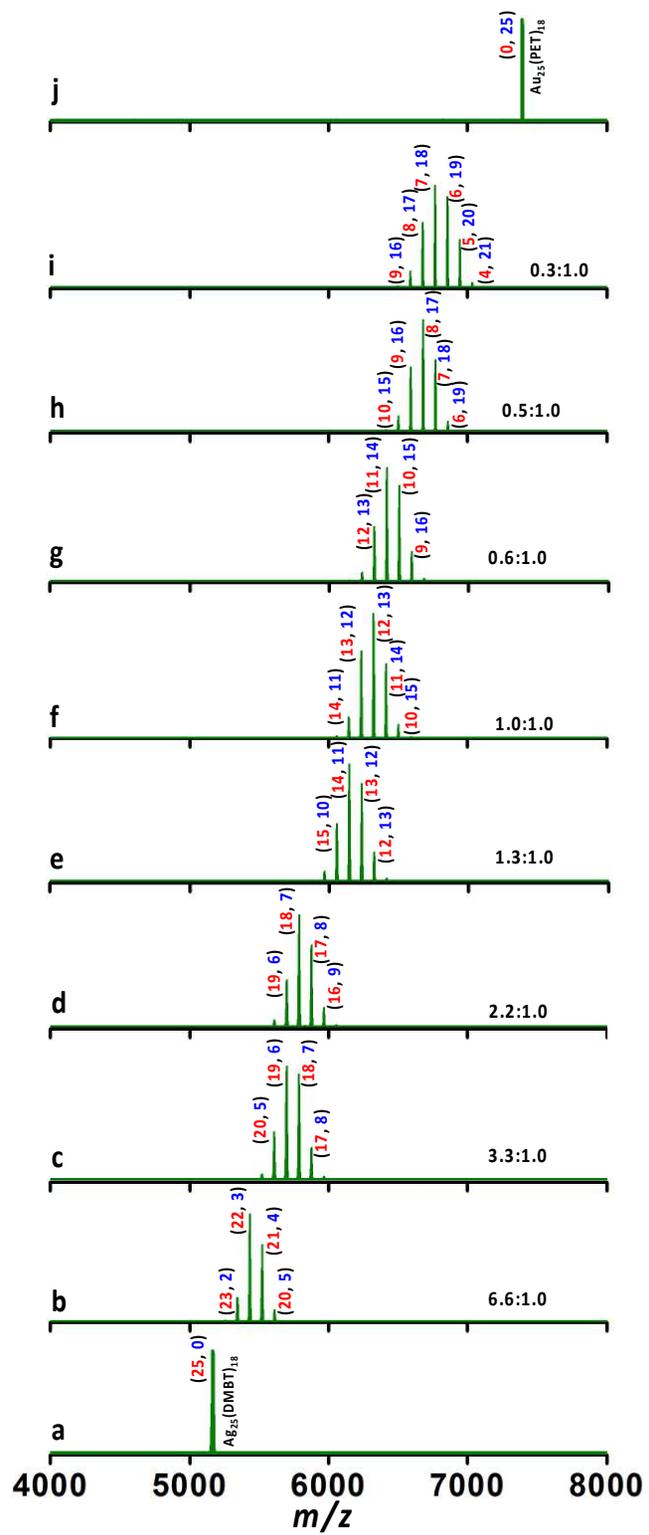


within 5 min

within 2 min

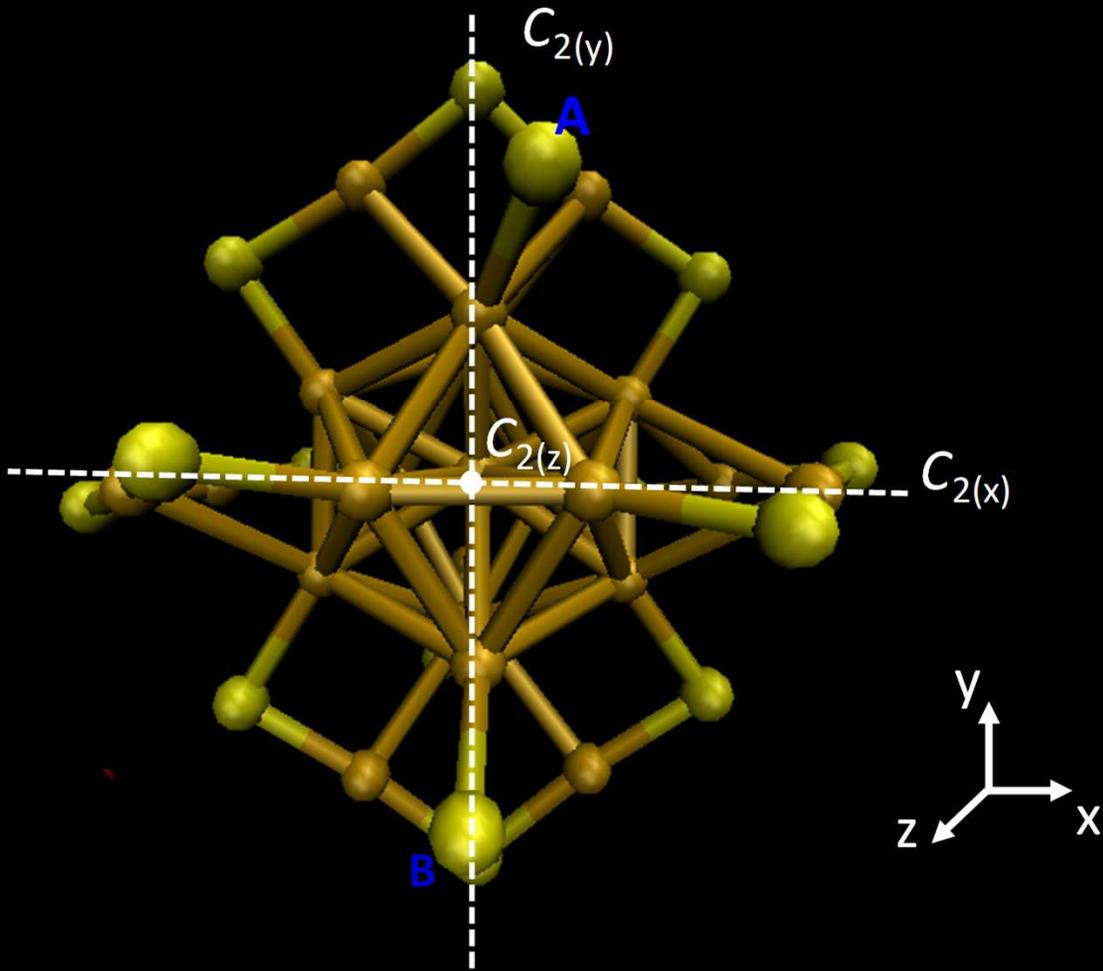
# Optimized structure of $[\text{Ag}_{25}\text{Au}_{25}(\text{DMBT})_{18}(\text{PET})_{18}]^{2-}$



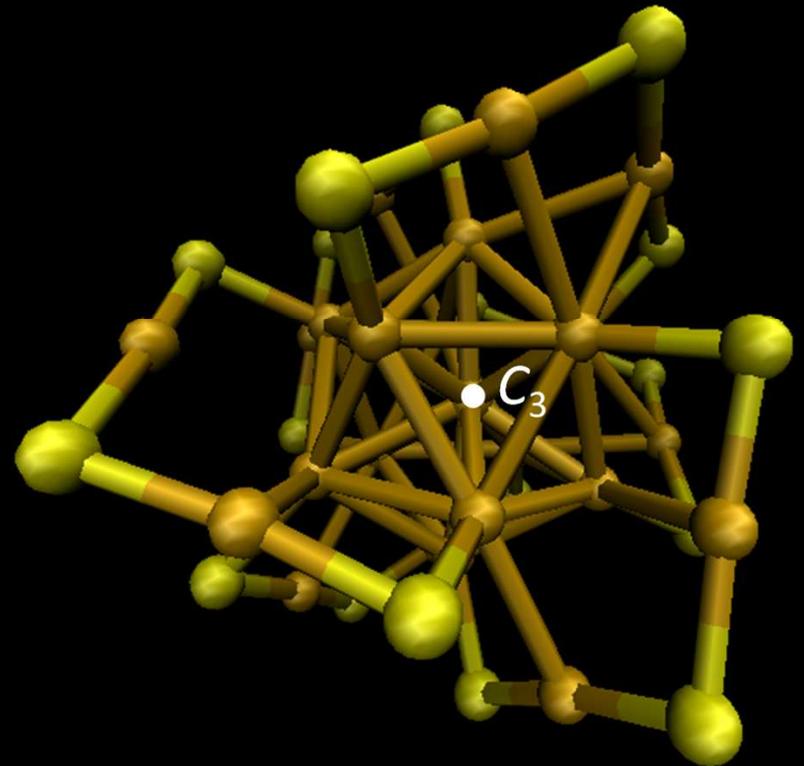


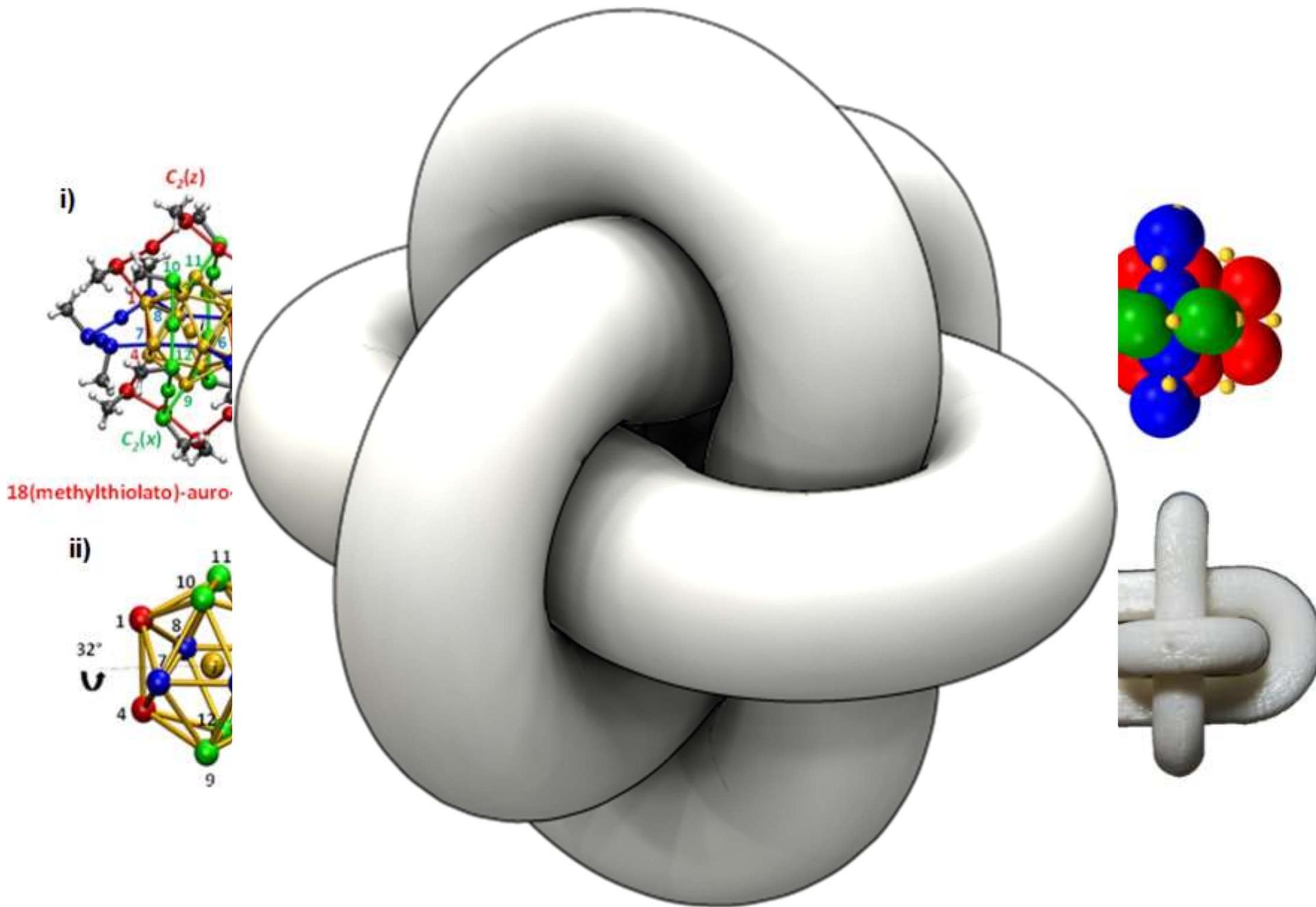
How do we comprehend this?

# 1) Edge projection



# 2) Face Projection





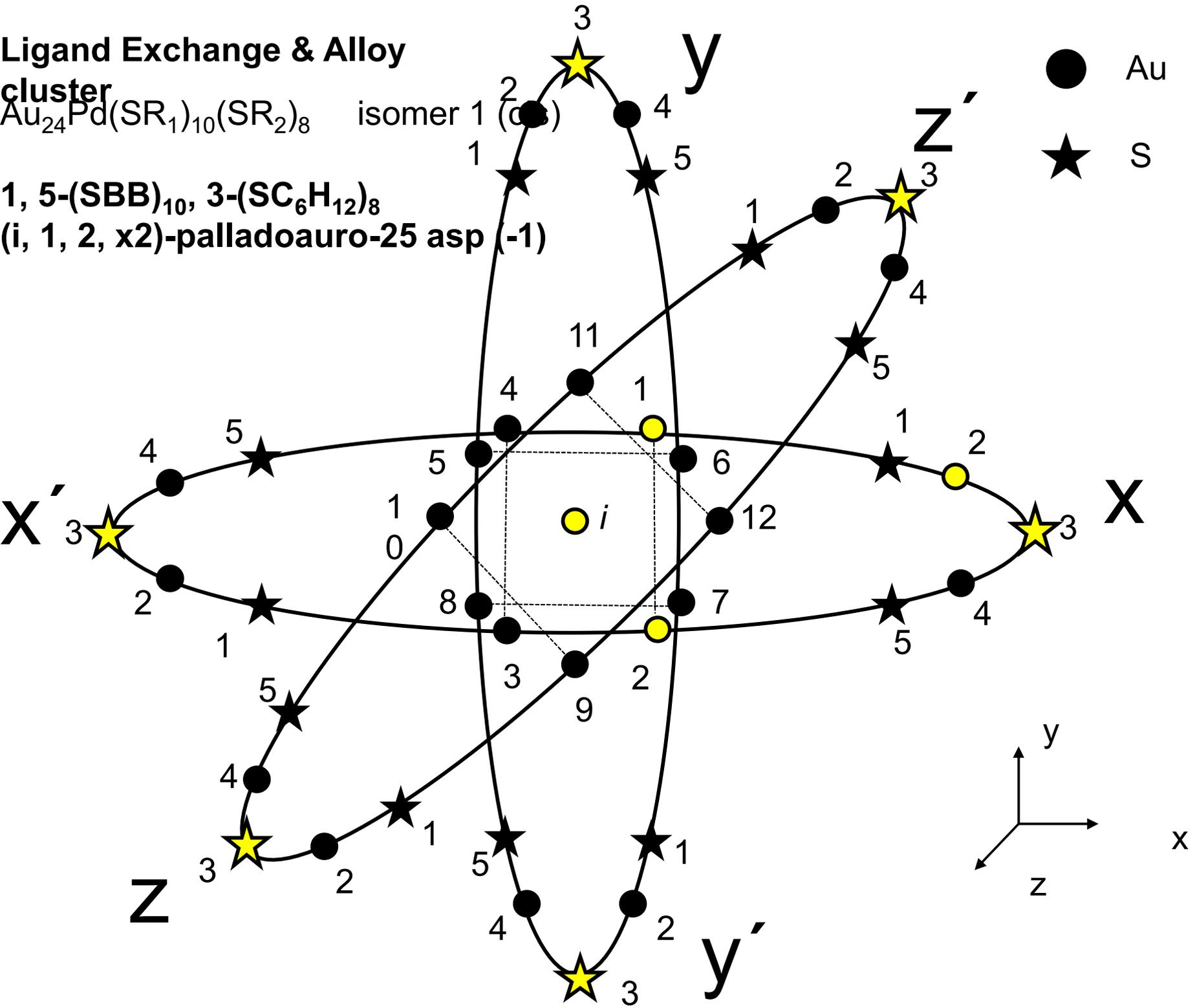
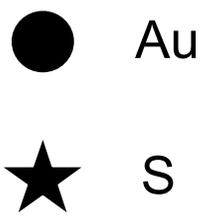
# Aspicules

(D1-3,D2-3)-di(2-phenylethylthiolato), 16(methylthiolato)-auro-25 aspicule(1-)  
(D1-3,D2-3)-(PET)<sub>2</sub>,(SMe)<sub>16</sub>-auro-25 aspicule(1-)

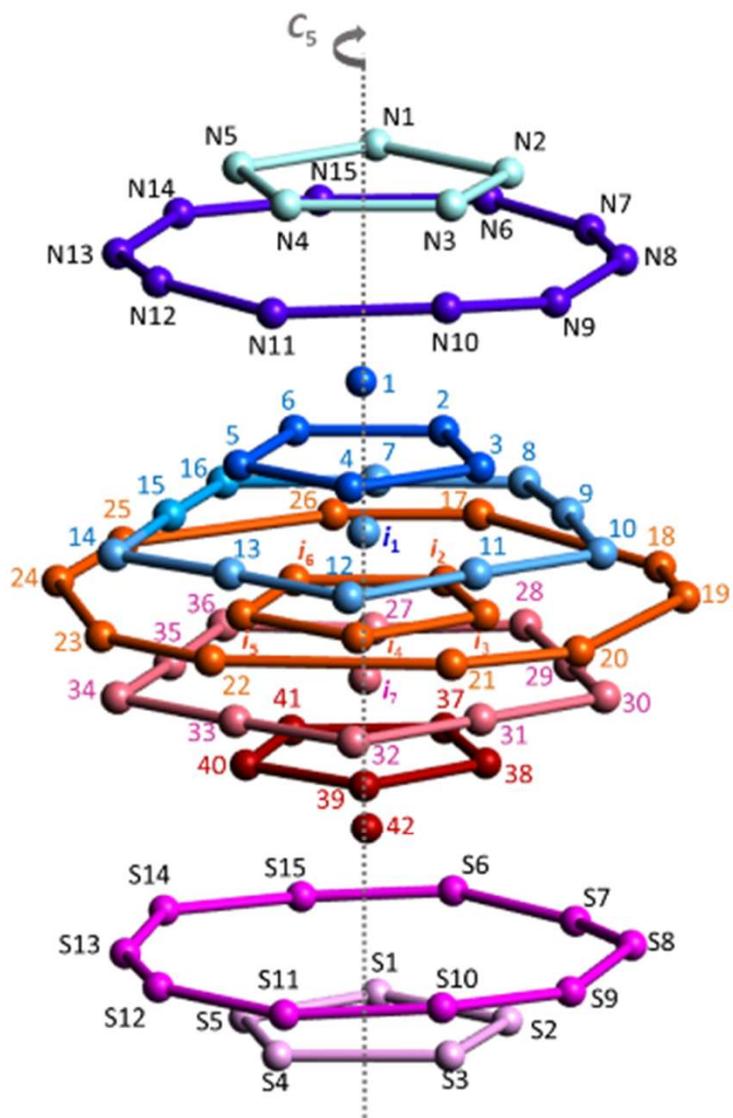
**Ligand Exchange & Alloy**

**cluster**  
 $\text{Au}_{24}\text{Pd}(\text{SR}_1)_{10}(\text{SR}_2)_8$  isomer 1 (c)

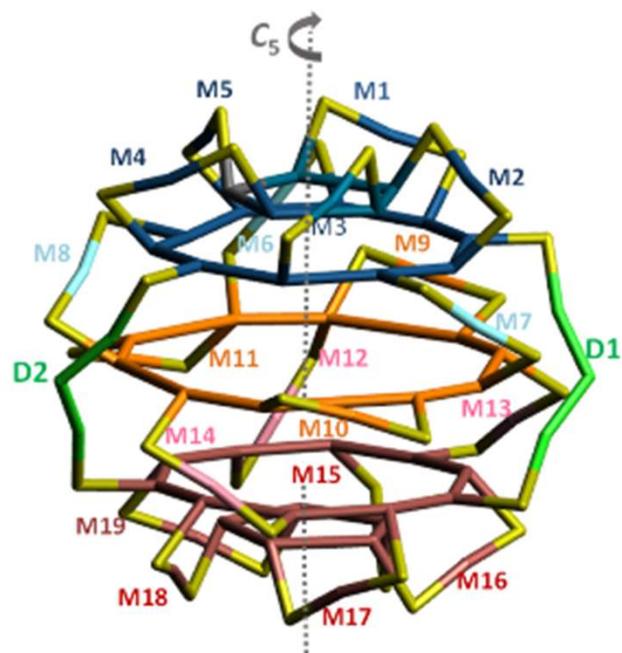
**1, 5-(SBB)<sub>10</sub>, 3-(SC<sub>6</sub>H<sub>12</sub>)<sub>8</sub>**  
**(i, 1, 2, x2)-palladoauro-25 asp (-1)**



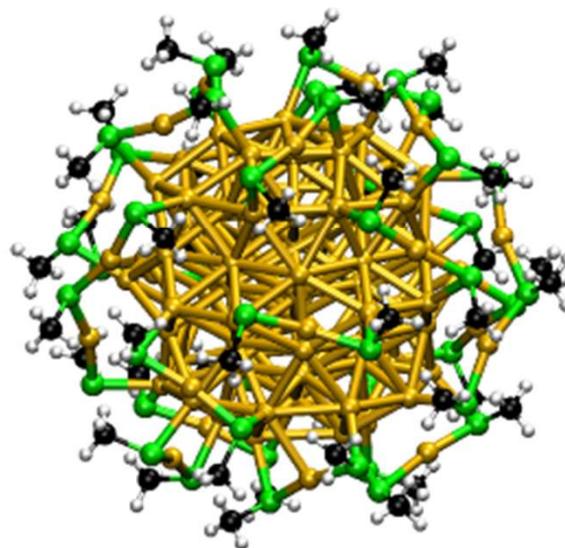
(A)



(B)



(C)



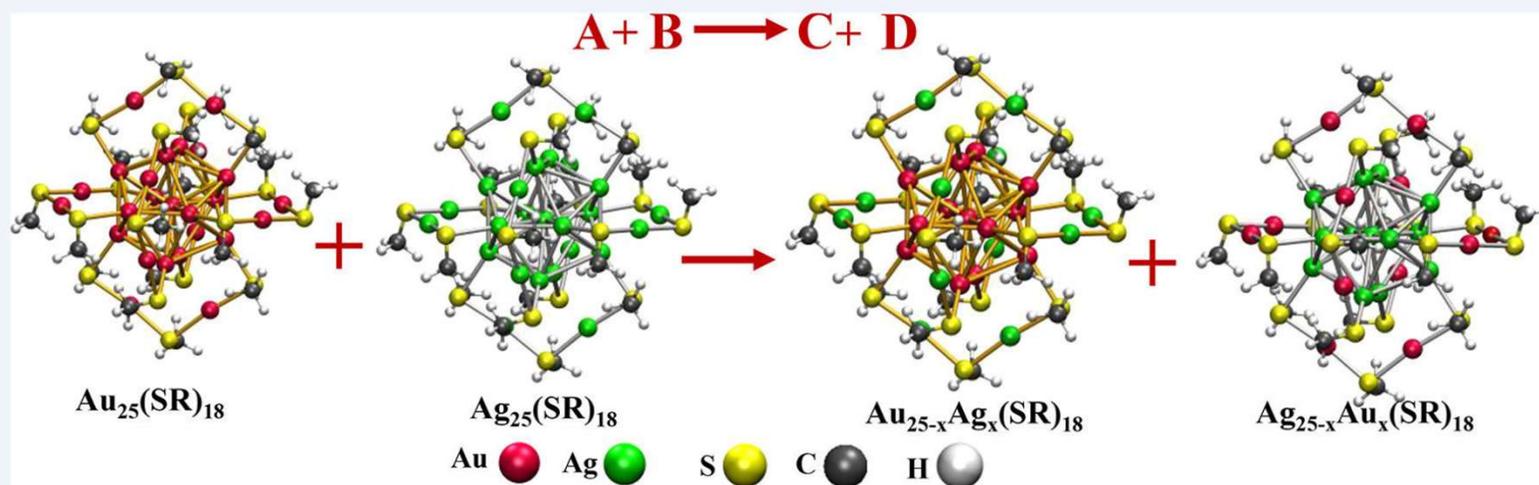
R-44(methylthiolato)-auro-102 aspicule(0)

R-(SMe)<sub>44</sub>-auro-102 aspicule(0) and L-(SMe)<sub>44</sub>-auro-102 aspicule(0) <sup>29</sup>

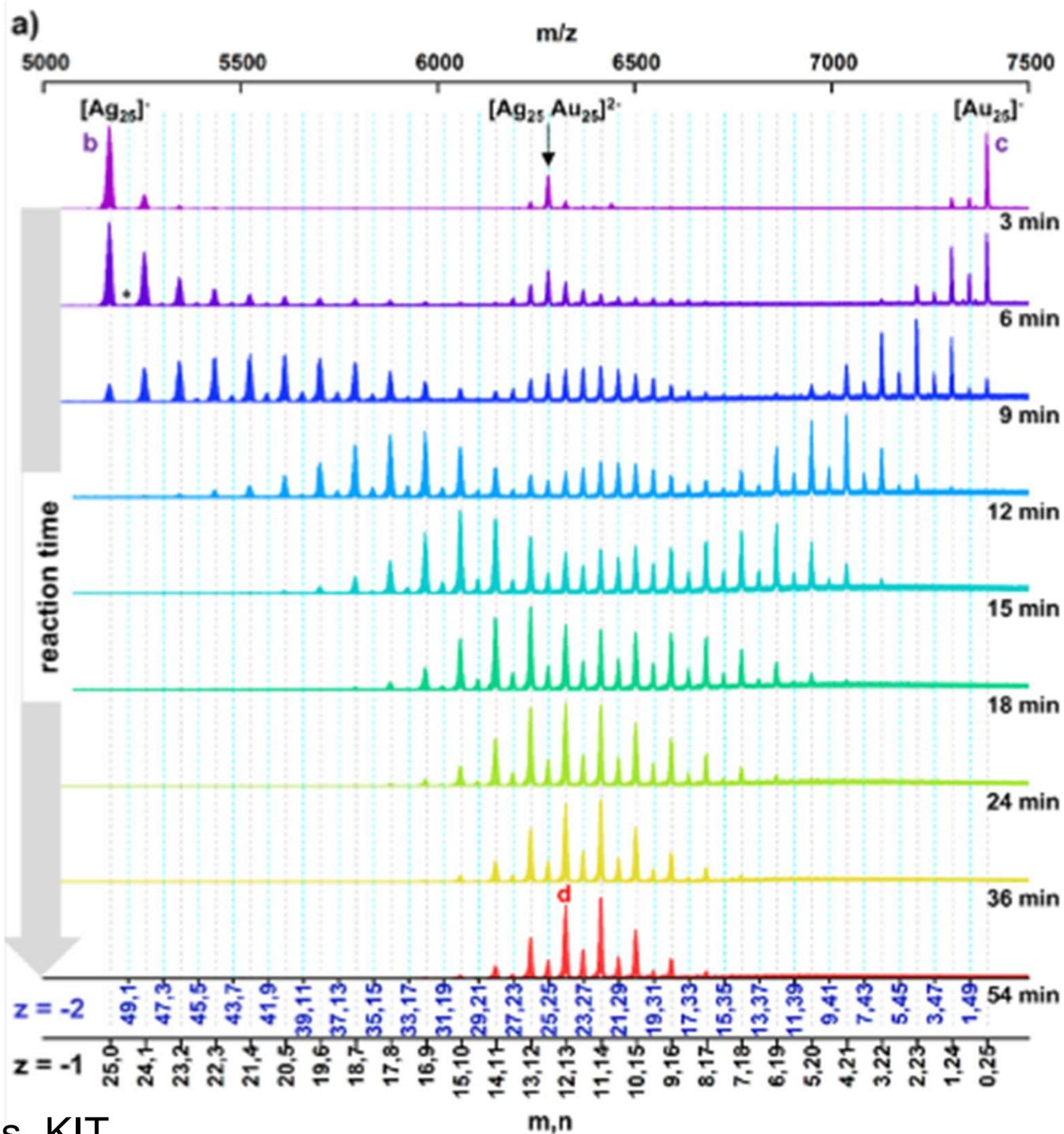
## Interparticle Reactions: An Emerging Direction in Nanomaterials Chemistry

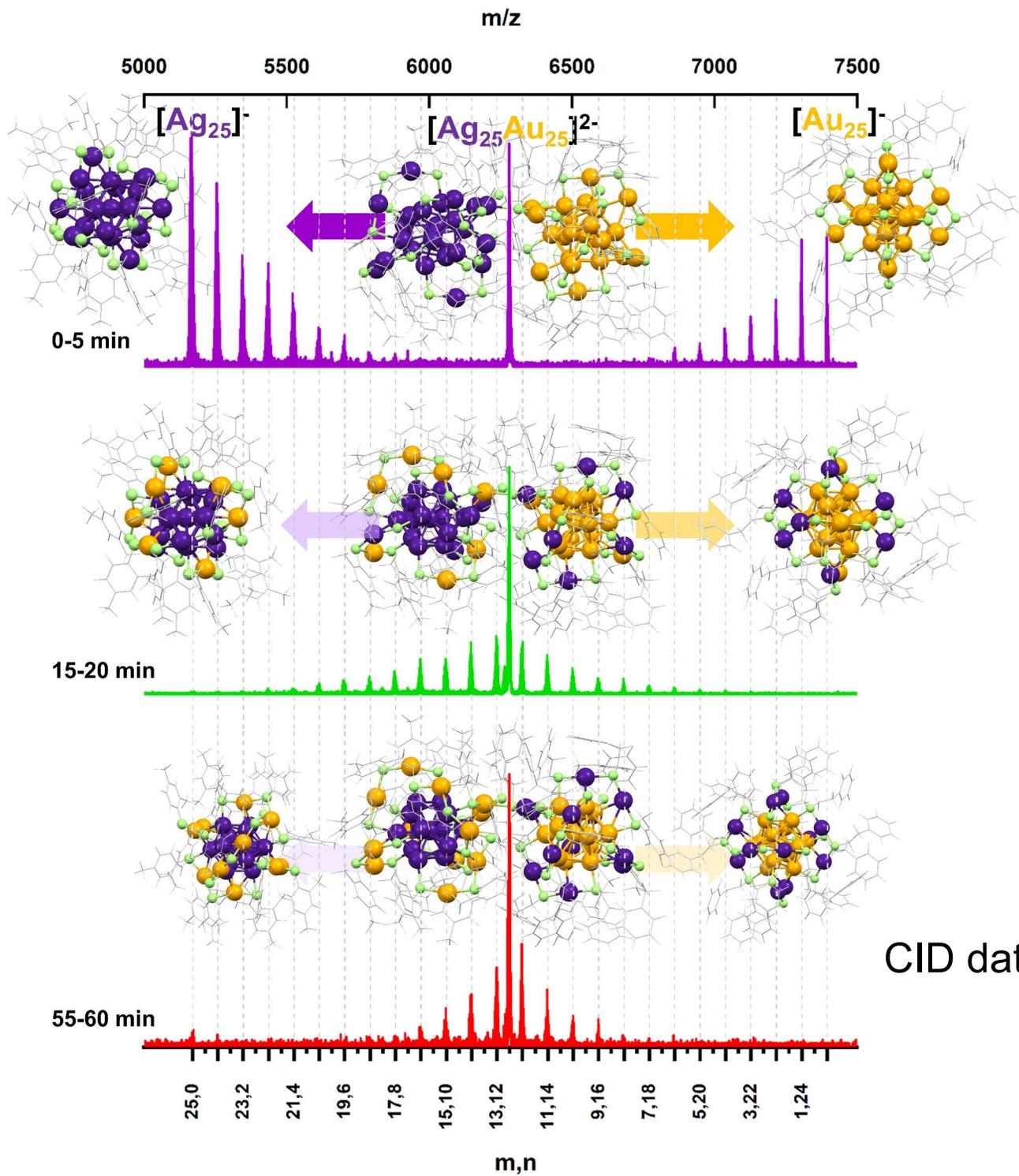
K. R. Krishnadas, Ananya Bakshi,<sup>†</sup> Atanu Ghosh, Ganapati Natarajan, Anirban Som, and Thalappil Pradeep\*<sup>ID</sup>

Department of Chemistry, DST Unit of Nanoscience (DST UNS) and Thematic Unit of Excellence (TUE) Indian Institute of Technology Madras, Chennai 600 036, India



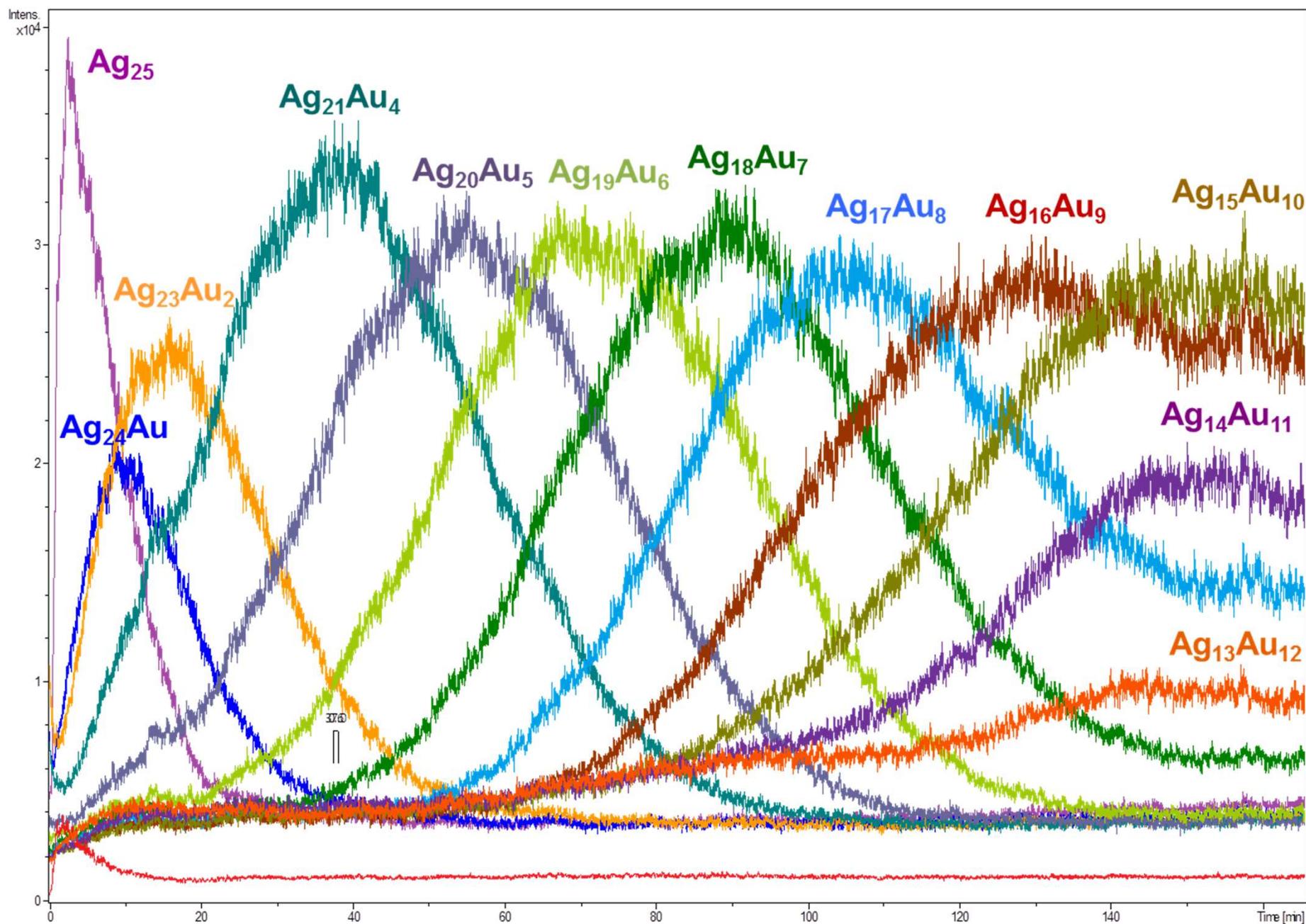
**CONSPECTUS:** Nanoparticles exhibit a rich variety in terms of structure, composition, and properties. However, reactions between them remain largely unexplored. In this *Account*, we discuss an emerging aspect of nanomaterials chemistry, namely, interparticle reactions in solution phase, similar to reactions between molecules, involving atomically precise noble metal clusters.





CID data of [Au<sub>25</sub>Ag<sub>25</sub>]<sup>2-</sup>

# Kinetics of the exchange (monitored on the $\text{Ag}_{25}$ side)



# Expanding reactions

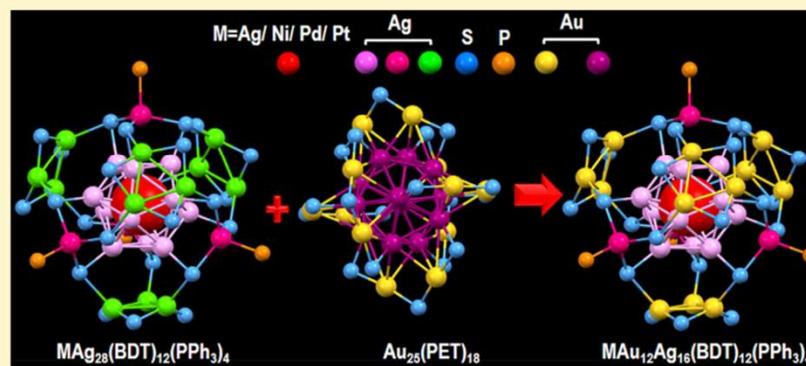
## Intercluster Reactions Resulting in Silver-Rich Trimetallic Nanoclusters

Esma Khatun, Papri Chakraborty, Betsy Rachel Jacob, Ganesan Paramasivam, Mohammad Bodiuzzaman, Wakeel Ahmed Dar, and Thalappil Pradeep\*<sup>1b</sup>

Department of Chemistry, DST Unit of Nanoscience (DST UNS) and Thematic Unit of Excellence (TUE), Indian Institute of Technology Madras, Chennai 600036, India

### S Supporting Information

**ABSTRACT:** Herein, we present an intercluster reaction leading to new trimetallic nanoclusters (NCs) using bimetallic and monometallic NCs as reactants. Dithiol protected bimetallic  $\text{MAg}_{28}(\text{BDT})_{12}(\text{PPh}_3)_4$  (BDT = 1,3-benzenedithiol and  $M = \text{Ni}, \text{Pd},$  or  $\text{Pt}$ ) and monothiol protected  $\text{Au}_{25}(\text{PET})_{18}$  (PET = 2-phenylethanethiol) were used as model NCs. A mixture of trimetallic  $\text{MAu}_x\text{Ag}_{28-x}(\text{BDT})_{12}(\text{PPh}_3)_4$  ( $x = 1-12$ ) and bimetallic  $\text{Ag}_x\text{Au}_{25-x}(\text{PET})_{18}$  ( $x = 1-7$ ) NCs were formed during the reaction as understood from time-dependent electrospray ionization mass spectrometry (ESI MS). Detailed studies of intercluster reaction between  $\text{Ag}_{29}(\text{BDT})_{12}(\text{PPh}_3)_4$  and  $\text{Au}_{25}(\text{PET})_{18}$  were also performed. Although both  $\text{MAg}_{28}(\text{BDT})_{12}(\text{PPh}_3)_4$  ( $M = \text{Ag}, \text{Ni}, \text{Pd},$  or  $\text{Pt}$ ) and  $\text{Au}_{25}(\text{PET})_{18}$  contain 13 atoms icosahedral core, only a maximum of 12 Au doped NCs were formed for the former as a major product and not the 13 Au doped one, unlike the previous reports of intercluster reaction. The transfer of Ni, Pd, or Pt atom from the center of icosahedron of  $\text{MAg}_{28}(\text{BDT})_{12}(\text{PPh}_3)_4$  to  $\text{Au}_{25}(\text{PET})_{18}$  was not observed, which suggests that the central atom is not involved in the reaction. Density functional theory (DFT) calculations were performed to know structures



# Interparticle Reactions Forming Product Co-Crystals

ACS NANO

Cite This: *ACS Nano* 2019, 13, 13365–13373

www.acsnano.org

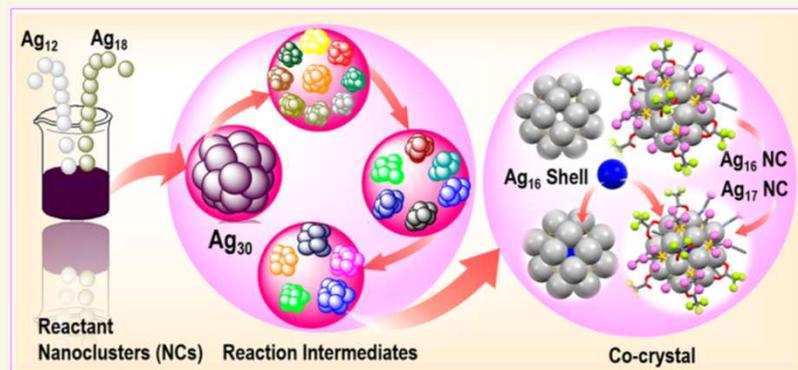
## Interparticle Reactions between Silver Nanoclusters Leading to Product Cocystals by Selective Cocrystallization

Wakeel Ahmed Dar,<sup>†</sup> Mohammad Bodiuzzaman,<sup>†</sup> Debasmita Ghosh, Ganesan Paramasivam, Esma Khatun, Korath Shivan Sugi, and Thalappil Pradeep\*<sup>ID</sup>

Department of Chemistry, DST Unit of Nanoscience and Thematic Unit of Excellence, Indian Institute of Technology Madras, Chennai 600036, India

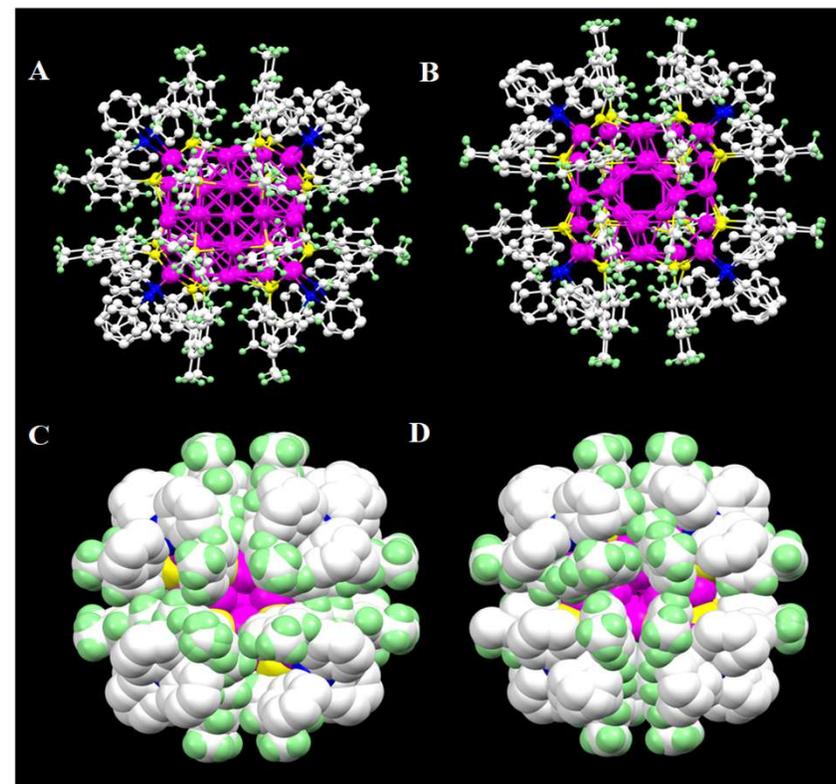
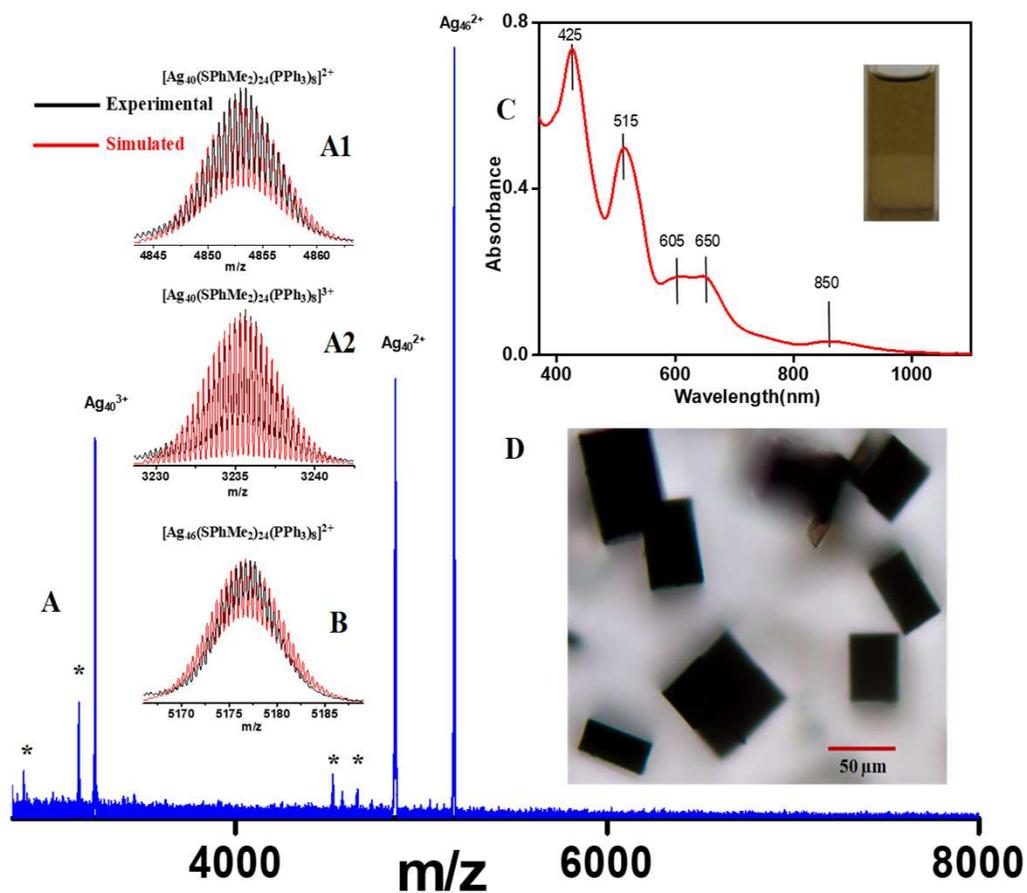
**S** Supporting Information

**ABSTRACT:** We present an example of an interparticle reaction between atomically precise nanoclusters (NCs) of the same metal, resulting in entirely different clusters. In detail, the clusters  $[\text{Ag}_{12}(\text{TBT})_8(\text{TFA})_5(\text{CH}_3\text{CN})]^+$  (TBT = *tert*-butylthiolate, TFA = trifluoroacetate,  $\text{CH}_3\text{CN}$  = acetonitrile) and  $[\text{Ag}_{18}(\text{TPP})_{10}\text{H}_{16}]^{2+}$  (TPP = triphenylphosphine) abbreviated as  $\text{Ag}_{12}$  and  $\text{Ag}_{18}$ , respectively, react leading to  $[\text{Ag}_{16}(\text{TBT})_8(\text{TFA})_7(\text{CH}_3\text{CN})_3\text{Cl}]^+$  and  $[\text{Ag}_{17}(\text{TBT})_8(\text{TFA})_7(\text{CH}_3\text{CN})_3\text{Cl}]^+$ , abbreviated as  $\text{Ag}_{16}$  and  $\text{Ag}_{17}$ , respectively. The two product NCs crystallize together as both possess the same metal chalcogenolate



# Co-crystals

## Ag<sub>40</sub> and Ag<sub>46</sub> with the same shell



M. Bodiuzzaman, et. al. *Angew. Chem. Int. Ed.* 2018

# Atom transfer dynamics



They are indeed molecules!

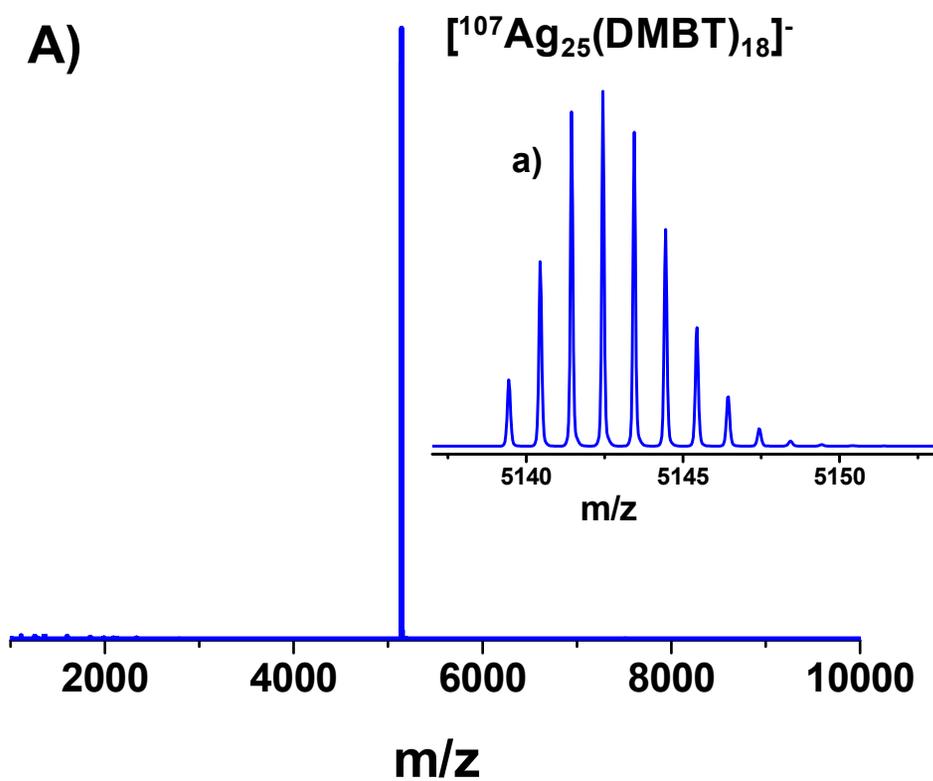
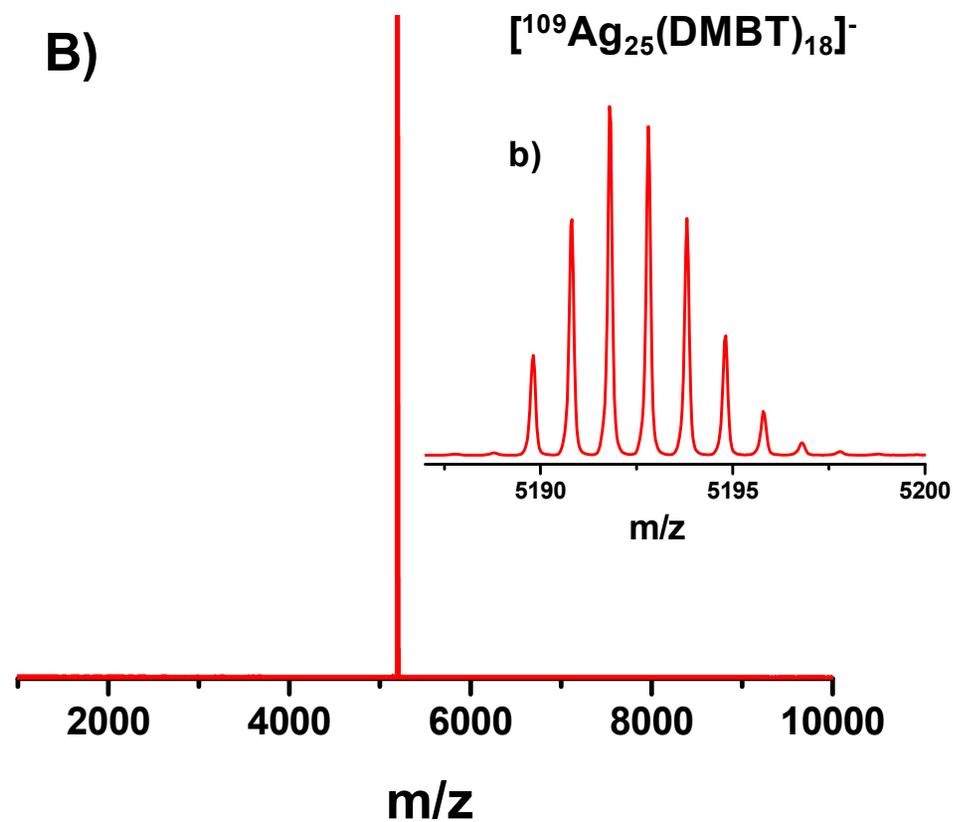
CONDENSED MATTER PHYSICS

# Rapid isotopic exchange in nanoparticles

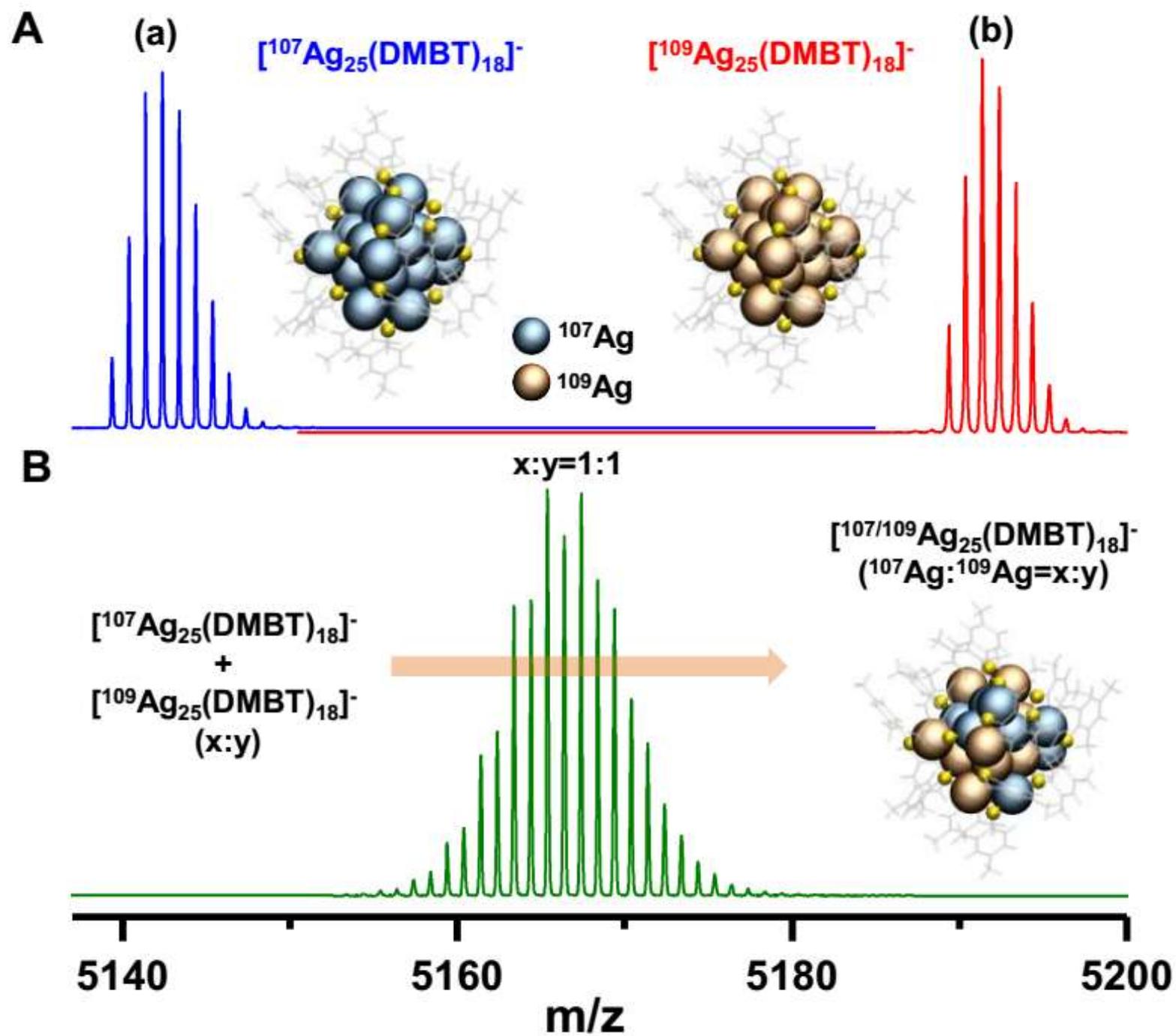
**Papri Chakraborty<sup>1</sup>, Abhijit Nag<sup>1</sup>, Ganapati Natarajan<sup>1</sup>, Nayanika Bandyopadhyay<sup>1</sup>, Ganesan Paramasivam<sup>1</sup>, Manoj Kumar Panwar<sup>1</sup>, Jaydeb Chakrabarti<sup>2</sup>, Thalappil Pradeep<sup>1\*</sup>**

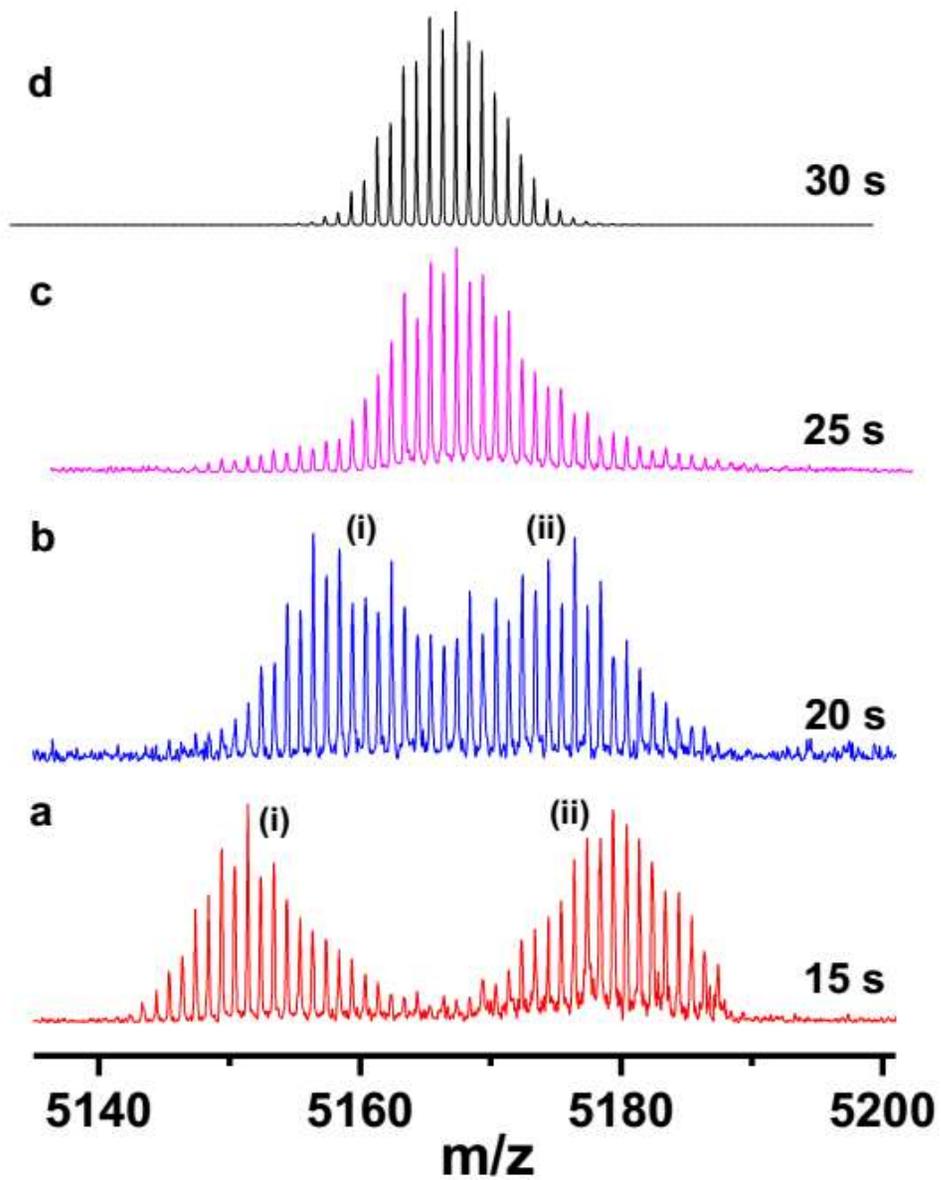
Rapid solution-state exchange dynamics in nanoscale pieces of matter is revealed, taking isotopically pure atomically precise clusters as examples. As two isotopically pure silver clusters made of  $^{107}\text{Ag}$  and  $^{109}\text{Ag}$  are mixed, an isotopically mixed cluster of the same entity results, similar to the formation of HDO, from  $\text{H}_2\text{O}$  and  $\text{D}_2\text{O}$ . This spontaneous process is driven by the entropy of mixing and involves events at multiple time scales.

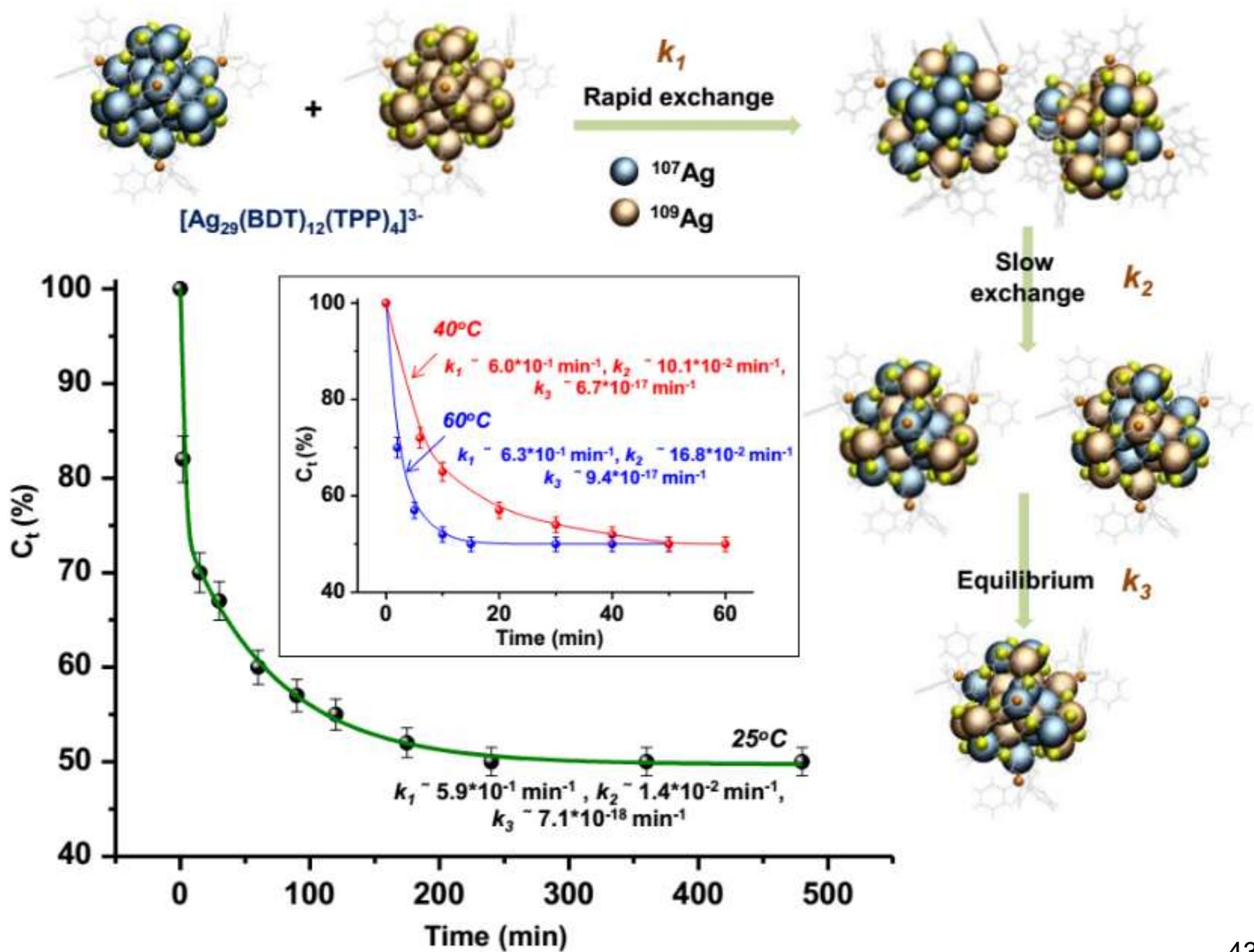
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**A)****B)**

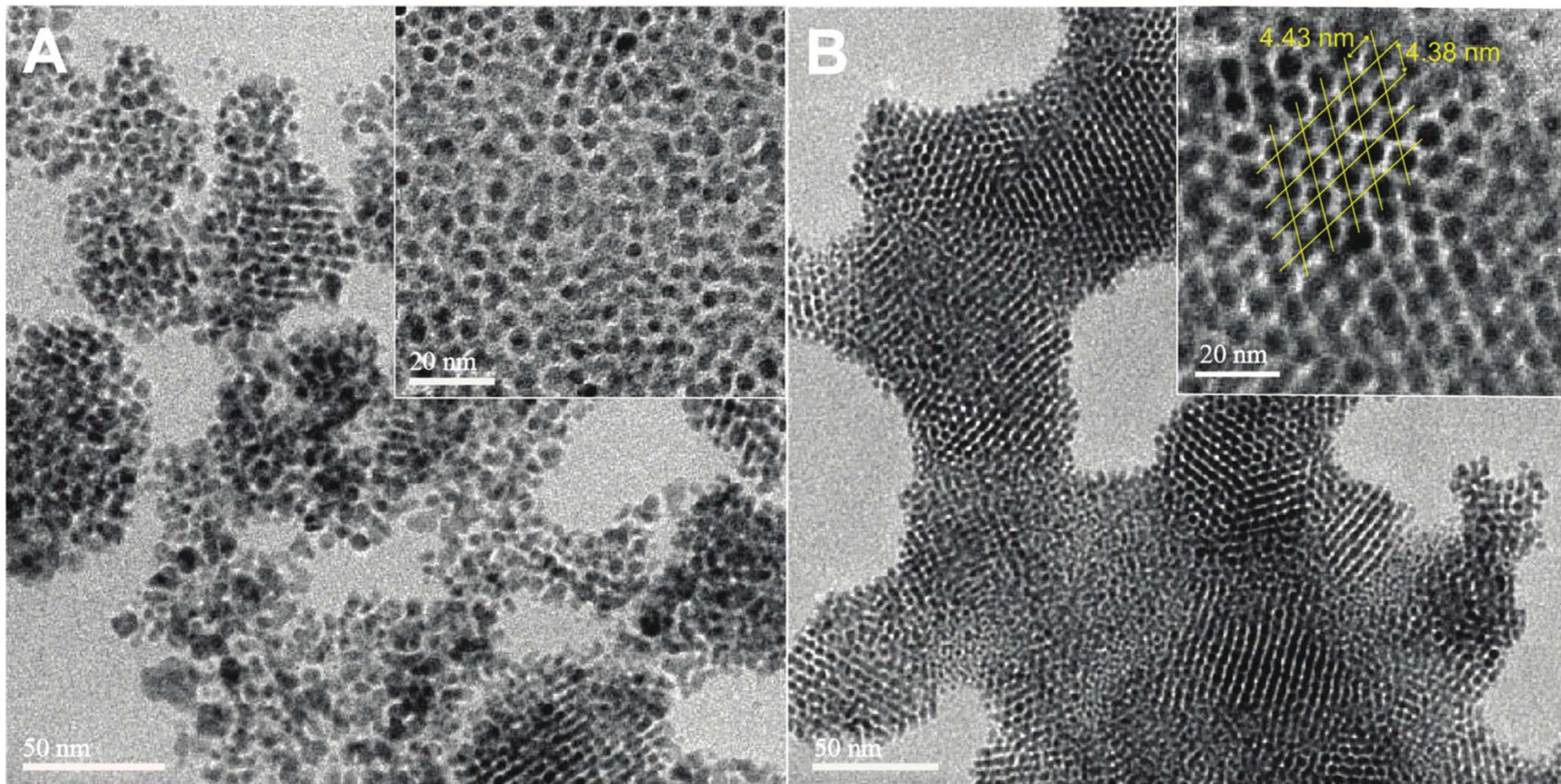
ESI MS of **A)**  $^{107}\text{Ag}_{25}(\text{DMBT})_{18}$  and **B)**  $^{109}\text{Ag}_{25}(\text{DMBT})_{18}$ . Insets shows the respective isotope patterns.

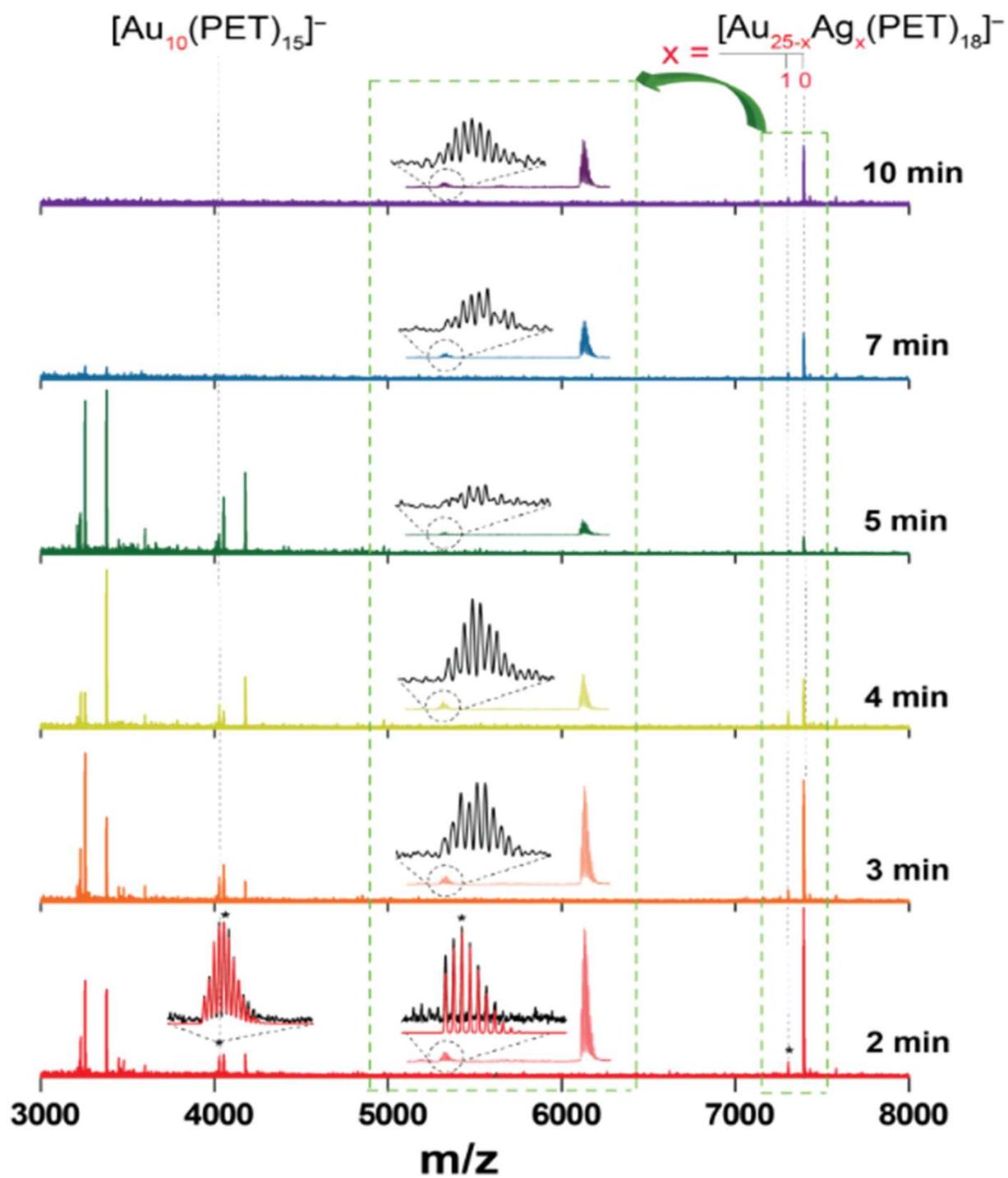


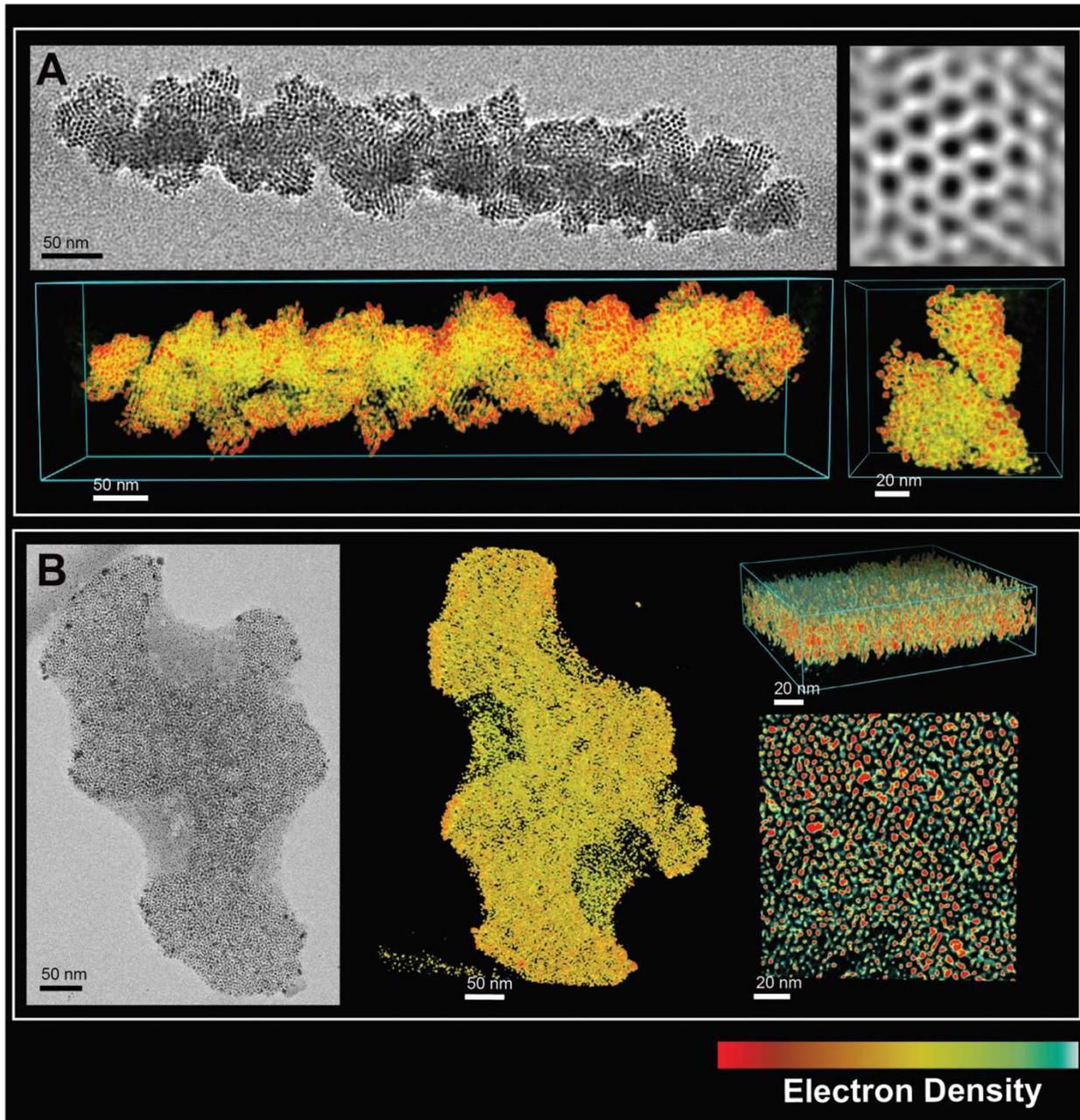




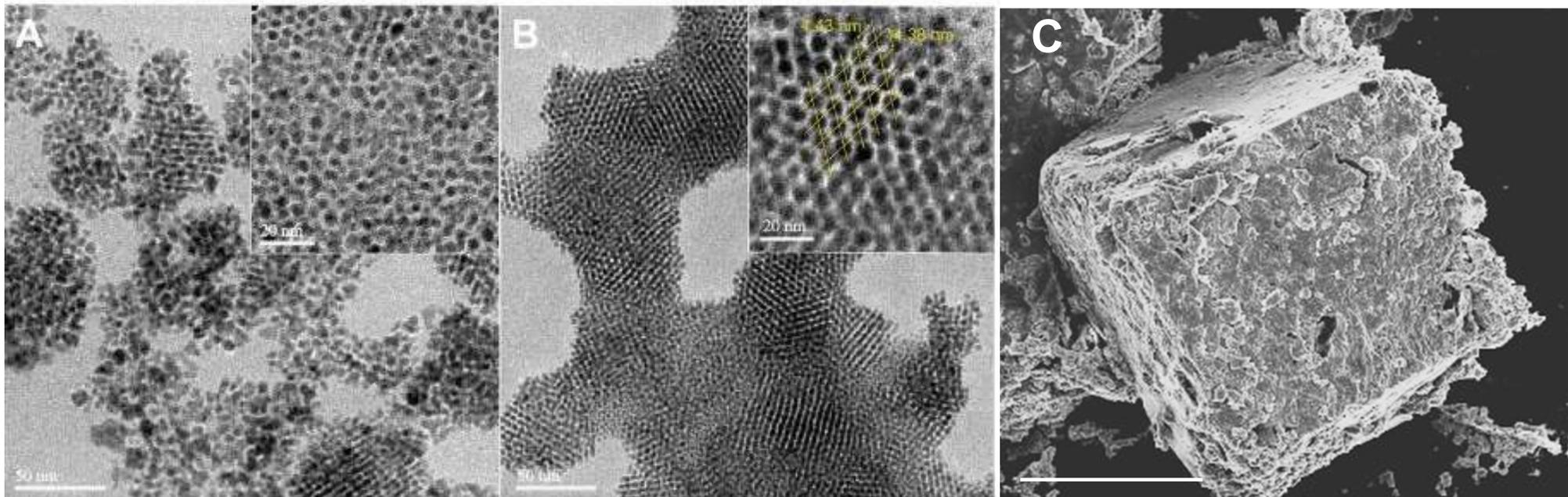
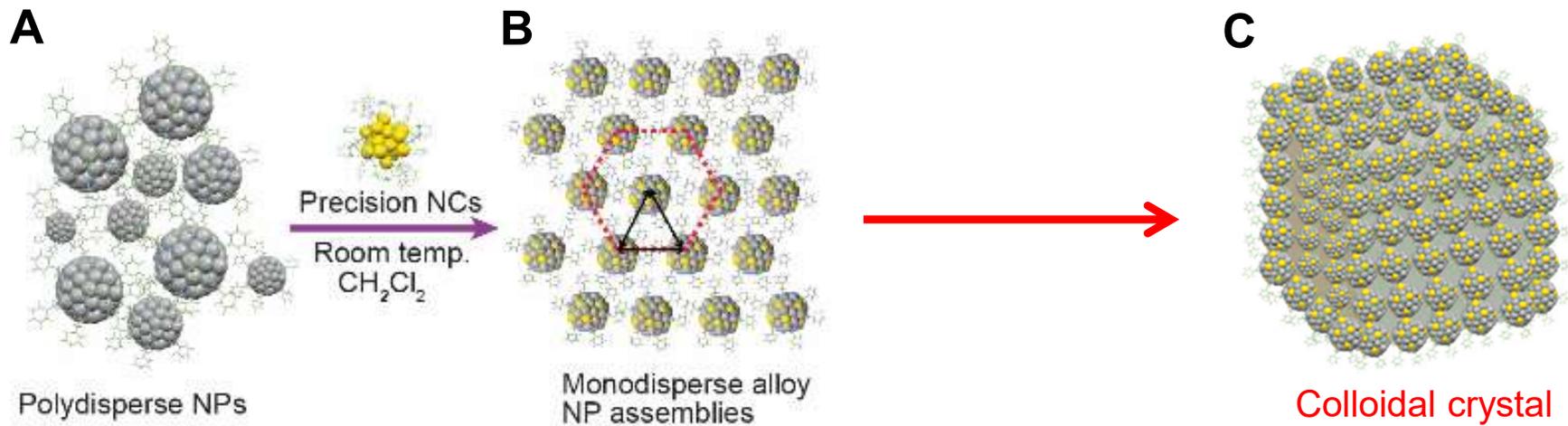
Can clusters react with nanoparticles?



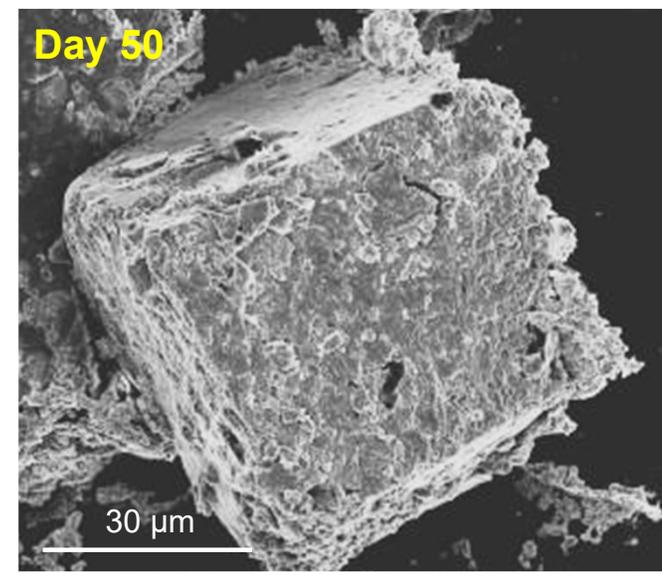
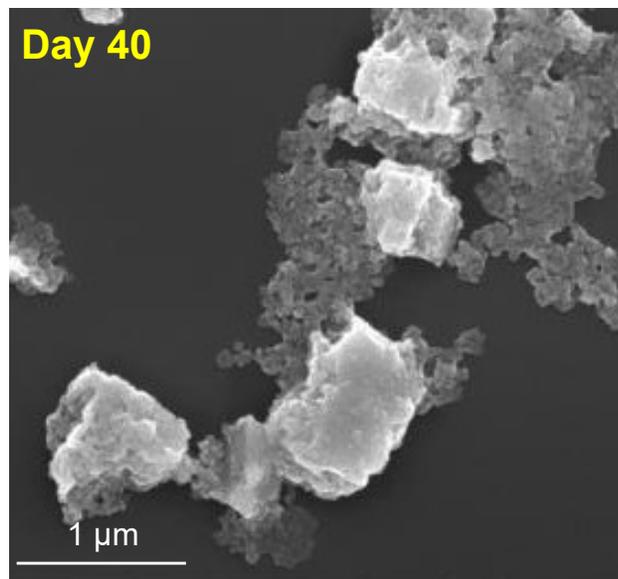
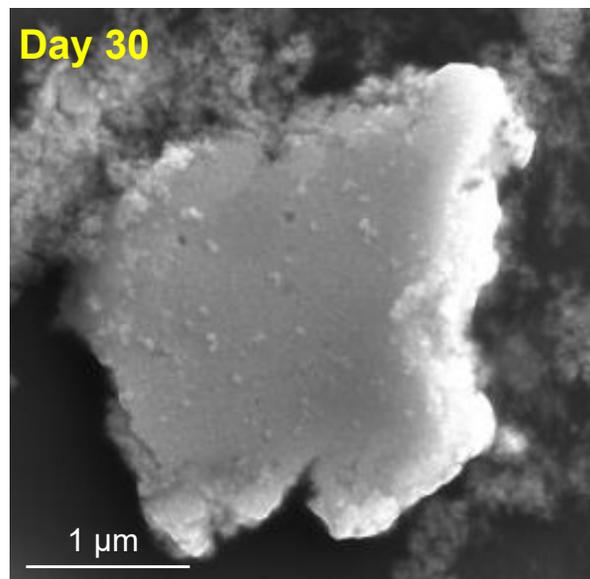
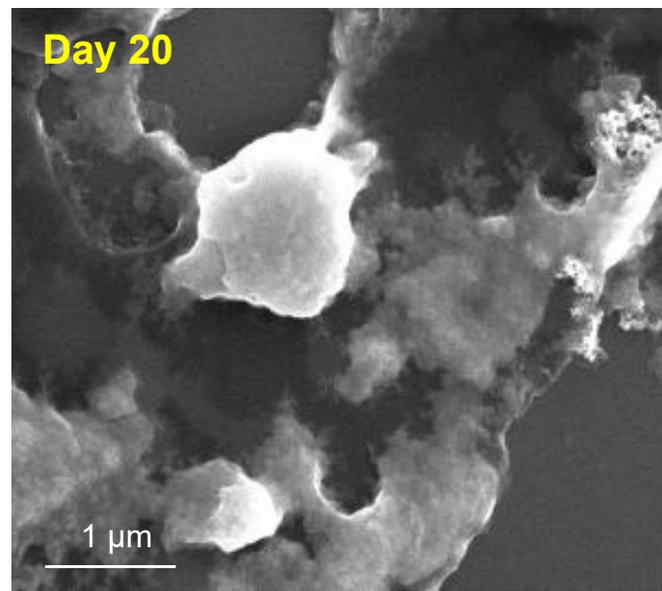
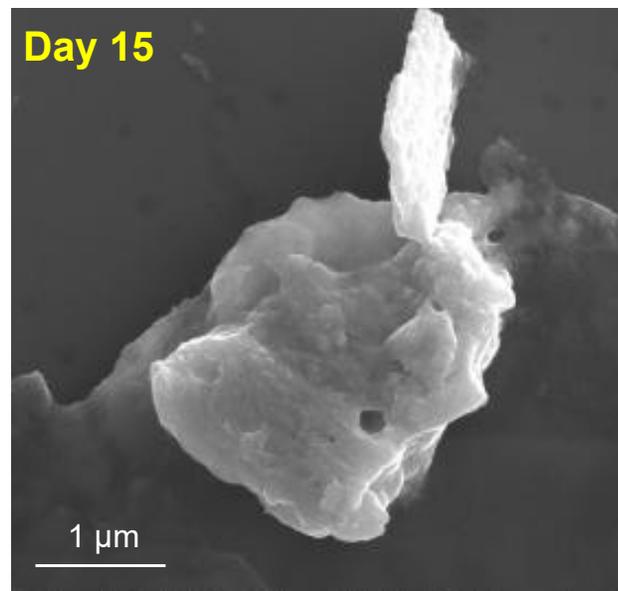
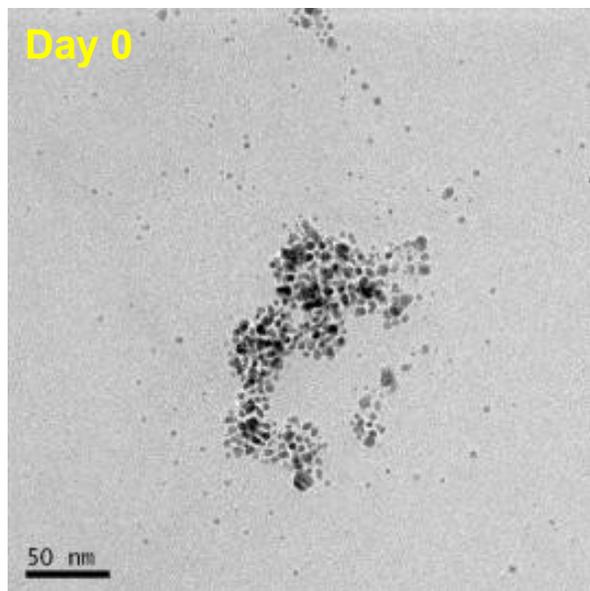




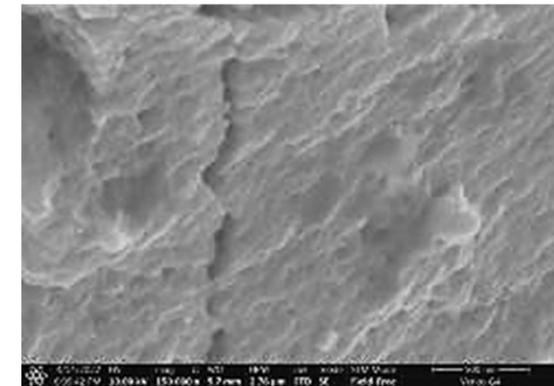
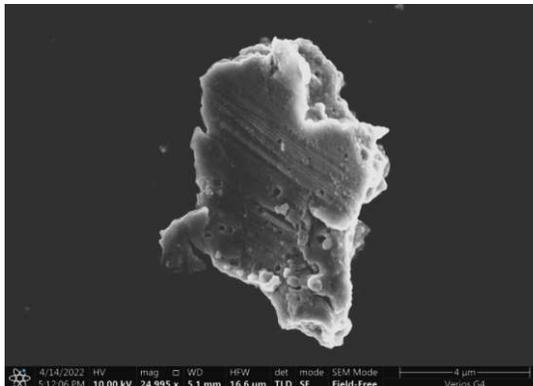
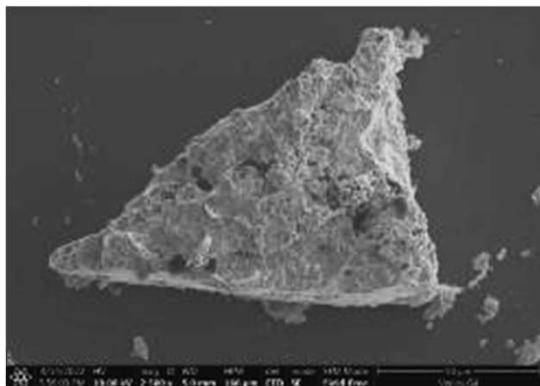
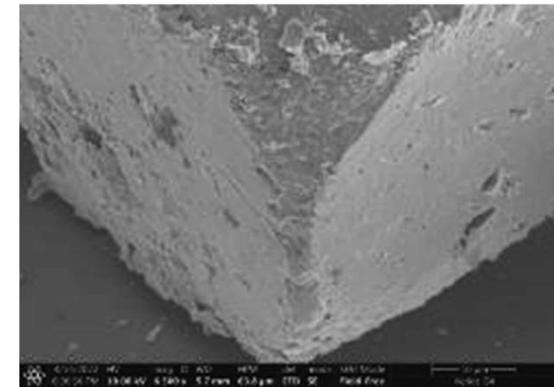
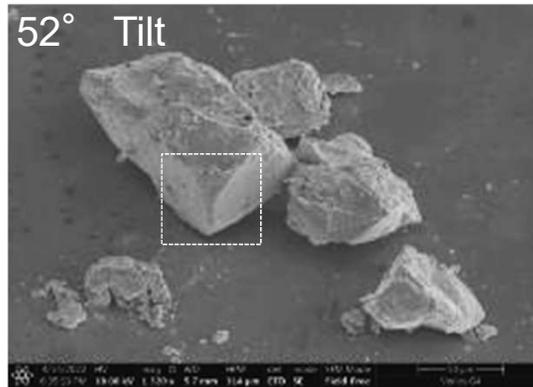
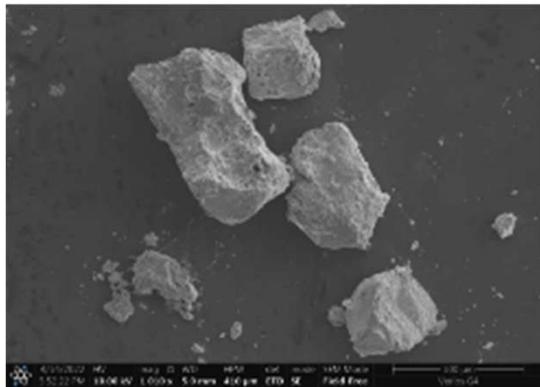
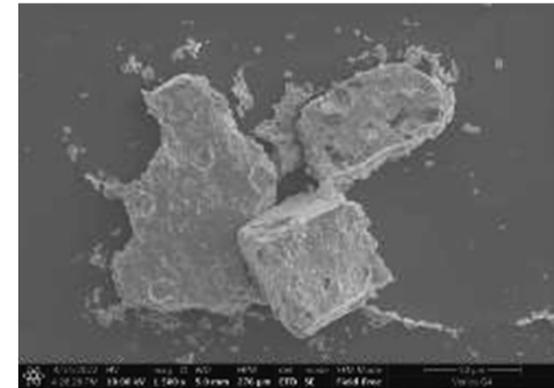
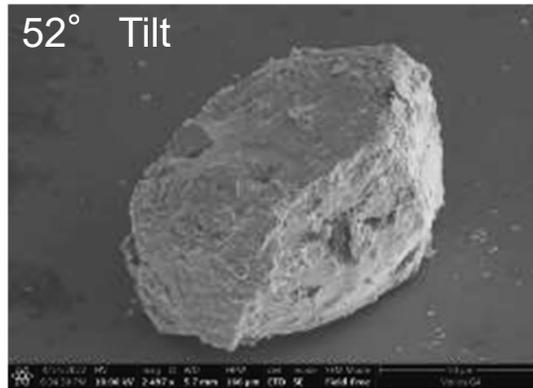
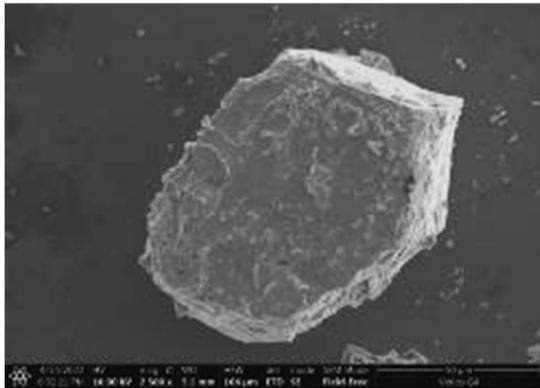
# Crystallization of AgAu NP



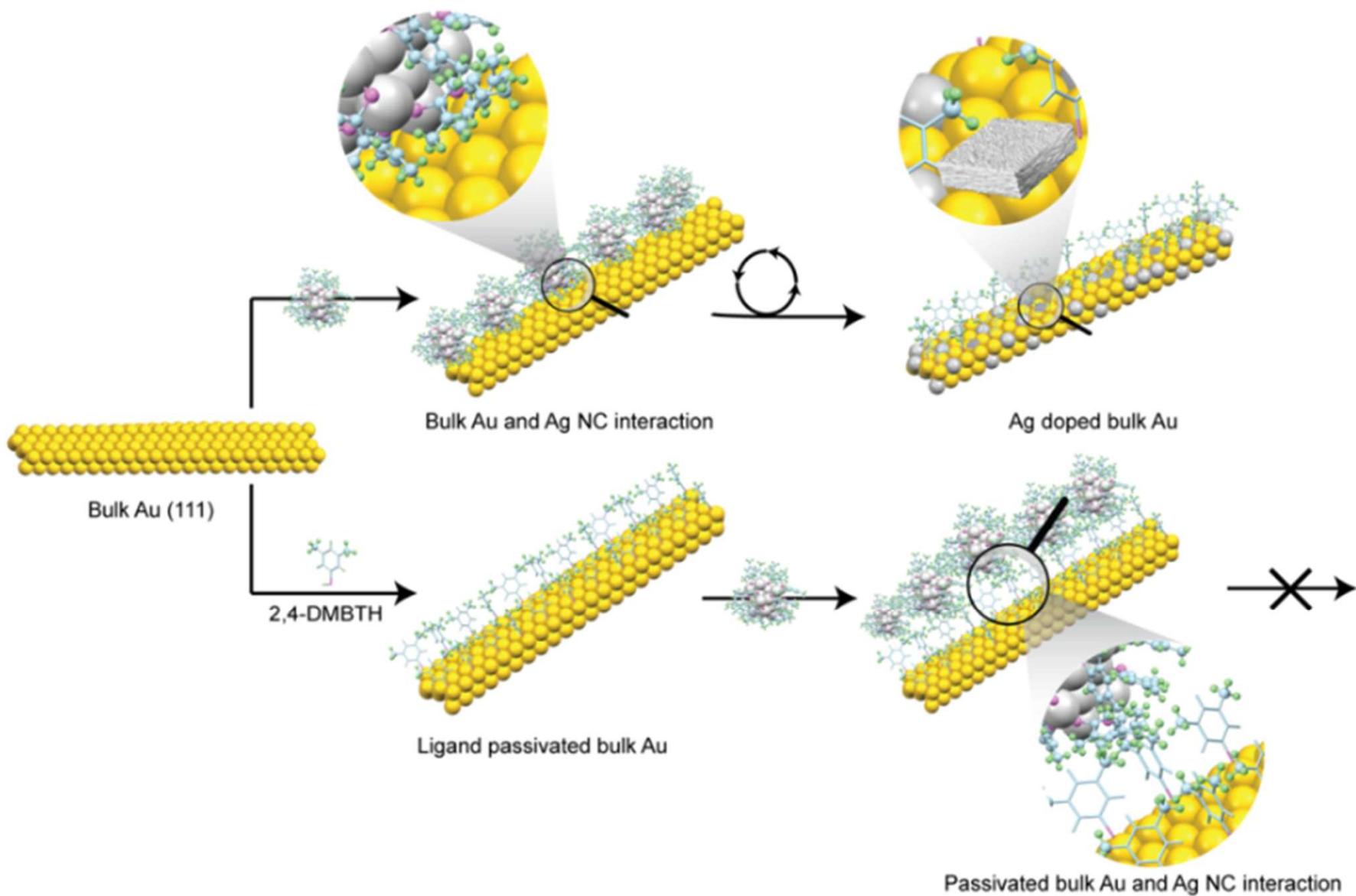
# Transformation of colloidal particles in solution



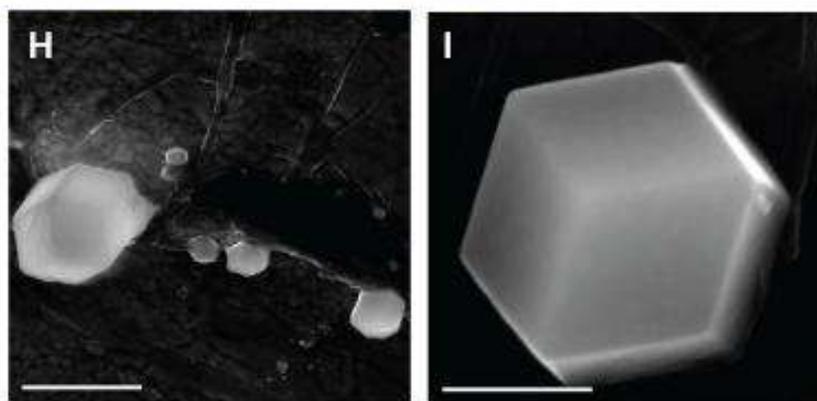
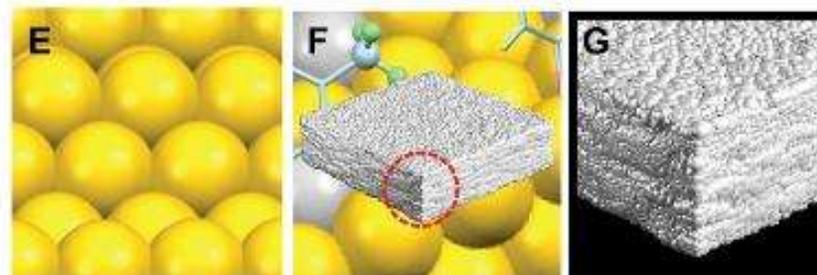
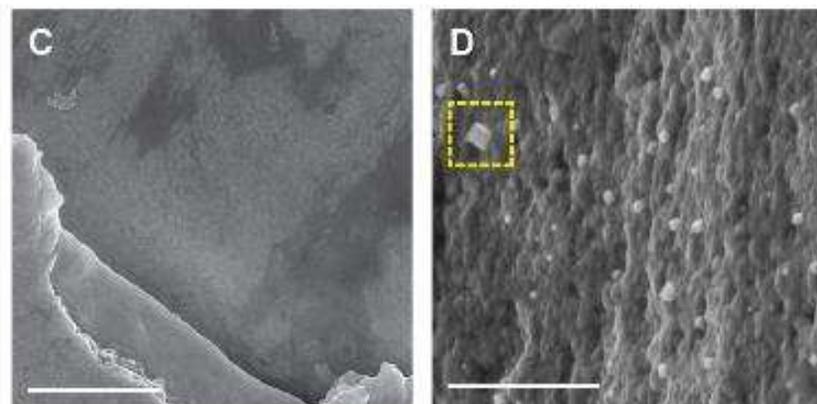
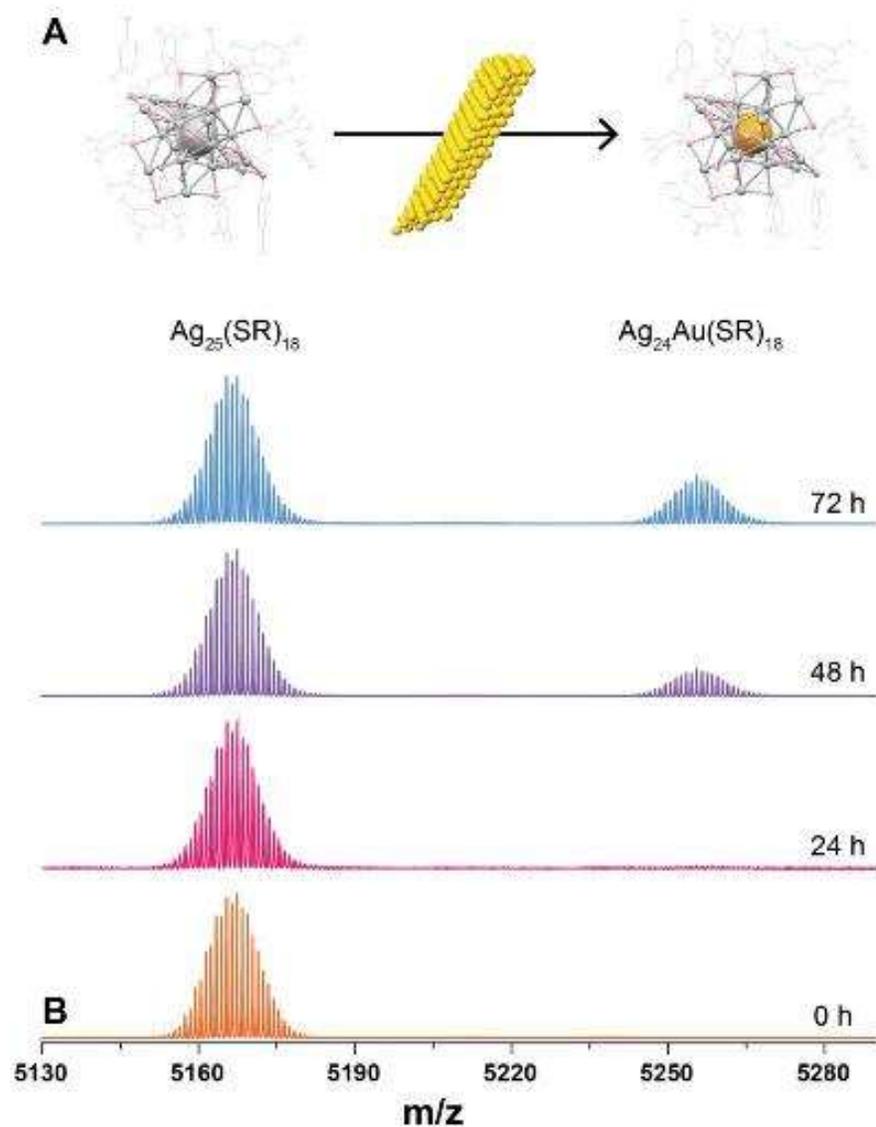
# Crystals imaged after 50 days



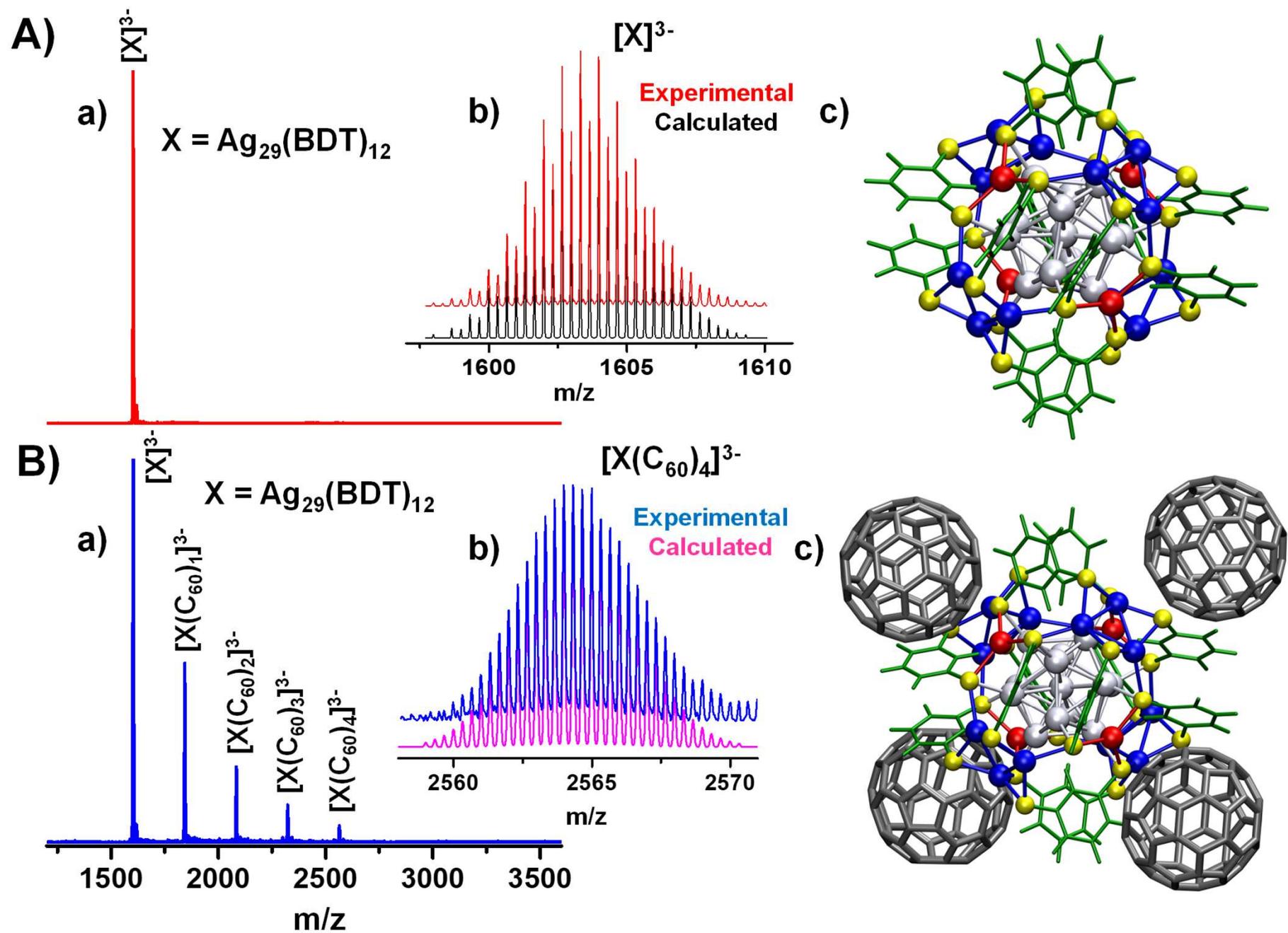
# Bulk Au—Ag NC reaction



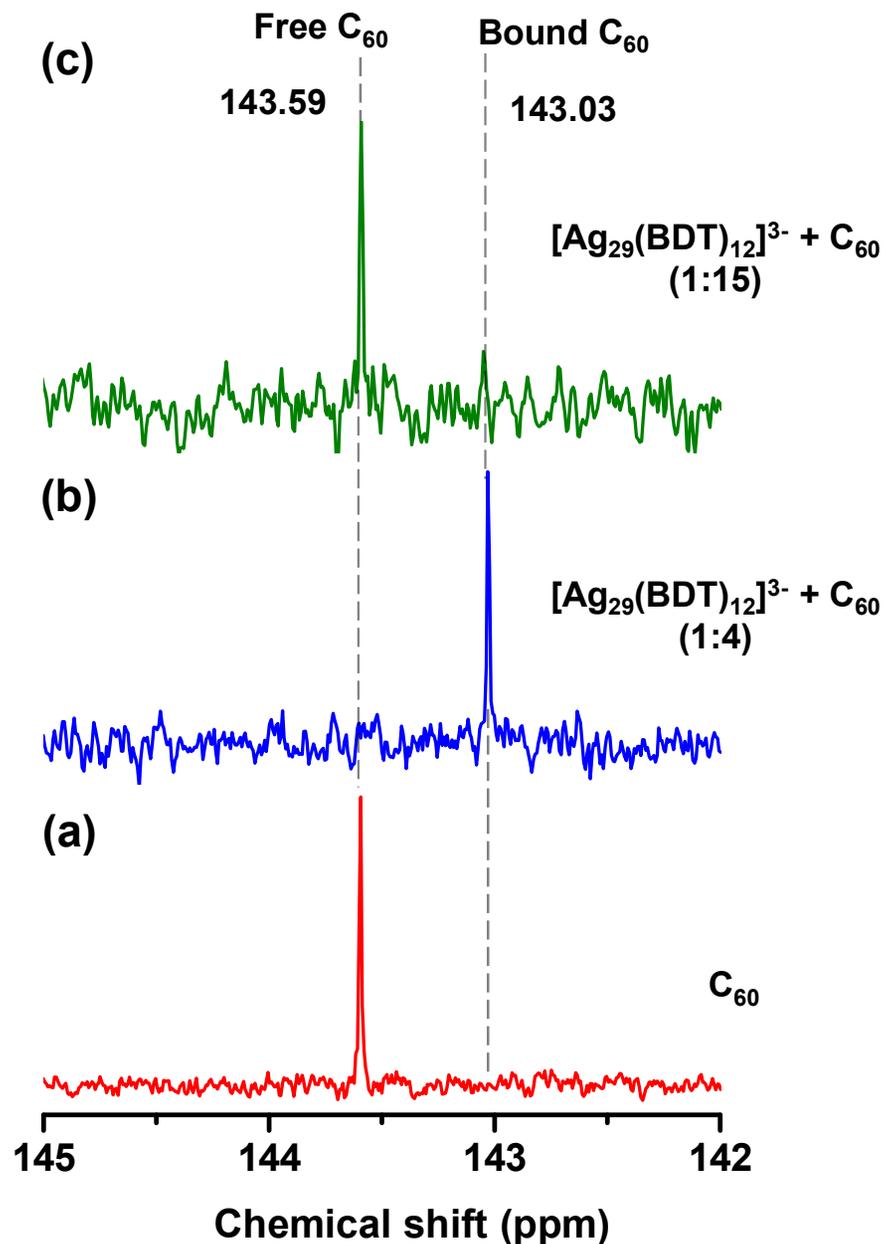
# Characterization of reaction products



# Supramolecular chemistry

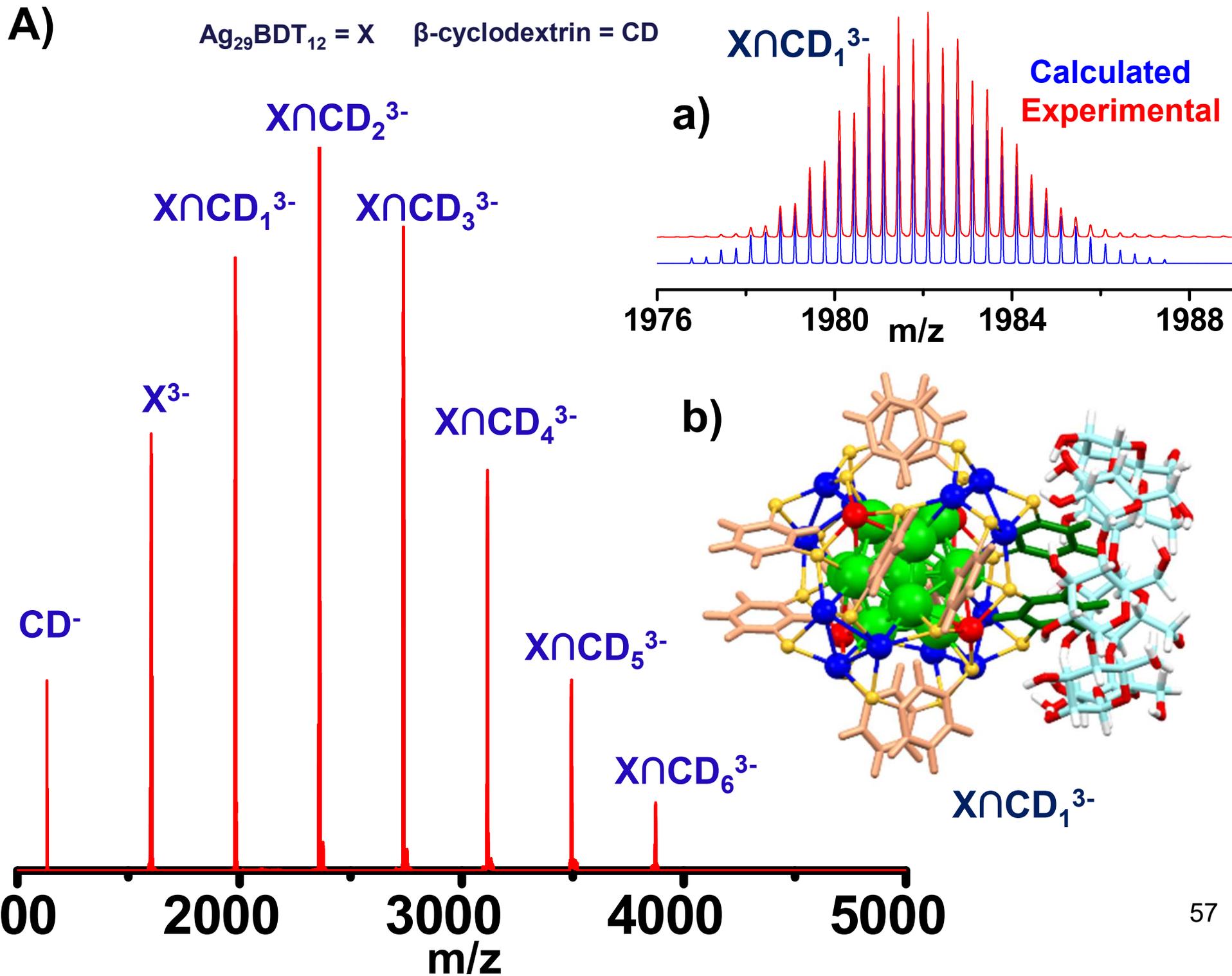


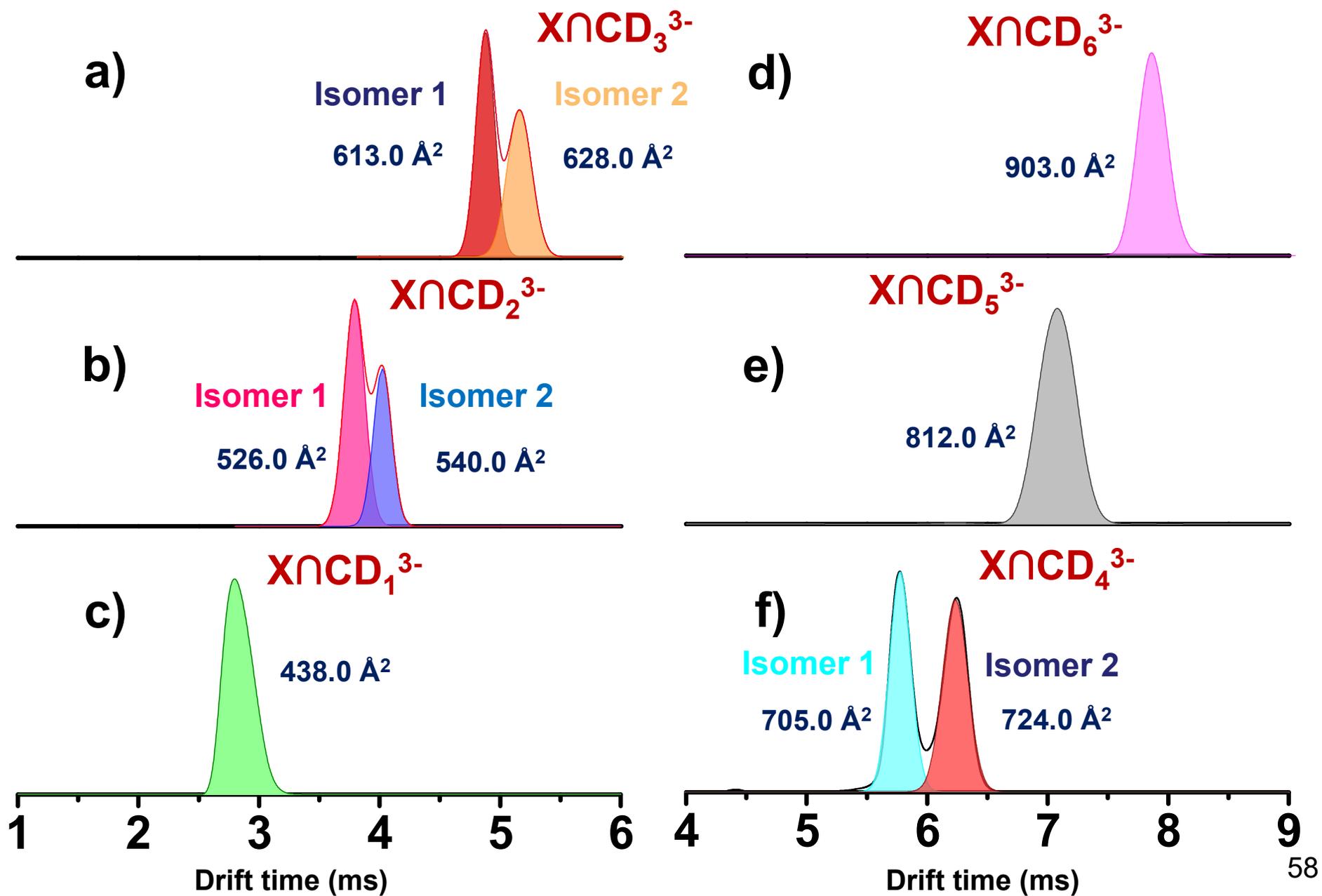
**Figure 1.** **A)** **(a)** Full range ESI MS, **(b)** experimental and calculated isotope patterns and **(c)** DFT optimized structure of  $[Ag_{29}(BDT)_{12}]^{3-}$  cluster. **B)** **(a)** ESI MS of  $[Ag_{29}(BDT)_{12}(C_{60})_n]^{3-}$  ( $n=1-4$ ) complexes, **(b)** experimental and calculated isotope patterns of  $[Ag_{29}(BDT)_{12}(C_{60})_4]^{3-}$  and **(c)** schematic of the possible structure of  $[Ag_{29}(BDT)_{12}(C_{60})_4]^{3-}$ .



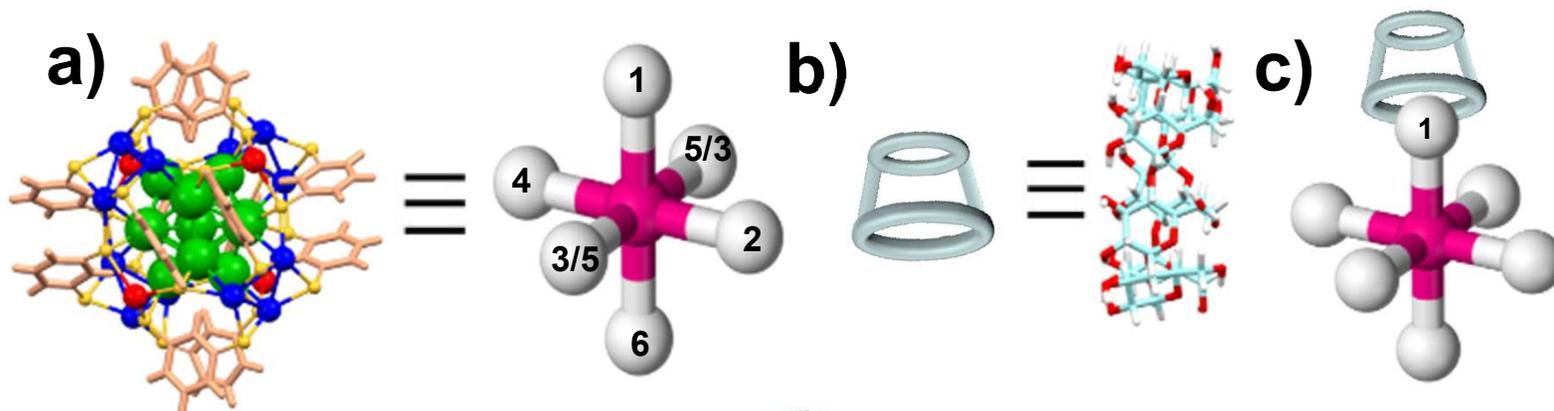
**Figure 3.** NMR of (a) C<sub>60</sub> showing peak at 143.59 ppm, (b) the adducts at a cluster:fullerene molar ratio of 1:4 showing peak at 143.03 ppm for the C<sub>60</sub> molecules in bound state and (c) the adducts at an excess concentration of C<sub>60</sub> (cluster:fullerene molar ratio of 1:15) showing a predominant peak for free C<sub>60</sub> (143.59 ppm) and a less intense peak for bound C<sub>60</sub> (143.03 ppm).

# Isomerism in supramolecular adducts

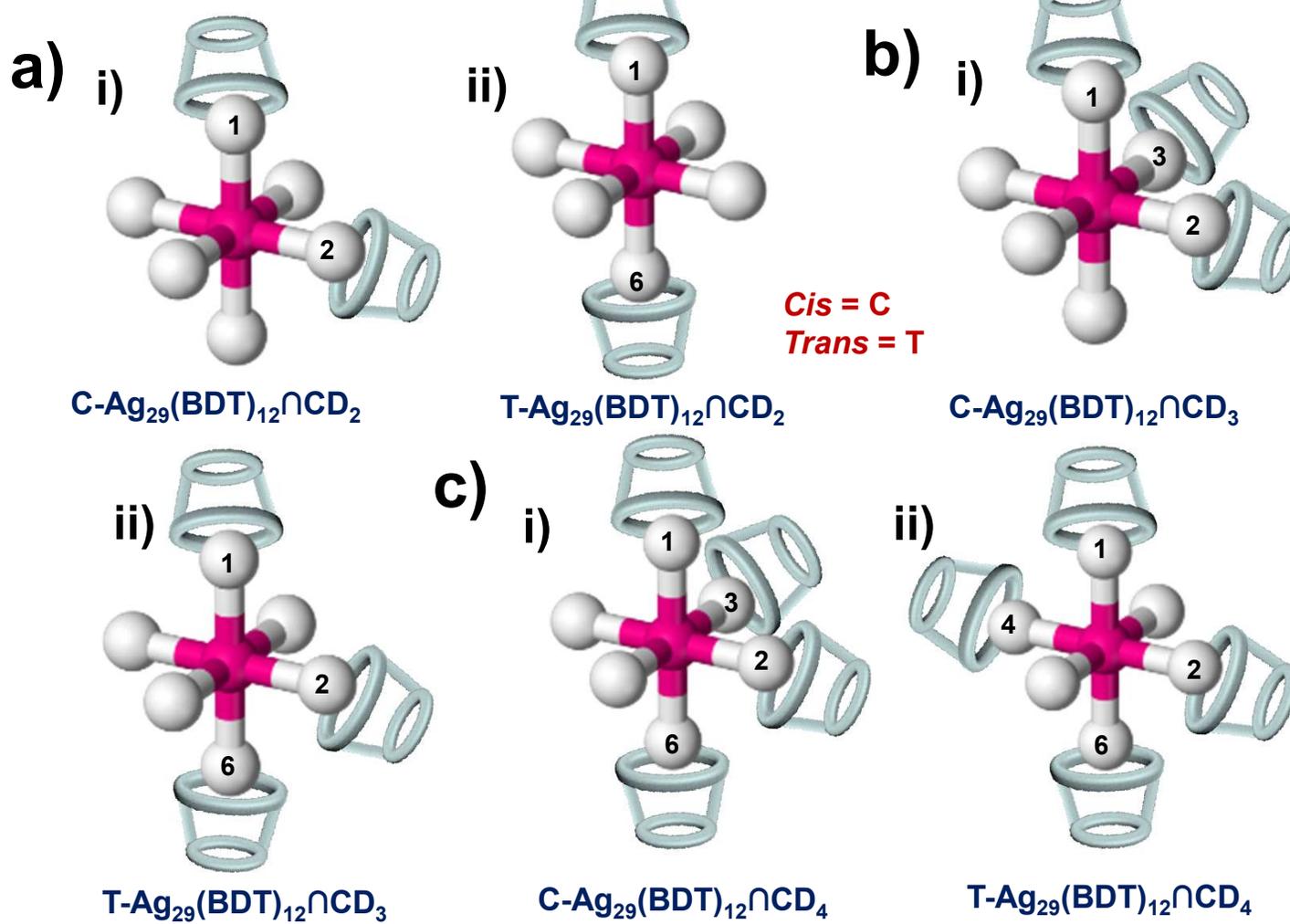




**A)**

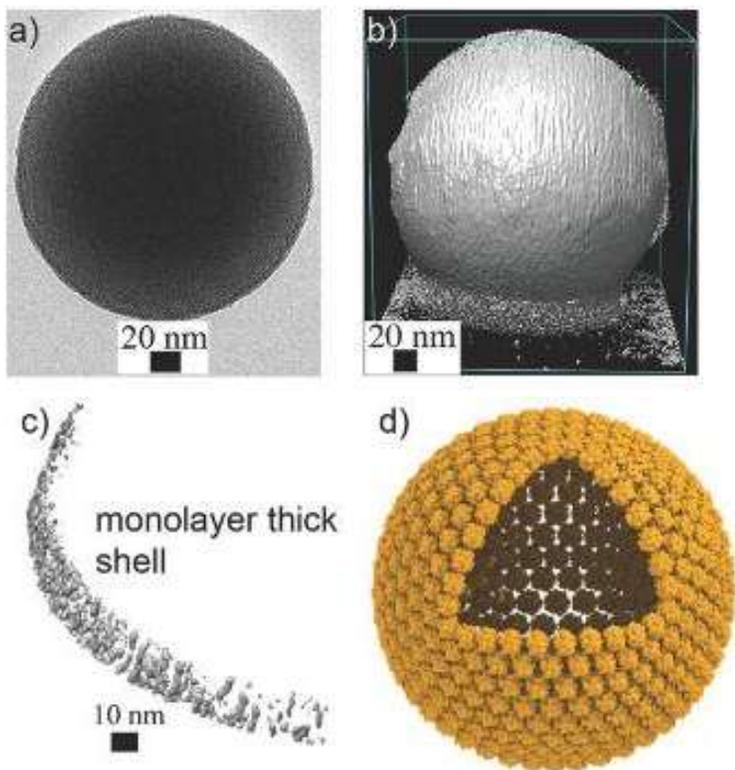
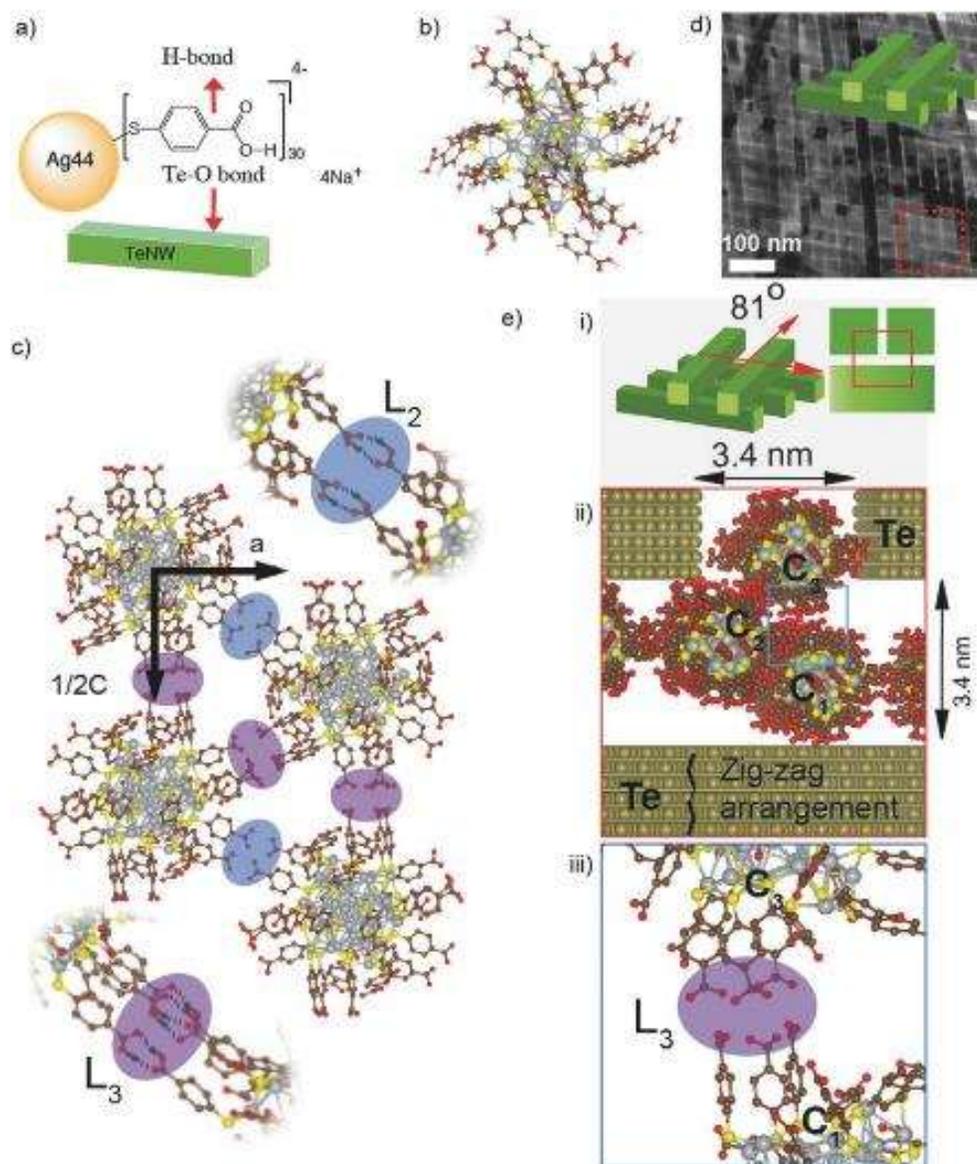
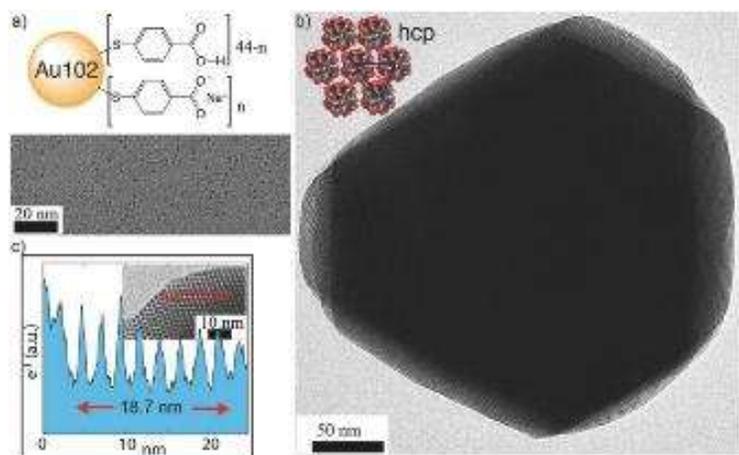


**B)**



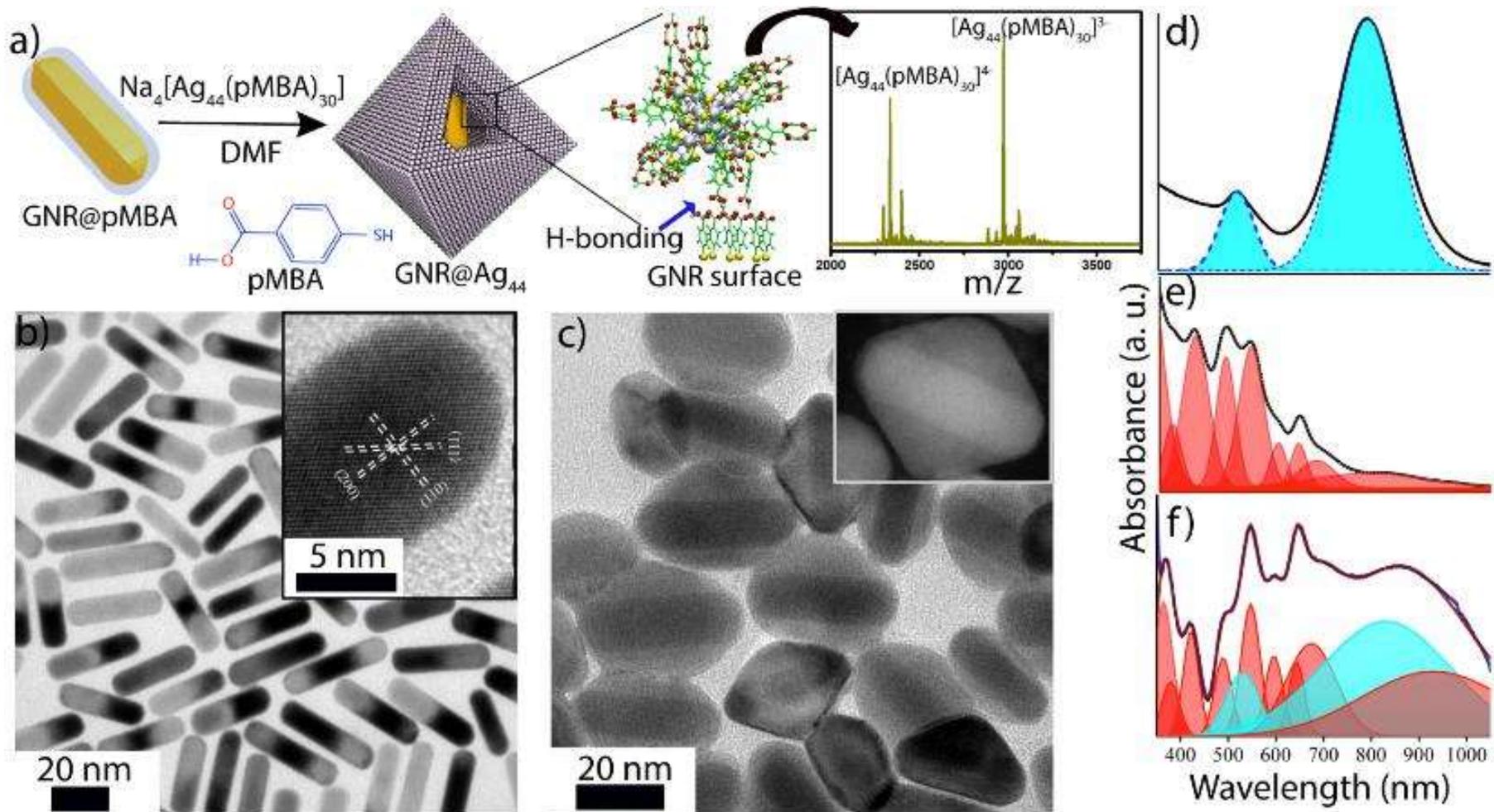
# **Assemblies and superstructures**

# Nanoclusters in colloidal assemblies



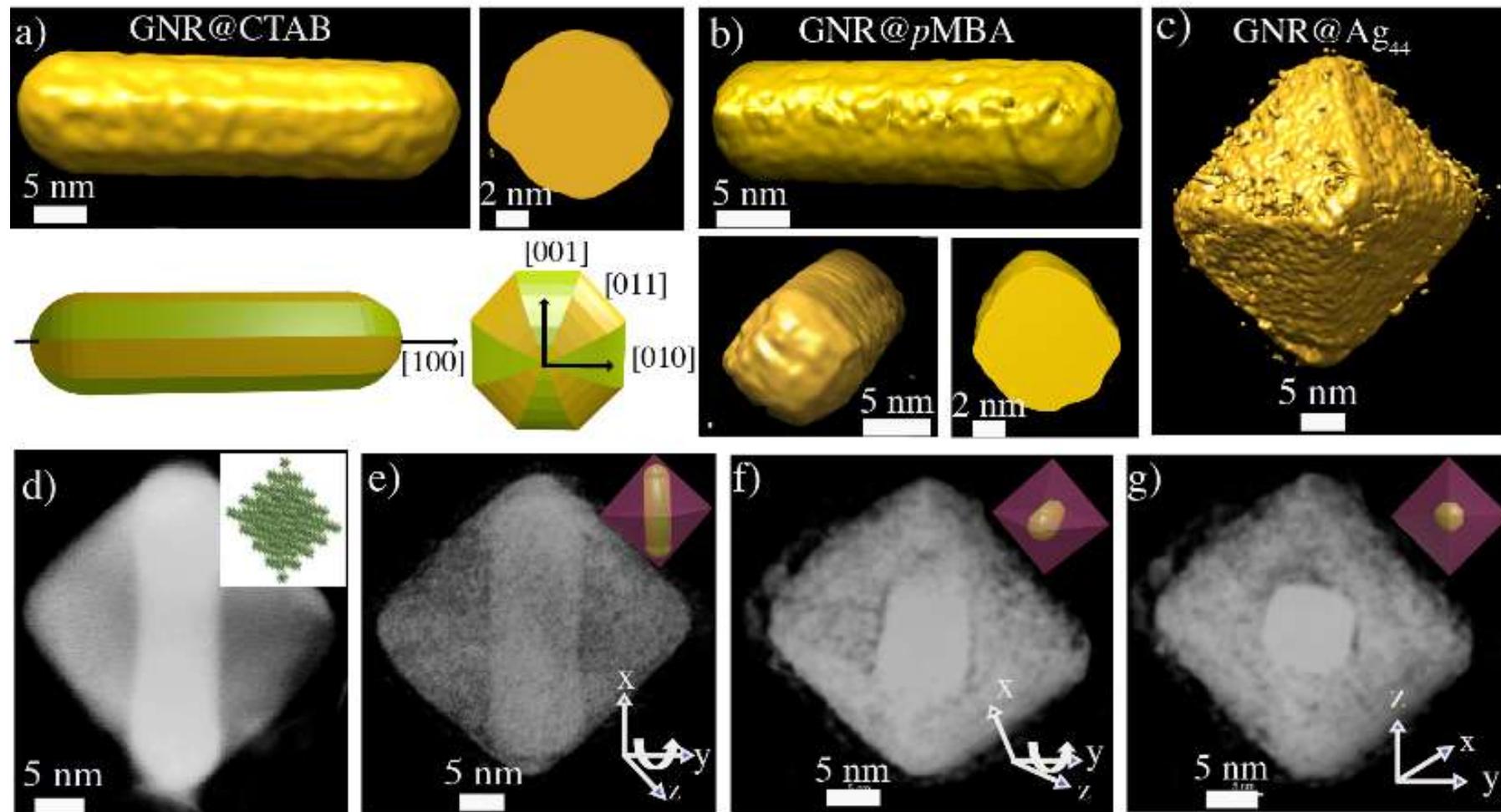
Som, A. et al., *Adv. Mater.* **2016**, 28, 2827–2833

# Atomically precise nanocluster assemblies encapsulating plasmonic gold nanorods

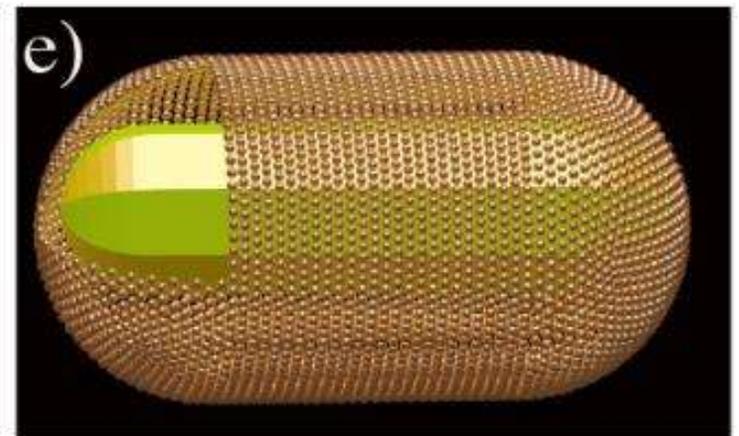
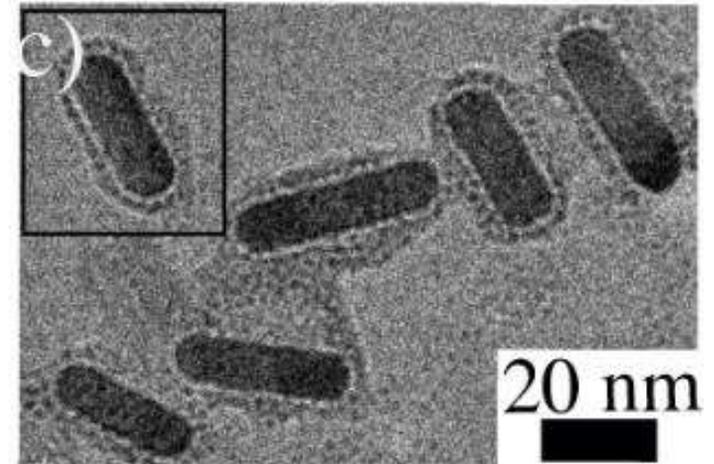
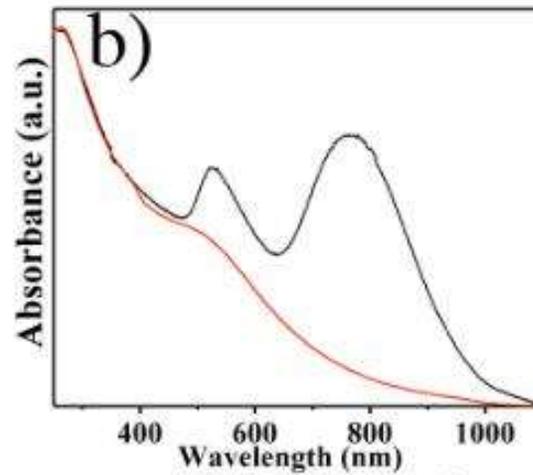
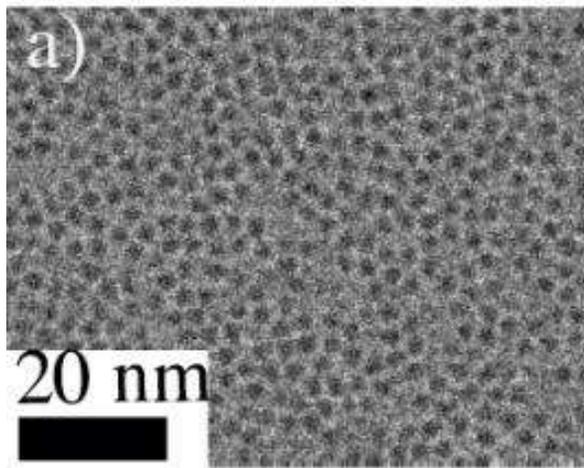


Chakraborty, A. et al., *Angew. Chem. Int. Ed.* 2018, 57, 6522–6526.

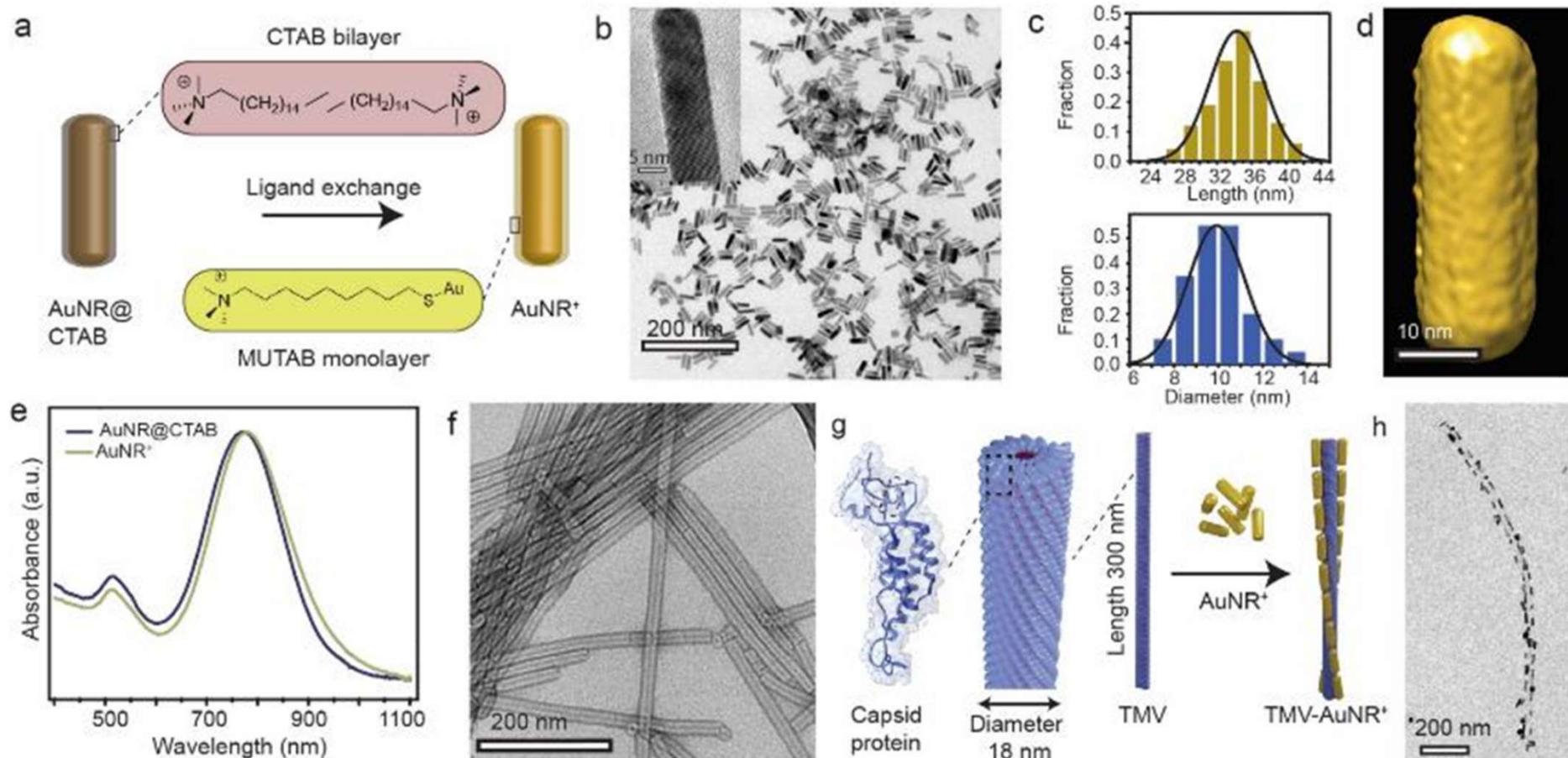
# 3D morphological analysis



# Works for $\text{Au}_{250}(\text{pMBA})_n$ and aqueous solvent



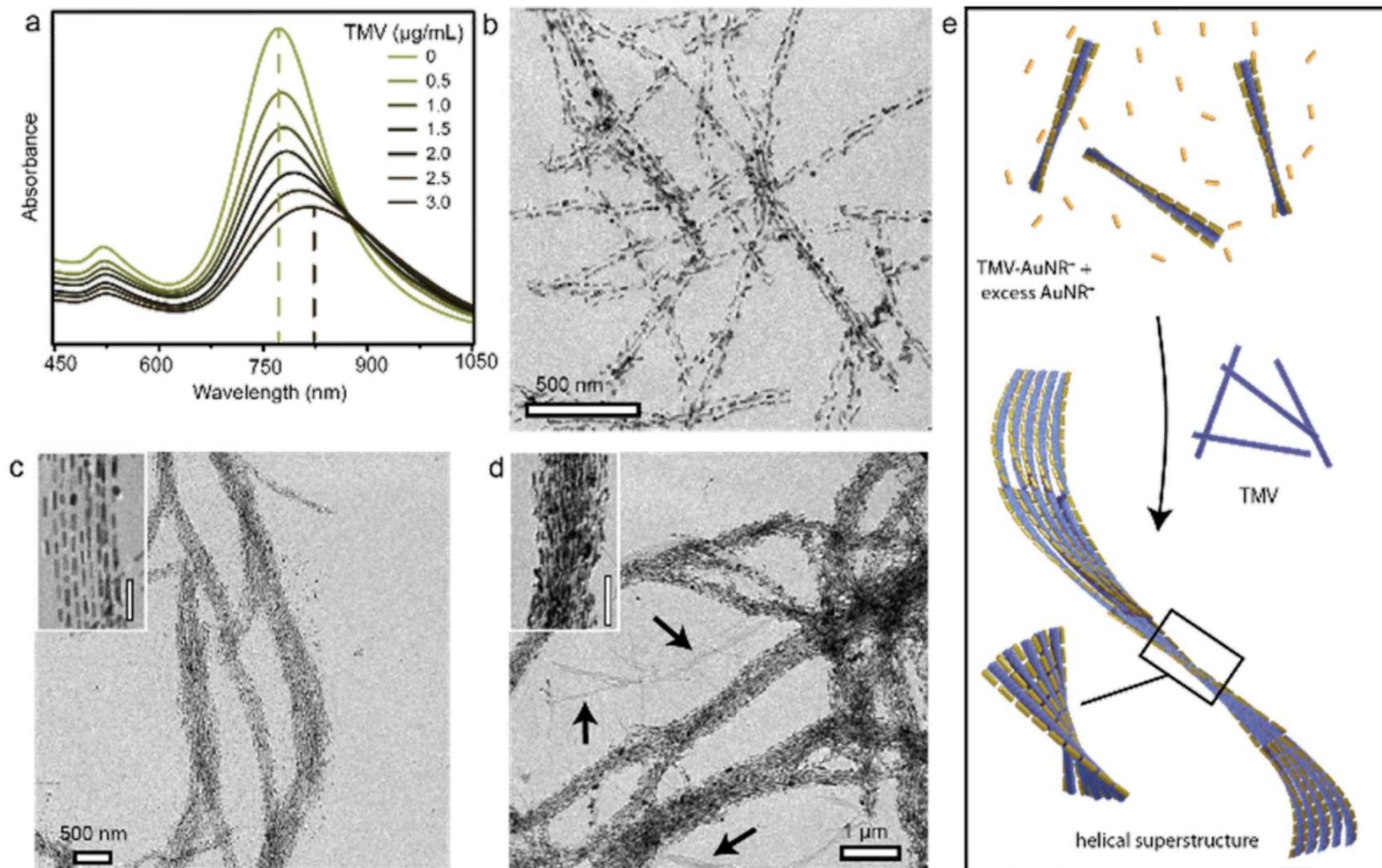
# Near-infrared chiral plasmonic microwires through precision assembly of gold nanorods on soft biotemplates



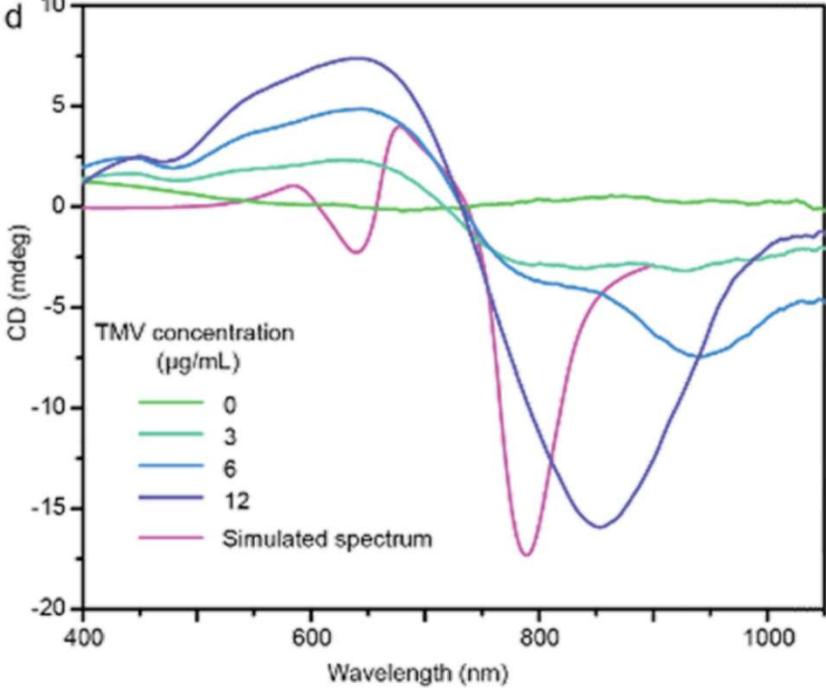
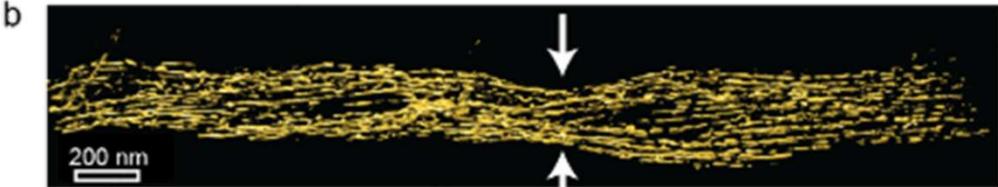
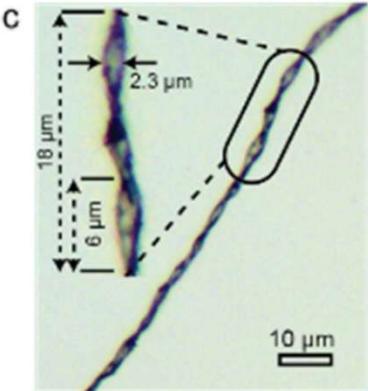
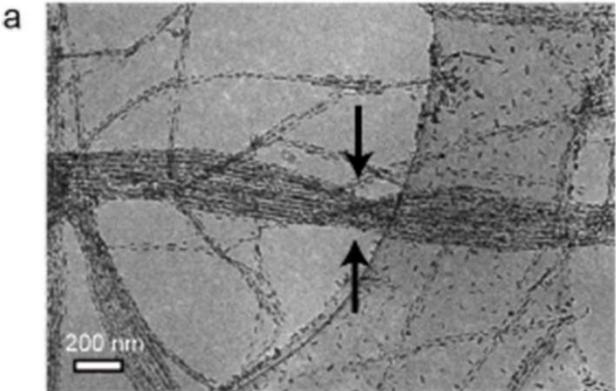
Chakraborty, A. et al., *J. Phys. Chem. C* 2021, 125, 3256–3267.

With Nonappa, Mauri A. Kostianen and Robin H. A. Ras

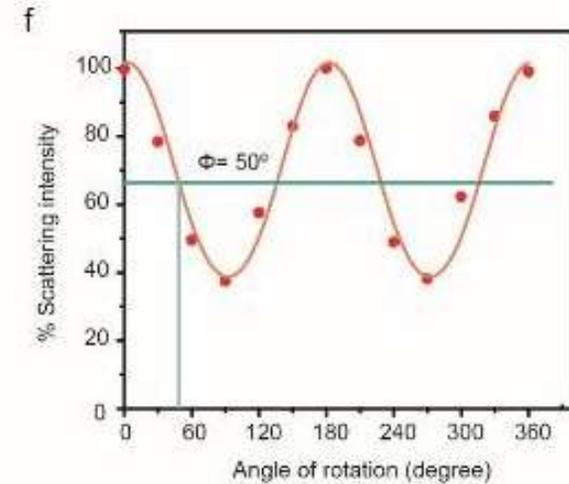
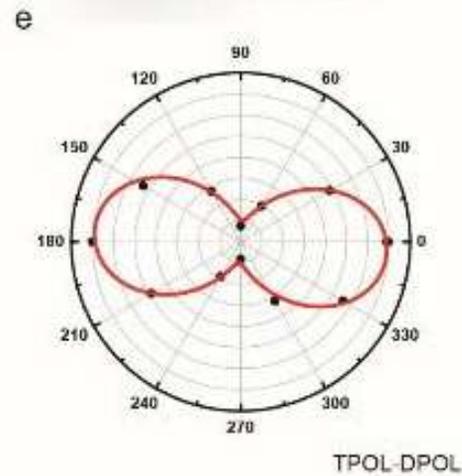
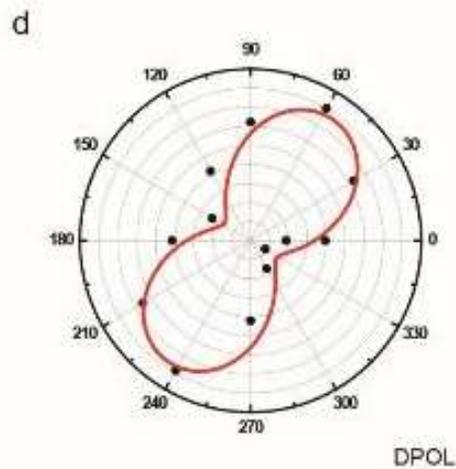
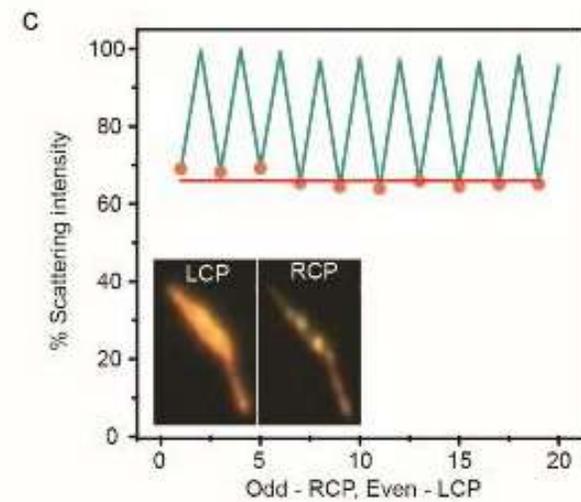
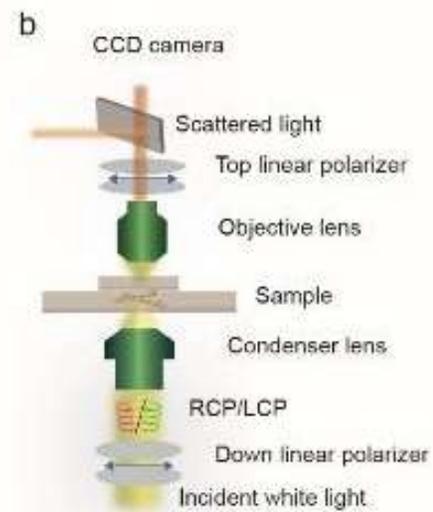
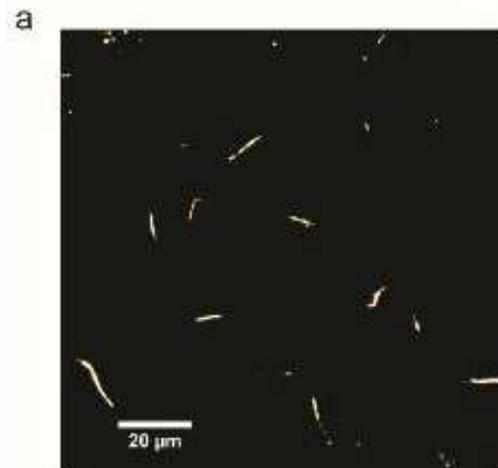
# Long-range assembly



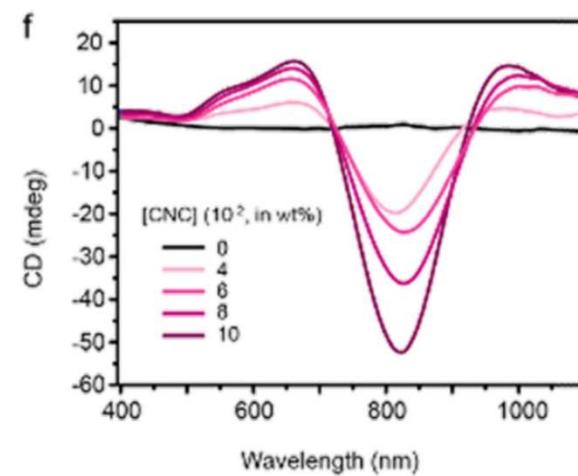
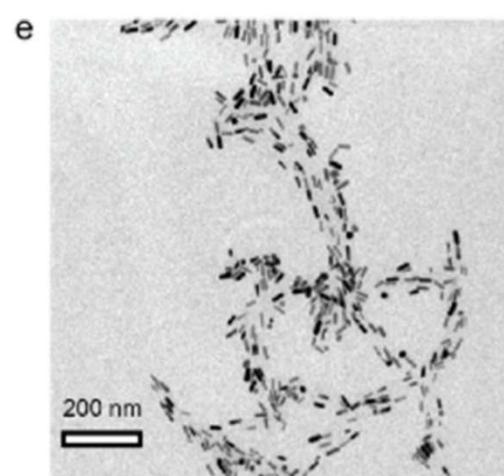
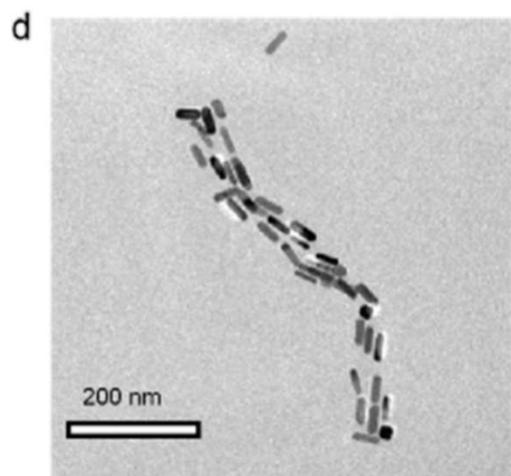
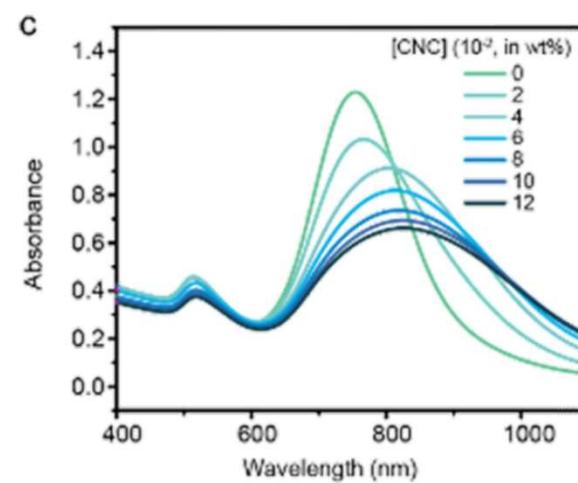
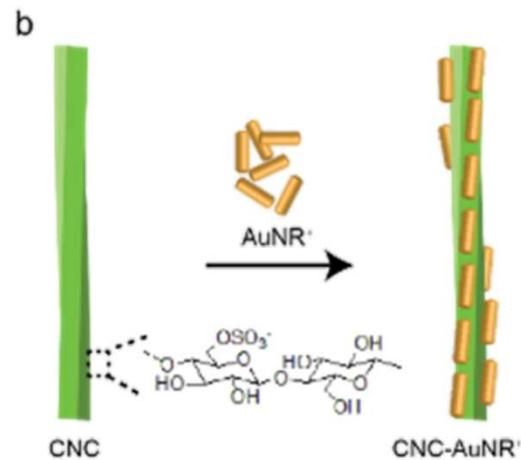
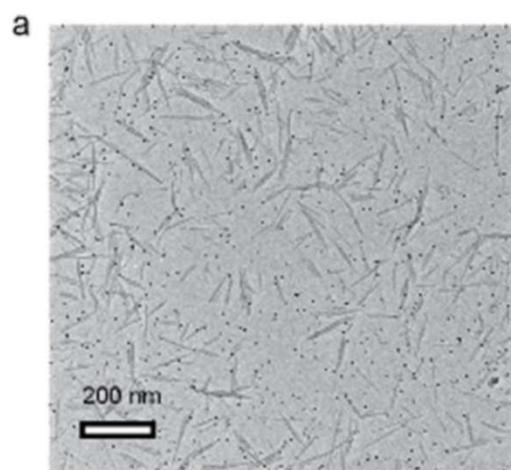
# NIR chirality



# Single particle chirality measurement



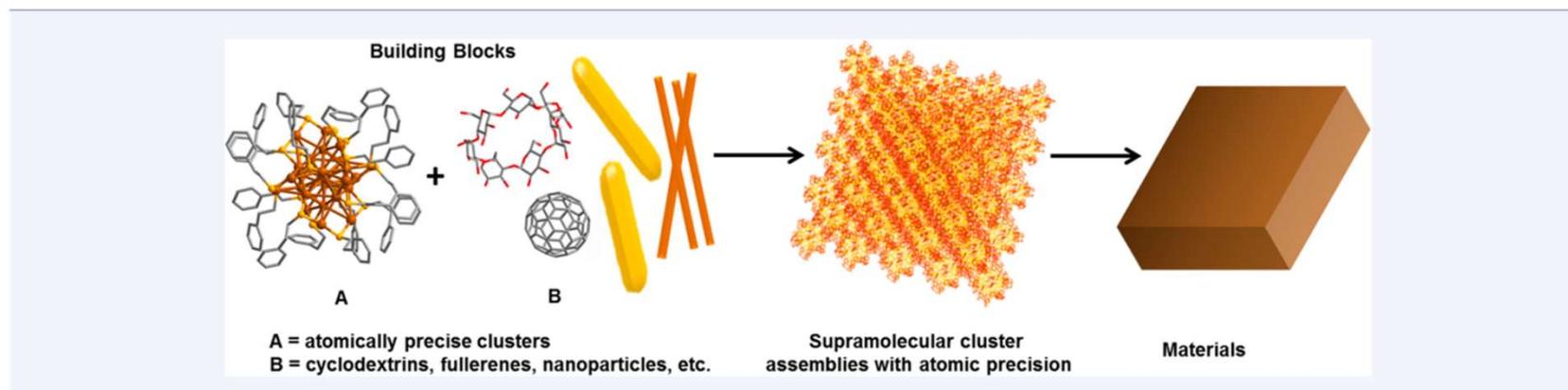
# With cellulose nanocrystals



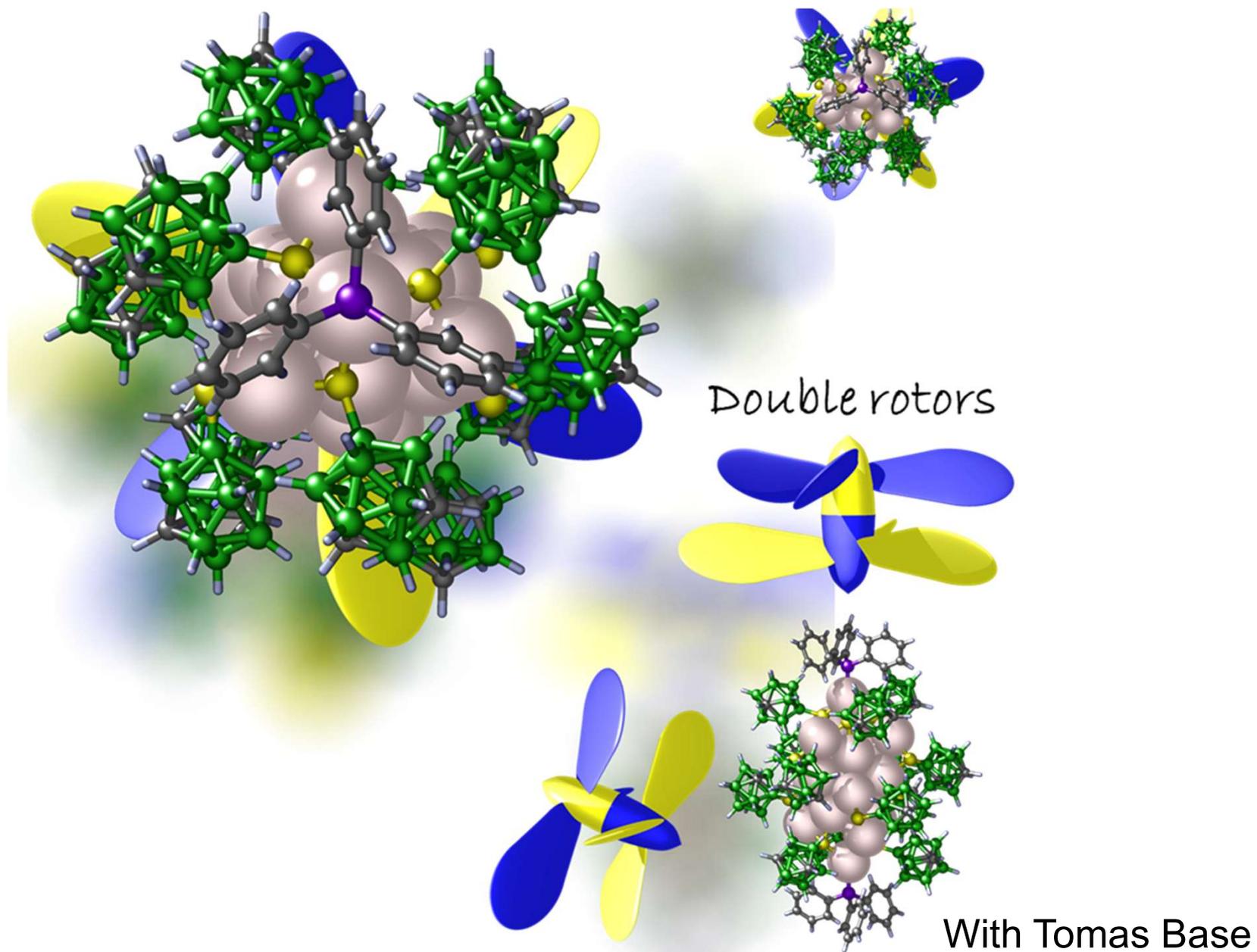
## 1 Approaching Materials with Atomic Precision Using Supramolecular 2 Cluster Assemblies 3

4 Papri Chakraborty, Abhijit Nag, Amrita Chakraborty, and Thalappil Pradeep\*<sup>ID</sup>

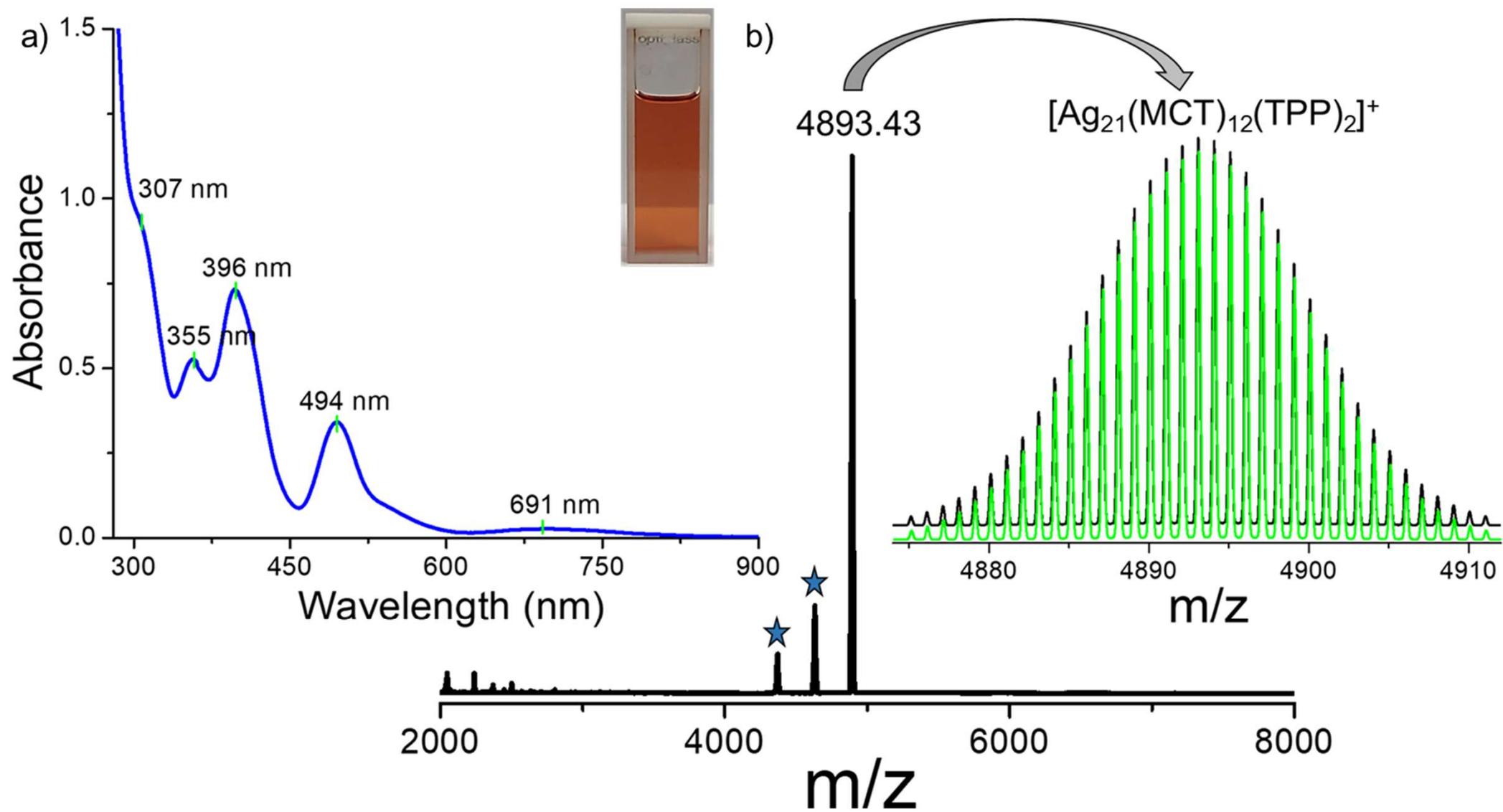
5 DST Unit of Nanoscience (DST UNS) and Thematic Unit of Excellence (TUE), Department of Chemistry, Indian Institute of  
6 Technology Madras, Chennai 600 036, India



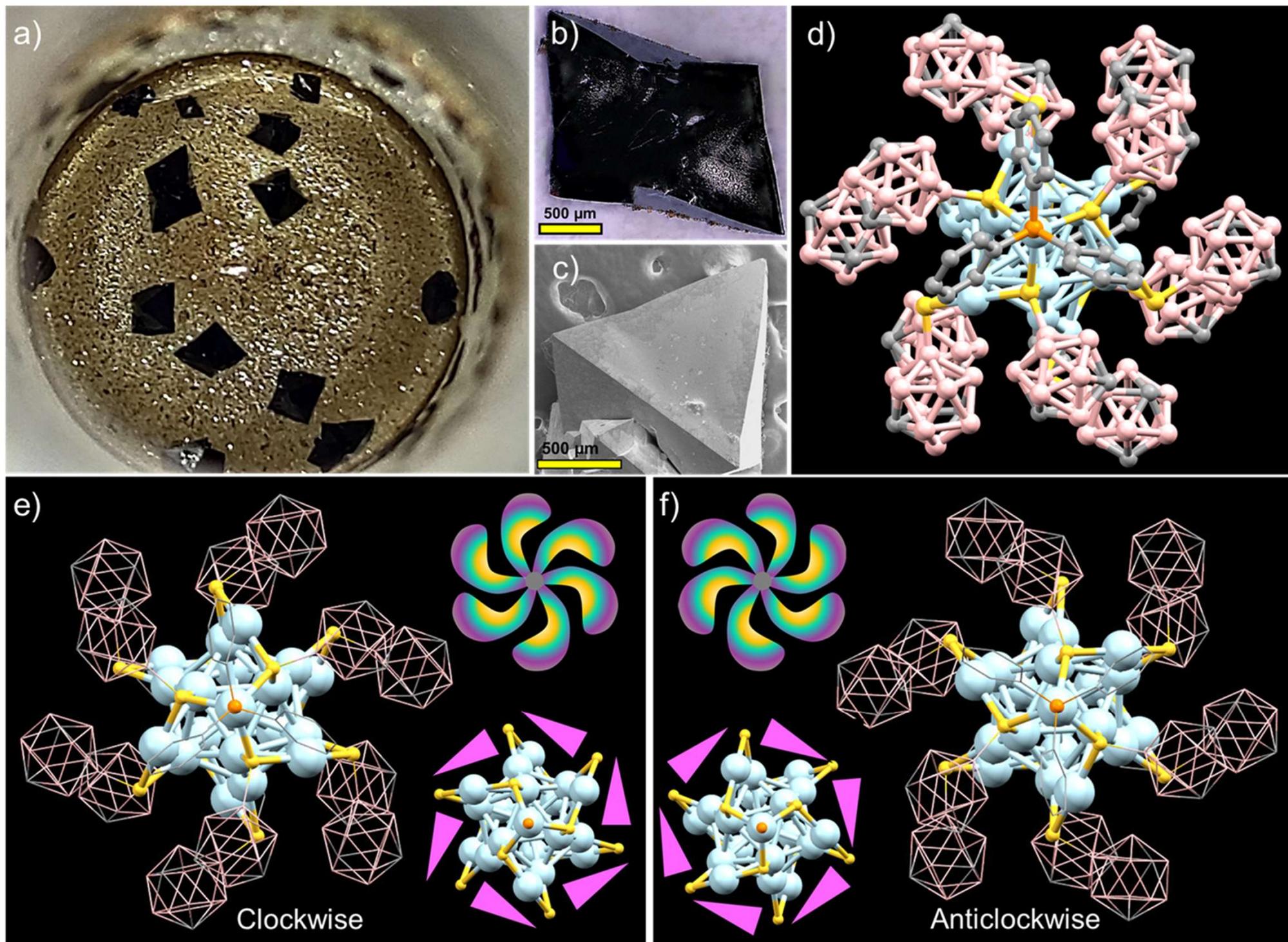
# Carborane-thiol protected propeller-shaped photoresponsive silver nanomolecule



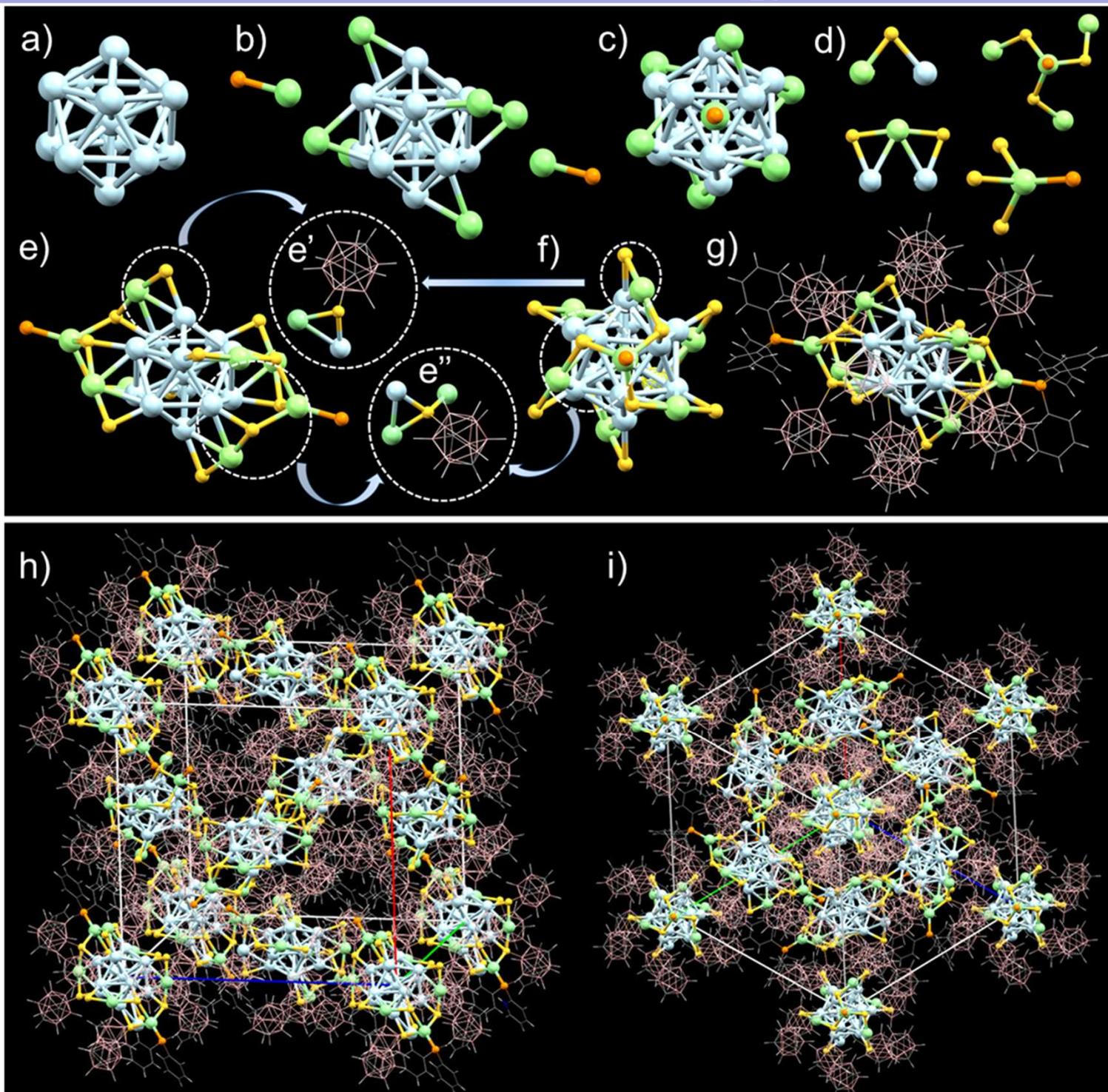
# Characterization of Ag<sub>21</sub>



# Structural details of $\text{Ag}_{21}$



# Structural details of $\text{Ag}_{21}$



# Molecules and their properties

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Chemical formula	H <sub>2</sub> O
Molecular weight	18.0148
Critical temperature	373.91°C
Critical pressure	22.05 MPa
Critical density	315.0 kg/m <sup>3</sup>
Triple point temperature	0.01°C
Triple point pressure	615.066 Pa
Normal boiling point	100.0°C
Normal freezing point	0.0°C
Density of ice at normal melting point	918.0 kg/m <sup>3</sup>
Maximum density, 3.98°C	999.973 kg/m <sup>3</sup>
Viscosity, 25°C	0.889 mN s/m <sup>2</sup>
Surface tension, 25°C	72 mN/m
Heat Capacity, 25°C	4.1796 kJ/kg.K
Enthalpy of vaporisation, 100°C	2,257.7 kJ/kg
Enthalpy of fusion, 0°C	333.8 kJ/kg
Velocity of sound, 0°C	1.403 km/s
Dielectric constant, 25°C	78.40
Electrical conductivity, 25°C	8 μS/m
Refractive index, 25°C	1.333
Liquid compressibility, 10°C	480. × 10 <sup>-12</sup> m <sup>2</sup> /N
Coefficient of thermal expansion, 25°C	256.32 × 10 <sup>-6</sup> K <sup>-1</sup>
Thermal Conductivity, 25°C	0.608 W/m.K

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Molecular formula  
Molecular weight  
Molecular structure  
Molecular absorption and emission  
Molecular reactions  
Molecular assembly  
Molecular co-crystals

-----  
Phases - phase transitions  
Physical properties  
Electrical, magnetic  
Mechanical properties  
Electrochemical properties

Future?

# Collaborators



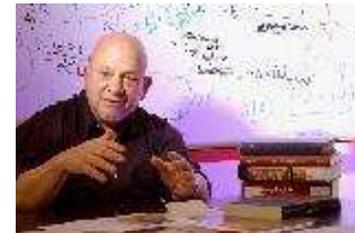
Robin Ras



Nonappa



Mauri Kostainen



Manfred Kappes



Olli Ikkala

Tatsuya Tsukuda, Keisaku Kimura, Yuichi Negishi, Uzi Landman, Hannu Hakkinen, Rob Whetten, Tomas Base





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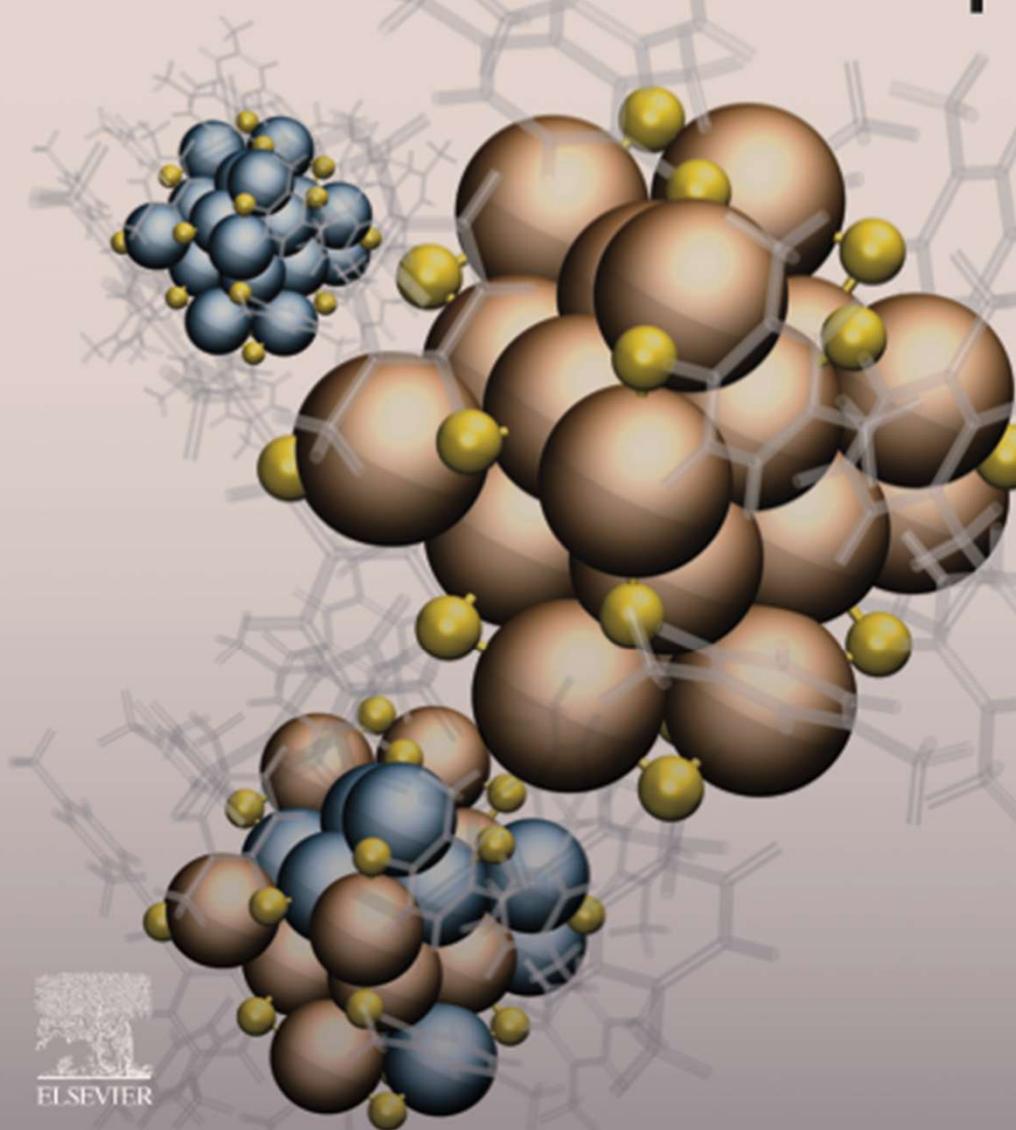
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# ATOMICALLY PRECISE METAL NANOCCLUSERS





## Department of Science and Technology

Collaborators: Tatsuya Tsukuda, Keisaku Kimura, Yuichi Negishi, Uzi Landman, Rob Whetten, Hannu Hakkinen, Robin Ras, Manfred Kappes, Horst Hahn, Tomas Base, Shiv Khanna, Umesh Waghmare, Chandrabhas Narayana, Giridhar U. Kulkarni, Reji Philip, Vivek Polshettiwar, R. Mukhopadhyay

**Thank you all**